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Effects of Globally Waste-Disturbing Activities on Gas Generation, Retention, and Release in Hanford Waste Tanks

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July 2005



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Richland, WA 99352

Abstract

Various operations are authorized in Hanford single- and double-shell tanks that disturb all or a large fraction of the waste. These globally waste-disturbing activities have the potential to release a large fraction of the retained flammable gas and to affect future gas generation, retention, and release behavior. This report presents analyses of the expected flammable gas release mechanisms and the potential release rates and volumes resulting from these activities. The background of the flammable gas safety issue at Hanford is summarized, as is the current understanding of gas generation, retention, and release phenomena. Considerations for gas monitoring and assessment of the potential for changes in tank classification and steady-state flammability are given.

Summary

This report presents an assessment of the effects of globally waste-disturbing activities on flammable gas generation, retention, and release. The assessment includes the potential for hazardous flammable gas releases, considerations for determining the frequency of gas monitoring during the disturbance, and the potential for long-term changes in tank classification and gas generation rate resulting from the disturbance. Analyses are based on historic waste conditions; “gel” and other waste forms outside of the Hanford experience base are not considered.

Global waste disturbances are those that affect all or a large fraction of the waste in a tank. The disturbance may involve draining the supernatant or interstitial liquid, dissolving solids, or mobilizing the settled sediment. Nine specific activities that occur or are approved in Hanford double-shell (DST) and single-shell (SST) waste tanks are considered in this report. While many of these operations could be performed in both SSTs and DSTs, all but two are considered specific to one type. The nine globally waste-disturbing activities are:

- Waste removal—pumping supernatant or slurry out of a DST
- Waste addition—adding liquid or slurry into a DST or SST
- Saltwell pumping—removing interstitial liquid from SSTs
- Saltcake dissolution—adding water to dissolve soluble solids in an SST while pumping out the resulting brine
- Water addition—adding water to a DST or SST without concurrent brine removal
- Mixer pump operation—mobilizing sediment in DSTs by the hydraulic jet forces of a mixer pump
- Airlift circulator operation—mobilizing sediment by airlift pumping in several DSTs
- Chemical addition—adding sodium hydroxide or nitrite solution to DSTs to prevent corrosion
- Evaporation—gradually concentrating supernatant in DSTs by evaporating water from the waste surface.

The considerations for each of these operations depend on the waste group classification of the tank. Three waste groups are defined based on the potential for flammability and whether the tank exhibits buoyant displacement gas release events (BDGReS). Waste Group A and B tanks store sufficient gas to make the headspace flammable if all of it were suddenly released; Waste Group C tanks do not. Waste Group A tanks exhibit spontaneous BDGReS, while tanks in Waste Group B do not. Tanks are assigned to waste groups based on several criteria for BDGRE behavior and the waste configuration and volume.

A waste-disturbing activity can cause or influence gas releases in three ways: 1) BDGReS in Waste Group A tanks may be amplified by decreasing the headspace and increasing the hydrostatic pressure on the gas (e.g., by adding waste, water, or chemicals), 2) BDGReS can be induced during the operation as a result of sediment dissolution or lowering the neutral buoyancy

gas fraction by suspending sediment in the supernatant (e.g., by mixer pump or airlift circulator operation) or otherwise increasing the supernatant density (e.g., by waste or chemical additions), and 3) the waste disturbance itself can produce gas releases that can eventually make the headspace flammable if the ventilation rate is low (e.g., saltwell pumping and saltcake dissolution in SSTs; sediment uncovery and dissolution in DSTs). The main concern is for operations in Group A and B tanks. Though gas releases occur in Group C tanks, they are inconsequential because these tanks do not have enough retained gas to make their headspaces flammable.

Besides releasing gas, a waste disturbance can also change a tank's waste group classification in the long term. A Group C tank experiencing BDGReS can move all the way into Group A by a waste addition that reduces its headspace sufficiently. Any operation that makes the supernatant denser can potentially decrease the neutral buoyancy gas fraction, initiate BDGReS, and move a tank from Group B to Group A. However, initiating BDGReS in a Group C tank will not change its classification. Waste, water, and chemical addition prove to be the only activities capable of moving a tank to a higher waste group except for waste removal in specific instances. Conversely, any of the waste removal activities could move a tank to Group C.

The hydrogen generation rate depends both on the amount and temperature of the waste, specifically the liquid portion, and the concentration of dissolved salts and total organic carbon in the liquid. Waste addition may cause an increase in the hydrogen generation rate, and any major removal of waste, especially liquid, will decrease it. Water additions that do not dissolve solids typically increase the total hydrogen generation rate. Caustic additions that do not dissolve or precipitate solids may potentially increase the hydrogen generation. Waste temperature can increase considerably from energy dissipated during mixer pump operation. Because hydrogen generation is sensitive to temperature, a relatively short period of continuous mixer pump operation could easily increase the hydrogen generation rate by an order of magnitude. If the waste temperature remains elevated for an extended period, the hydrogen generation rate and steady-state flammability hazard of the tank should be evaluated at the new temperature.

Contents

Abstract.....	iii
Summary	v
1.0 Introduction.....	1.1
1.1 History of the Flammable Gas Safety Issue.....	1.1
1.1.1 The Flammable Gas Watch List.....	1.2
1.1.2 The Flammable Gas USQ	1.2
1.1.3 Tri-Party Agreement Milestone M-40-00.....	1.3
1.1.4 DNFSB Recommendation 93-5	1.3
1.1.5 Flammable Gas Safety Issue Resolution.....	1.4
1.2 Local and Global Waste Disturbances.....	1.4
1.2.1 Gas Releases Observed During Local Disturbances.....	1.4
1.2.2 Gas Releases Observed During Global Waste Disturbances	1.5
1.3 Current and Planned Operations in Hanford Tanks.....	1.7
1.3.1 Waste Removal	1.7
1.3.2 Waste Addition	1.7
1.3.3 Saltwell Pumping	1.8
1.3.4 Saltcake Dissolution.....	1.8
1.3.5 Water Addition.....	1.9
1.3.6 Mixer Pump Operation	1.9
1.3.7 Airlift Circulator Operation	1.9
1.3.8 Chemical Addition	1.9
1.3.9 Natural Evaporation	1.10
2.0 Physics of Gas Generation, Retention, and Release	2.1
2.1 Gas Generation	2.1
2.1.1 Gas Generation Theory	2.2
2.1.2 Hydrogen Generation Model	2.3
2.2 Gas Bubble Retention Mechanisms	2.4
2.2.1 Bubbles Retained by Capillary Forces.....	2.5
2.2.2 Gas Retained by Waste Strength.....	2.6
2.3 Gas Release Mechanisms and Behavior	2.7
2.3.1 Liquid Waste Configuration	2.8
2.3.2 Liquid-over-Solid Configuration	2.8
2.3.3 Wet Sediment Configuration	2.11
2.3.4 Pumped Configuration	2.12
2.4 Waste Classification	2.12
2.4.1 Waste Group Definition	2.12
2.4.2 Criteria for Waste Group Classification.....	2.13

3.0 Gas Releases Induced by Global Waste Disturbances.....	3.1
3.1 Waste Removal	3.2
3.1.1 Mechanisms for Gas Releases During Waste Removal.....	3.2
3.1.2 Expected Gas Releases During Waste Removal.....	3.6
3.1.3 Gas Monitoring Considerations During Waste Removal	3.9
3.1.4 Potential Changes in Waste Group Classification After Waste Removal	3.12
3.1.5 Potential Changes in Hydrogen Generation After Waste Removal.....	3.12
3.2 Waste Addition	3.13
3.2.1 Mechanisms for Gas Releases During Waste Addition.....	3.13
3.2.2 Expected Gas Releases During Waste Addition	3.17
3.2.3 Gas Monitoring Considerations During Waste Addition.....	3.18
3.2.4 Potential Changes in Waste Classification After Waste Addition.....	3.19
3.2.5 Potential Changes in Hydrogen Generation After Waste Addition	3.21
3.3 Saltwell Pumping.....	3.21
3.3.1 Mechanisms for Gas Release During Saltwell Pumping	3.22
3.3.2 Expected Gas Releases During Saltwell Pumping.....	3.23
3.3.3 Gas Monitoring Considerations During Saltwell Pumping	3.26
3.3.4 Potential Changes in Waste Classification After Saltwell Pumping.....	3.27
3.3.5 Potential Changes in Hydrogen Generation After Saltwell Pumping.....	3.27
3.4 Saltcake Dissolution	3.27
3.4.1 Mechanisms for Gas Release During Saltcake Dissolution.....	3.28
3.4.2 Expected Gas Releases During Saltcake Dissolution	3.28
3.4.3 Gas Monitoring Considerations During Saltcake Dissolution.....	3.29
3.4.4 Potential Changes in Waste Classification After Saltcake Dissolution	3.32
3.4.5 Potential Changes in Hydrogen Generation After Saltcake Dissolution	3.32
3.5 Water Addition	3.32
3.5.1 Mechanisms for Gas Release During Water Addition.....	3.32
3.5.2 Expected Gas Releases During Water Addition	3.33
3.5.3 Gas Monitoring Considerations During Water Addition.....	3.36
3.5.4 Potential Changes in Waste Classification After Water Addition	3.36
3.5.5 Potential Changes in Hydrogen Generation After Water Addition	3.39
3.6 Mixer Pump Operation	3.39
3.6.1 Mechanisms for Gas Releases During Mixer Pump Operation.....	3.39
3.6.2 Expected Gas Releases During Mixer Pump Operation	3.40
3.6.3 Gas Monitoring Considerations During Mixer Pump Operation.....	3.43
3.6.4 Potential Changes in Tank Classification After Mixing	3.44
3.6.5 Potential Changes in Hydrogen Generation After Mixing.....	3.45
3.7 Airlift Circulator Operation	3.45
3.7.1 Mechanisms for Gas Releases During ALC Operation	3.45
3.7.2 Expected Gas Releases During ALC Operation	3.47
3.7.3 Gas Monitoring Considerations During ALC Operation.....	3.53
3.7.4 Potential Changes in Tank Classification After ALC Operation.....	3.53
3.7.5 Potential Changes in Hydrogen Generation After ALC Operation	3.53

3.8 Chemical Addition.....	3.53
3.8.1 Mechanisms for Gas Release During Chemical Addition	3.54
3.8.2 Expected Gas Releases During Chemical Addition.....	3.55
3.8.3 Gas Monitoring Considerations During Chemical Addition	3.55
3.8.4 Potential Changes in Waste Classification After Chemical Addition.....	3.55
3.8.5 Potential Changes in Hydrogen Generation After Chemical Addition.....	3.56
3.9 Natural Evaporation.....	3.56
3.9.1 Mechanisms for Gas Releases During Evaporation.....	3.57
3.9.2 Expected Gas Releases During Evaporation.....	3.59
3.9.3 Gas Monitoring Considerations During Evaporation	3.59
3.9.4 Potential Changes in Tank Classification After Evaporation	3.59
3.9.5 Potential Changes in Hydrogen Generation After Evaporation.....	3.62
4.0 Summary and Conclusions	4.1
4.1 Potential for Significant Gas Releases.....	4.1
4.2 Potential Changes in Waste Group Classification	4.3
4.3 Potential Changes in Hydrogen Generation.....	4.4
4.4 Implications for Other Operations	4.5
4.4.1 Local Waste Disturbances.....	4.5
4.4.2 Other Global Waste Disturbances.....	4.6
4.4.3 Summary	4.7
5.0 References.....	5.1
Appendix - Technical Background for Dissolution and Precipitation.....	A.1

Figures

3.1	Sensitivity of Peak Hydrogen Concentration to Ventilation Rate in AN-105	3.10
3.2	Predicted Hydrogen Concentration in AN-103 Resulting from Sediment Draining	3.11
3.3	Mechanism for Inducing BDGReS During Waste Addition	3.14
3.4	Tank Waste Configuration During and After Waste Transfer.....	3.15
3.5	Effects of Waste Addition on Buoyancy Ratio for Waste Group C Tank	3.20
3.6	Effects of Waste Addition on Buoyancy Ratio for Waste Group B Tank	3.20
3.7	U-105 Hydrogen Release and Saltwell Pumping Rates.....	3.23
3.8	Tank 241-S-106 Estimated Hydrogen Release Rate and Pumping Rate	3.24
3.9	U-105 Ammonia Release and Waste Volume Pumped	3.25
3.10	Hydrogen Concentration and Waste Level History in S-111	3.26
3.11	Hydrogen Concentration Versus Time for Saltcake Dissolution; Group B Tank	3.30
3.12	Required Gas Monitoring Frequency Versus Brine Pumping Rate.....	3.31
3.13	Sensitivity of Hydrogen Concentration to Ventilation in AN-103 During Mixing	3.44
3.14	Schematic of Typical Airlift Circulator Operation	3.46
3.15	Location of ALCs in AY and AZ DSTs	3.47
3.16	Stages of Waste Disturbance During ALC Operation	3.48
3.17	ALC Flow Rate and Velocity	3.51
3.18	Gas Release Rate from a Single ALC	3.52
3.19	Tank Waste Configuration Before and After Evaporation	3.58
3.20	Effects of Evaporation on Buoyancy Ratio	3.60
3.21	Water Partial Pressure Versus Solution Density.....	3.61

Tables

3.1	Summary of Globally Waste-Disturbing Activities.....	3.1
3.2	Summary of Supernatant Decant Analysis Results	3.7
3.3	Inputs Used for Waste Transfer Analysis	3.20
3.4	Peak Hydrogen Concentrations with Dissolution After Water and Caustic Addition.....	3.36
3.5	Peak Hydrogen Concentrations Without Dissolution After Water and Caustic Addition	3.38
3.6	Time to Return to BDGRE Behavior After Degassing.....	3.40
3.7	Summary of Mixing Gas Release Analysis Results	3.42
3.8	Waste Properties for ALC Operation Analysis.....	3.50
3.9	Disturbed Sediment Volumes	3.50
3.10	Change in Buoyancy Ratio for Addition of 100 kgal at SpG = 1.5	3.56
3.11	Tank Data Used for Evaporation Analysis	3.60
4.1	Potential Gas Releases During Globally Waste-Disturbing Activities.....	4.2
4.2	Potential Changes in Waste Group from Globally Waste-Disturbing Activities	4.3
4.3	Changes in Gas Generation from Globally Waste-Disturbing Activities.....	4.5

1.0 Introduction

Various operations in Hanford high-level radioactive waste storage tanks may disturb a large fraction of the waste. These “globally waste-disturbing activities” are of interest because they are now being performed or are planned as part of the Hanford mission to retrieve waste from the older single-shell tanks (SSTs), transfer it to the newer double-shell tanks (DSTs), and deliver waste feed from the DSTs to the vitrification plant for treatment.

Most of the waste tanks retain some volume of flammable gas, and some tanks retain a substantial gas volume. Because globally waste-disturbing activities potentially can release a large fraction of that gas quickly, all such operations must be approached with care. The potential gas releases from these kinds of activities have long been subject to intense scrutiny brought about by the Flammable Gas Safety Issue. However, that safety issue has been formally closed along with the associated Unreviewed Safety Questions (USQ), and the remaining tanks were removed from the Flammable Gas Watch List on August 17, 2001.

This report presents an assessment of the predicted effects of globally waste-disturbing activities on flammable gas generation, retention, and release. The assessment includes the potential for hazardous flammable gas releases, considerations for determining the frequency of gas monitoring during the disturbance, and the potential for long-term changes in tank classification and gas generation rate resulting from the disturbance. Analyses are based on historic Hanford waste conditions and processes studied from the mid-1990s; “gel” and other waste forms outside of the Hanford experience base that might result from future retrieval and treatment scenarios are not considered.

The rest of the Introduction sets the stage for a detailed discussion of each globally waste-disturbing activity. Section 1.1 summarizes the history of the Flammable Gas Safety Issue from inception to closure. Section 1.2 describes the difference between global and local waste disturbances and the history of gas release observations for each type. Section 1.3 introduces the globally waste-disturbing activities to be assessed for gas releases.

The mechanisms for gas generation, retention, and release for the major Hanford waste types and waste configurations are reviewed, and the revised tank waste group classifications and criteria are outlined in Section 2. The potential for significant gas releases for each activity, gas-monitoring issues, and potential changes in tank classification and gas generation rate are discussed in Section 3. Section 4 summarizes our conclusions, and references are listed in Section 5.

1.1 History of the Flammable Gas Safety Issue

All radioactive wastes in the Hanford underground storage tanks slowly generate a gas mixture that typically includes hydrogen, nitrous oxide, ammonia, nitrogen, and traces of methane and other organic compounds. In most tanks, the gas evolves from the waste into the tank headspace at the same rate as it is generated, so the existing active and passive ventilation systems can keep the concentration far below the lower flammability limit (LFL). However,

some tanks may retain enough gas within the waste to cause worker injury or damage to equipment if it were suddenly released into the headspace and ignited. Gas releases may occur spontaneously or be induced by external forces (e.g., severe earthquakes) or waste intrusion activities (e.g., core sampling, equipment installation, liquid removal operations). The potential for spontaneous releases and their undesirable consequences constituted the Flammable Gas Safety Issue.

The Flammable Gas Safety Issue was born in early 1990 when the large gas releases occurring in Tank SY-101 were declared hazardous. Since then this issue has been documented or acknowledged in the “Watch List,” several USQs, a major milestone of the Tri-Party Agreement (Ecology et al. 1996), and a Defense Nuclear Facilities Safety Board (DNFSB) recommendation, as summarized below. A complete discussion of the evolution and closure of the Flammable Gas Safety Issue is given by Johnson et al. (2001).

1.1.1 The Flammable Gas Watch List

In November 1990, Public Law 101-510 was passed. Section 3137 of this law, also known as the Wyden Amendment, required the Secretary of Energy to identify the high-level nuclear waste tanks that may have “serious potential for release of high-level waste due to uncontrolled increases in temperature or pressure.”

Twenty-three tanks were identified in 1991 by a variety of criteria in response to the public law, and two more tanks were added in 1993, for a total of 25 on the Watch List. Tank SY-101 was removed from the list in January 2001. The remaining 24 tanks, A-101, AN-103, AN-104, AN-105, AX-101, AX-103, AW-101, S-102, S-111, S-112, SX-101, SX-102, SX-103, SX-104, SX-105, SX-106, SX-109, SY-103, T-110, U-103, U-105, U-107, U-108, and U-109, were removed in August 2001. These tanks were also the subject of TPA Milestone M-40-00.

1.1.2 The Flammable Gas USQ

An unreviewed safety question (USQ) exists when DOE determines that the nuclear safety Authorization Basis governing the activities at a facility or site may not adequately protect the health and safety of the public, workers, or the environment. DOE declared a USQ in 1990 because the existing controls and analyses did not consider the observed gas releases in SY-101 and postulated phenomena in other tanks. In June 1996, DOE-RL closed the USQ for SY-101 because the mixer pump had mitigated the spontaneous gas release hazard.^(a)

The original USQ was updated in July 1996, consolidating earlier determinations into one overall flammable gas USQ determination that was adopted by DOE-RL on November 1, 1996.^(b) The expanded USQ applied to 149 SSTs and 27 DSTs (excluding SY-101). The USQ

(a) Wagoner JD. June 21, 1996. *Closure of Hanford Tank 241-SY-101 Flammable Gas Unreviewed Safety Question (USQ)*. Letter 96-WSD-060 to AL Trego (WHC). DOE-RL, Richland, WA.

(b) Wagoner JD. November 1, 1996. *The U.S. Department of Energy (DOE) Richland Operations Office (RL) Definition and Declaration of Flammable Gas Unreviewed Safety Question (USQ)*. Letter 96-WSD-283 to HJ Hatch (FDH). DOE-RL, Richland, WA.

for AX-104 was closed by DOE-RL in April 1998^(a) and for all the other tanks in September 1998^(b) based on the adequacy of flammability, ignition, and monitoring controls implemented as part of the Authorization Basis. Another USQ (TF-97-0975) was declared for Tank SY-101 in 1998 because of an unexpected rise in the waste level.^(c) This USQ was closed in November 2000.^(d)

1.1.3 Tri-Party Agreement Milestone M-40-00

The Tri-Party Agreement (Ecology et al. 1996) was established between the Washington State Department of Ecology, the U.S. Environmental Protection Agency, and DOE in 1989. TPA Milestone M-40-00, “Mitigate/Resolve Tank Safety Issues for High Priority Watch List Tanks,” was among the specific milestones established for the various programs at the Hanford Site. This milestone, established in January 1994 with a due date of September 30, 2001, required that corrective action plans be developed for the watch list tanks and that mitigation activities, if required, be implemented to ensure safe storage of waste until retrieval began for treatment and/or disposal operations.

The safety issues for ferrocyanide and organic/nitrate were resolved by showing that these compounds did not exist or that their concentrations were so low that self-sustaining propagating chemical reactions were not plausible. The high-heat issue was limited to one SST (C-106) and was resolved by transferring most of the waste to a DST (AY-102). The flammable gas issue was quite different in that flammable gases are present in all of the waste tanks. However, evaluations showed that the hazard could be managed by following the established control process. When the Flammable Gas Safety Issue was closed in August 2001, the TPA milestone was completed one month ahead of schedule.

1.1.4 DNFSB Recommendation 93-5

In July 1993, the DNFSB transmitted Recommendation 93-5 (Conway 1993) on the Hanford Waste Tank Characterization Studies to DOE. Recommendation 93-5 noted that 1) technical information available for the Hanford tank waste was insufficient to ensure that wastes could be safely stored and associated operations could be conducted safely and 2) the characterization effort should be upgraded and expedited.

(a) Wagoner JD. April 21, 1998. *Closure of Flammable Gas Unreviewed Safety Question (USQ) for Tank 241-AX-104*. Letter 98-TWR-011 to HJ Hatch (FDH). DOE-RL, Richland, WA.

(b) Wagoner JD. September 25, 1998. *Closure of Flammable Gas Unreviewed Safety Question (USQ) for Tank Waste Remediation System (TWRS) Single-Shell Tanks (SST) and Double-Shell Tanks (DST)*. Letter 98-SCD-111 to RD Hanson (FDH). DOE-RL, Richland, WA.

(c) Kinzer JE. February 26, 1998. *Declaration of an Unreviewed Safety Question (USQ) Related to Waste Surface Level Changes in Tank 241-SY-101*. Letter 98-WSD-070 to HJ Hatch (FDH). DOE-RL, Richland, WA.

(d) Boston HL. November 30, 2000. *Approval of Authorization Basis (AB) Amendment for Closure of Tank 241-SY-101 Waste Surface Change Unreviewed Safety Question (USQ) and Removal of Requirement to Operate Tank Mixer Pump*. Letter 00-SHD-124 to MP DeLozier, CHG, Richland, WA.

The DNFSB accepted implementation plans (DOE/RL 1996) that focused characterization efforts on understanding safety-related phenomena to expedite resolution of waste tank safety issues. The plan included 12 deliverables associated with the Flammable Gas Safety Issue. All items, including the ones for flammable gas, were completed and the DNSFB closed 93-5 in November 1999 (Conway 1999).^(a) Satisfactory resolution of DNFSB 93-5 was basic to resolution of the Flammable Gas Safety Issue.

1.1.5 Flammable Gas Safety Issue Resolution

The report documenting all the data and analyses required to close the Flammable Gas Safety Issue was submitted in July 2001 (Johnson et al. 2001). Resolution of the Safety Issue was based on showing that the tanks and tank farm operations were within an approved safety envelope, that waste conditions were properly analyzed and mitigated as necessary, and that loss of primary or secondary containment of waste could not occur.

The closure of all open flammable gas USQs satisfied the requirement that the tanks be operated within their approved Authorization Basis safety envelope. The report discussed the characterization and evaluation of SY-101 and the remediation of the tank such that it could return to normal service. Evaluations for the rest of the DSTs and SSTs were also summarized from prior reports (Hedengren et al. 2000, 2001). In response to this information, DOE Headquarters closed the Flammable Gas Safety Issue on August 17, 2001.

1.2 Local and Global Waste Disturbances

If waste that contains gas bubbles is disturbed, some gas will be released from the volume that is actually disrupted. Whether a disturbance is local or global depends on the size of the disturbance relative to the waste volume. For example, penetration of the waste by a 3-inch-diameter core drill string is a local disturbance, while mixer pump operation, which eventually disturbs the entire waste volume, is global. The history of observations of gas releases during local and global disturbances is summarized below.

1.2.1 Gas Releases Observed During Local Disturbances

A 1996 review of 77 waste-intrusive activities in 47 separate SSTs showed three probable releases of gas associated with the waste disturbance detected with headspace gas monitoring equipment. The study then examined gas releases from an additional 61 core-sampling events and four additional liquid observation well (LOW) installations. These measurements were made using a hand-held combustible gas meter at the riser level before the drill string was flushed and after the sample was removed. Only three samples showed a measurable rise in hydrogen concentration in SSTs (Hedengren et al. 2001).

Additionally, the database of reportable events was reviewed for the time since extensive monitoring controls were implemented on the tanks (about five years). The only instances

(a) Conway JT. 1999. Letter (no title) to B. Richardson, U.S. Department of Energy, November 15, 1999, Defense Nuclear Facilities Safety Board, Washington, D.C.

identified in SSTs where monitored flammable gas levels were >25% of the LFL were in sealed risers or other sealed equipment that was left in a condition where it penetrated the waste for an extended period of time. The gas levels were determined to have resulted from a buildup of gases that were not able to escape or to mix with headspace gases. There were no reports identified of >25% of the LFL in dome or ex-tank regions during waste-intrusive events during this time. While the specific number of waste-intrusive events has not been calculated, it is estimated that the number could easily be in the many hundreds.

The record is similar for the DSTs. During the 1990s, most of the DSTs were sampled with full-length core samplers. The core sampling consisted of penetrating the waste surface crust with a water lance followed by inserting the ~3-inch-diameter core sampler slowly into the waste until the tank bottom was reached. At least two cores were taken in each sampling event. No significant gas release (i.e., a surface level drop of 0.1 inch or greater or measured hydrogen concentration greater than 1,000 ppm hydrogen) was observed for any of these events. Video monitoring of some of the sample events did show bubbling around the lance or sampler shaft, but insufficient gas was released to be detectable even by the headspace gas monitors.

Each DST on the Flammable Gas Watch List was also investigated using the ball rheometer and the void fraction instrument (VFI). The ball rheometer is a 4-inch-diameter, 16-lb ball attached to a cable. The ball is lowered and raised through the waste at varying speeds. The VFI has a 3-ft arm that allows it to sweep an area 6-ft in diameter around its support shaft. The VFI is slowly lowered and raised through the waste taking measurements about every six inches. Each tank was examined through at least two risers. Measurable gas releases were observed only during one deployment of the ball rheometer and VFI in Tank AW-101 when the hydrogen concentration reached 4,600 ppm. The release did not affect operations. Because gas release activity in this tank was particularly active during this period, the small disturbance probably triggered an event that would have occurred spontaneously at about the same time (Hedengren et al. 2000).

1.2.2 Gas Releases Observed During Global Waste Disturbances

Global waste disturbances have been mainly connected with retrieval operations and have included saltwell pumping, mixing, transfers, dissolution, and sluicing. This section summarizes some of the activities and the associated gas releases that might indicate the potential behavior of similar future operations.

Saltwell pumping has been performed for many years in a large number of SSTs, and a considerable experience base has been accumulated for gas releases. Saltwell pumping drains supernatant and as much of the interstitial liquid as is practical with a jet pump installed in a central well. The process usually requires about a year to complete. The maximum hydrogen concentration that has been observed in these passively ventilated tanks during saltwell pumping is 7,200 ppm, less than 25% of the LFL but greater than the action level of 6,250 ppm. Typically, the headspace is monitored periodically and pumping halted if hydrogen levels approach the action level. The fraction of stored gas released during saltwell pumping has ranged from 30 to 50%, implying that over 50% of the gas remains in the waste after pumping (Hedengren et al. 2001).

Jet pump mixing of a tank with a large retained gas volume has been done only in SY-101. This tank was degassed by mixing the waste with a single 150-hp jet pump during October through December 1993 (Allemann et al. 1994). The headspace hydrogen concentration remained at less than 1,000 ppm throughout and exceeded 500 ppm on only three occasions in the initial stages of mixing. Repeated mixer pump runs performed on an average of three times per week for 25 minutes created a relatively uniform slurry with 1 to 2 m of sediment settling out between pump runs. The more recent mixer pump tests in Tank AZ-101 were not representative of a significant degassing operation. Tank AZ-101 had a relatively thin sediment layer and low retained gas volume, and the mixer pump jets were placed just above the sediment.

While mixing in SY-101 prevented gas accumulation in the sediment layer, which had been producing large buoyant displacement gas release events (BDGReS), it caused or allowed the slow growth of a floating crust layer. This caused an accelerating level rise that began to challenge the waste level limits in 1998–1999. The 3-m-thick crust contained as much gas or more than was stored in the waste prior to mixing. The surface level rise issue in SY-101 was remediated by a large-volume dilution and transfer that dissolved the thick crust layer and most of the soluble sodium salts in the tank (Mahoney et al. 2000). Dissolution-induced gas releases were largest when the crust was dissolved in January 2000. The dissolution process lasted several days and raised the hydrogen concentration as high as 3,000 ppm at one point. The larger gas releases were apparently caused by a breakup of large sections of crust, which liberated large volumes of “bubble slurry” (slurry that is approximately 50% gas) that had accumulated beneath it.

Sluicing was performed regularly in the 1950s, ‘60s, and ‘70s. The only recent example that was sufficiently instrumented to assess gas release was sluicing of sludge from C-106 into AY-102 to resolve the high-heat safety issue. C-106 contained about 1.8 m of sludge containing a 2 to 5% gas volume fraction (Stewart and Chen 1998). Sluicing was performed very gently in this tank because of concern for postulated “steam bumps” resulting from reducing the hydrostatic head on sludge near the local boiling point. Each sluicing batch removed several inches to a foot of sludge and resulted in an increase in headspace hydrogen concentration from a baseline of 20–30 ppm to 400–500 ppm. Hydrogen was also liberated in the receiver tank, AY-102, usually increasing to about 100 ppm during sluicing (Cuta et al. 2000).

The saltcake dissolution proof-of-concept demonstration in U-107 was performed between December 2002 and February 2003. Restrictions limited the extent of the demonstration, and failure of the hydrogen sensor precluded gas concentration results (Baide et al. 2003). Modified sluicing retrieval is currently being employed to retrieve Tank S-112, and modified sluicing has been initiated in S-102. Limited gas concentration data is available from S-112. Based on twice-daily combustible gas meter readings, a maximum of 4.4% LFL (roughly equivalent to 1,800 ppm hydrogen) occurred during retrieval from August 20, 2003, to December 8, 2003.

Numerous waste transfer or decanting operations have been conducted in the DSTs. No BDGReS have been documented for these operations. In most instances, however, the only data available is waste surface level, which may not accurately reflect BDGRE behavior (Hedengren et al. 2000).

It is apparent that, even though some of these global waste disturbances were relatively rapid and severe, the peak hydrogen concentrations remained far below the LFL in all cases. The hydrogen concentration was more dependent on the ventilation rate than on the waste disturbance. The foregoing history does not point to an undue or unique hazard potential for future globally disturbing retrieval operations.

1.3 Current and Planned Operations in Hanford Tanks

Retrieval methods include saltcake dissolution, modified sluicing, supernatant decant, vacuum retrieval, which uses an articulating mast system and, if necessary, the Mobile Retrieval System (MRS) or “crawler,” and others. In addition, other globally disturbing operations such as transfers and mixing have been performed historically and are being brought to the same level of analysis as the operations now under scrutiny. This section briefly summarizes the nine globally waste-disturbing activities that are described and evaluated thoroughly in Section 3. Analyses of gas releases during modified sluicing are presented in Huckaby and Wells (2004).

In addition to these activities, other potential globally waste-disturbing activities may be proposed whose potential for gas release will need to be evaluated. However, the associated global waste disturbances are not sufficiently well defined to be analyzed at this time. Finally, major earthquakes also represent a globally waste-disturbing activity that can induce large gas releases (Reid and Deibler 1997). However, these events cannot be planned or controlled and are not considered in this report.

1.3.1 Waste Removal

Fluid waste (liquid or liquid-solid slurry) is removed from a DST by pumping. The pumping action itself is assumed not to cause a significant waste disturbance. Transferring waste out of a tank disturbs the waste globally by reducing the hydrostatic pressure in the sediment. The resulting expansion of retained gas bubbles can cause sections of the sediment layer to become buoyant and release a large fraction of its gas in a BDGRE. At the same time, the first BDGRE can also suspend sediment that increases the bulk supernatant density. This reduces the gas fraction that is required to make the sediment buoyant (the neutral buoyancy gas fraction) and can thereby induce secondary BDGREs. Liquid removal which uncovers sediment, either due to an uneven sediment surface or the pump inlet being located in an excavated area or cavity in the sediment, may induce gas release by drainage of interstitial liquid. Waste removal includes the proposed supernatant decant operation, which has been analyzed in detail (Wells et al. 2002).

1.3.2 Waste Addition

Fluid waste (liquid waste or liquid-solid slurry) is transferred into a tank such that the actual introduction of waste causes only a local disturbance. The introduction of dilute waste may dissolve soluble solids with subsequent gas release. BDGREs may also be triggered by sediment dissolution. Waste addition also increases the hydrostatic pressure on the entire waste volume and potentially changes the density of the liquid layer. Increasing hydrostatic pressure, which compresses retained gas, or adding waste that is less dense than the existing supernatant, which

increases the neutral buoyancy gas fraction, both act to prevent gas release from BDGReS during the addition. However, adding fluid that is denser than the supernatant reduces the sediment neutral buoyancy gas fraction, which can induce BDGReS. The first buoyant displacement can suspend additional sediment, further reducing the neutral buoyancy gas fraction to induce additional gas releases.

1.3.3 Saltwell Pumping

Saltwell pumping is designed to remove supernatant and drainable interstitial liquid from SSTs to reduce the potential for releases of waste into the soil if a tank leaks. Saltwell pumping at Hanford as part of Interim Stabilization has been completed. Saltwell pumping is accomplished with a jet pump installed in a centrally located saltwell screen. The global waste disturbance of saltwell pumping comprises the removal of liquid itself and the resulting decrease in hydrostatic pressure in the liquid.

In saltcake, the liquid drains from the pores between the particles. This removes the buoyant force on solid particles and increases the lithostatic load in the waste column. As described in Section 2, this changes the retained gas bubble configuration to a form with different release behavior. Retained gas is released from the draining saltcake at a rate roughly proportional to the brine-drainage rate.^(a) Dissolved ammonia is also released from the moist, newly unsaturated waste by evaporation. In the history of saltwell pumping, monitored hydrogen concentrations have not exceeded 1 vol% (Hedengren et al. 2001).

In contrast to saltcake, sludge is observed to drain by consolidation of the matrix rather than by flow of liquid through a fixed porous structure, as in coarser saltcake, and is thus always effectively liquid-saturated. This is evidenced by theory and experiment (Simmons 1996). Thus the only mechanism for gas release during saltwell pumping of sludge tanks might be bubble percolation resulting from decreased hydrostatic pressure. Little enhancement to ammonia evaporation would be expected. All references to gas release behavior during saltwell pumping in the balance of this report refer to saltcake.

1.3.4 Saltcake Dissolution

Saltcake dissolution is a method of waste retrieval from SSTs containing a large fraction of soluble waste (Estey et al. 2001). For SST retrieval, water is sprayed on the waste to dissolve the soluble salts, and the resulting brine is pumped out of the tank at about the same average rate that it is created. The global waste disturbance of saltcake dissolution is the destruction of the sediment microstructure by dissolution of solid particles. Any gas bubbles retained in the region of the waste in which soluble solids are dissolved are assumed to be released. This activity presently only applies to SSTs.

(a) Huckaby JL, LA Mahoney, and ER Siciliano. 2001. *Waste Gas Releases Associated with Recent Saltwell Pumping of Single-Shell Tanks*. Letter report TWS01.036, Pacific Northwest National Laboratory, Richland, WA.

1.3.5 Water Addition

Additions of relatively small volumes of water are sometimes necessary to flush transfer lines, lance instruments into the waste, decontaminate hardware removed from the waste, or install retrieval systems. Large water additions may be used during retrieval operations. The global waste disturbance potentially occurs as an increase in hydrostatic pressure and by dissolution of soluble solids, which is the primary gas release mechanism. BDGReS may also be triggered by sediment dissolution.

1.3.6 Mixer Pump Operation

DOE plans to install mixer pumps as part of DST waste retrieval and staging to the vitrification plant. A mixer pump initially degassed the SY-101 waste in 1993, and mixer pumps were tested in AZ-101 in 2001. Mixer pump jets globally disturb the waste by hydraulic erosion and suspension of previously settled sediment. Gas release during mixer pump operation is mainly caused by disruption of the waste structure that retains the gas bubbles. However, a major secondary release mechanism is the increase in effective liquid density by suspension of solid particles. This reduces the neutral buoyancy gas fraction, which may induce BDGReS. Wells et al. (2002) performed detailed analyses of these effects for four DSTs.

1.3.7 Airlift Circulator Operation

Airlift circulators (ALCs) were designed into the construction of the AY and AZ DST farms as well as in Tanks AN-107 and AW-102 and SSTs in several farms. ALCs mix the waste by introducing a stream of air bubbles into large cylindrical tubes that extend from near the tank bottom well up into the supernatant. The bubble stream reduces the average density inside the large tubes, causing a strong upward flow. As many as 22 ALCs may be installed in a tank. The global disturbance by ALC operation is similar to mixing except that the hydraulic action of a high-velocity jet is absent. The main gas release results from mobilizing the sediment around the circulator tubes. As with mixer pump operation, the suspension of solids by the ALCs could also induce BDGReS by reducing the neutral buoyancy gas fraction.

1.3.8 Chemical Addition

Concentrated sodium hydroxide and possibly sodium nitrite solutions need to be added to DSTs occasionally to keep the waste chemistry within corrosion control limits. Each addition is typically on the order of 50,000 gallons of solution with a specific gravity as high as 1.5 (Fort 2001). The global disturbances from chemical addition may include possible solid precipitation, solid dissolution, and, as with waste addition, relatively small increase in the hydrostatic pressure. However, the high density of the solution also increases the bulk density of the supernatant. The corresponding decrease in the neutral buoyancy gas fraction could induce a buoyant displacement in a tank that already experiences them.

1.3.9 Natural Evaporation

Evaporation of water from the waste is included for completeness. It occurs naturally in all tanks but is of potential concern only in tanks with dilute waste and a high heat loading. Evaporation has been insignificant in concentrated saltcake tanks. Evaporation of water increases the concentration of dissolved solids in the supernatant and raises its density. Carried to an extreme, the rising supernatant density could theoretically reduce the neutral buoyancy gas fraction sufficiently to allow BDGReS. However, the increasing concentration of the liquid eventually suppresses evaporation, so the process is self-limiting.

2.0 Physics of Gas Generation, Retention, and Release

Gas is generated in the waste by radiolysis of water, thermal and radiolytic decomposition of organic complexants, and corrosion of the steel tank walls. The gases consist mainly of hydrogen, nitrous oxide, and nitrogen with small amounts of ammonia, methane, and other hydrocarbons that are found in the tank headspaces and stored as bubbles in the liquid and solid wastes (Johnson et al. 1997; Mahoney et al. 1999).

The gas generation rate is an important parameter in determining whether a tank can retain sufficient gas to make the sediment buoyant. The steady-state hydrogen generation rate must also be considered in assessing 1) whether a tank's ventilation rate is sufficient to keep the headspace below the lower flammability limit (LFL) or, 2) if active ventilation fails, the time available to repair the system before the LFL is reached. Gas generation theory and modeling are summarized in Section 2.1.

Gas molecules are generated in the liquid. Because the gases, except ammonia, are not very soluble in concentrated salt solutions, most of the gas comes out of solution as bubbles. Ammonia is very soluble and remains mostly in solution. Though dissolved ammonia can be released by evaporation from a stirred liquid surface, it is not very flammable and is more a toxicological than a flammability hazard. Bubbles containing flammable gases are the most important mode of gas retention and release concerning flammability. Bubbles are retained only in those regions of the sediment that are otherwise saturated with liquid. The configuration and amount of gas retained depends mainly on the properties of the sediment, as described in Section 2.2.

The mechanisms of gas release and the behavior of a specific tank depend on the waste type and configuration. Potentially hazardous spontaneous releases are possible only in tanks where a deep layer of liquid overlies a deep sediment layer. These events (BDGReS) occur when portions of the sediment accumulate enough gas to become buoyant with respect to the liquid above. Models have been developed to determine whether a specific tank will exhibit these releases and whether they are likely to be induced by waste disturbing activities. Gas release behavior and the models currently used to evaluate it are described in Section 2.3.

Finally, a methodology was developed to classify tanks by "waste groups" based on the accumulated knowledge of gas generation, retention, and release behavior (Barker and Hedengren 2004). This methodology provides a sound quantitative basis with which to apply controls to tanks based on the actual flammability hazard they present. The waste group definitions and criteria are summarized in Section 2.4.

2.1 Gas Generation

Stock (2000) reviewed the work that has been done during the past decade at universities, national laboratories, and the Hanford Site to establish the chemical origins of the gases generated in Hanford waste. Section 2.1.1 briefly summarizes Stock's review. An empirical

model was developed by Hu (2000) based on this theory and comparisons with field observations of background hydrogen release rates. This model is summarized in Section 2.1.2.

2.1.1 Gas Generation Theory

Hydrogen is formed in Hanford Site waste in three distinct ways: first, by the radiolysis of water; second, during the corrosion of the steel tank walls; third, during the cascade of radiolytic and chemical reactions involved in the decomposition of organic compounds in the waste.

Organic complexants, organic phosphate esters, and organic hydrocarbons were used at Hanford during separations and other operations. The complex series of degradation reactions of the original compounds has created a broad array of fragmented and oxidized organic compounds. These compounds are also degraded so that the wastes now contain hundreds of different organic compounds. Eventually, slow aging reactions will convert organic carbon into inorganic carbonate ions. A product of many of these reactions is gas, such as hydrogen and other gases found in the tank waste.

Only a small portion of the hydrogen that is formed in the waste is produced through direct radiolysis of the organic compounds because the concentrations (electron density) of the organic compounds are very low compared with water and inorganic salts. Rather, most of the organic-derived hydrogen and other gases are formed in the reactions of the reactive organic intermediates in the later stages of the chemistry.

Methane and other volatile organic compounds are formed via the degradation of organic complexants, extractants, and solvents used in Hanford Site separation processes and in the waste tanks. Production rates and competing reactions (e.g., reactions that degrade the organics to nonvolatile species) are such that the volatile organic compounds are minor waste gases in all but a few passively ventilated SSTs and do not represent a flammability hazard in any tank.

Ammonia arises in part from the oxidation reactions of the nitrogen-containing complexants. The reaction sequences that degrade EDTA and HEDTA lead to the formation of a molecule with a primary amino group and eventually to ammonia. The radiochemical reactions of glycine, which is a common intermediate in the degradation of these complexants, provide ammonia as one of the reaction products. Ammonia also is formed by the hydrolysis of nitriles and amides that are produced during the oxidation reactions of other organic compounds in the waste.

Laboratory investigations indicate that organic compounds are intimately involved in the formation of nitrogen, nitrous oxide, and additional ammonia. Nitric oxide and nitrogen dioxide are important reagents in reaction cascades that lead to the oxidation of the organic compounds and simultaneously to the reduction of nitrate or nitrite ions. Nitrous oxide is also formed by reaction pathways that involve the nitrosyl anion. Studies have demonstrated that these gas generation rates are highly sensitive to temperature and depend on the concentrations of the radioactive isotopes, principally strontium and cesium; organic compounds; and inorganic reagents such as aluminate, nitrite, nitrate, and hydroxide ion.

2.1.2 Hydrogen Generation Model

Hu (2004a) presents a hydrogen generation rate model based on a large body of gas generation and tank waste characterization data. This model, summarized below, is used by Barker and Hedengren (2004) to determine a tank's waste group. The rate equations, which are formulated as a function of physical and chemical properties of tank waste, are used to estimate the hydrogen generation rate of current waste content as well as newly mixed waste of known waste properties. The model uses a set of semi-empirical rate equations to simulate the hydrogen generation mechanism of thermal chemical reactions, radiolysis of water and organic components, and corrosion processes. Hu (2004a) specified the total hydrogen generation rate (HGR) in units of hydrogen yield per kilogram total waste per day as a sum of the thermolysis, radiolysis and corrosion rates as shown in Eq. (2.1):

$$HGR = HGR_{thm} + HGR_{rad} + HGR_{corr} \quad (2.1)$$

where $HGR_{thm} = a_{thm} \times (r_f \times [TOC]) \times [Al^{+3}]^{0.4} \times L_f \times \exp^{(-E_{thm}/RT)}$ (2.2)

$$HGR_{rad} = [(G_{(H_2)}_{H_2O}^{\beta/\gamma} + G_{(H_2)}_{ORG}^{\beta/\gamma}) \times H_{Load}^{\beta/\gamma} + (G_{(H_2)}_{H_2O}^{\alpha} + G_{(H_2)}_{ORG}^{\alpha}) \times H_{Load}^{\alpha}] \times L_f \times CF_1 \quad (2.3)$$

$$HGR_{corr} = R_{corr} \times E_{H_2} \times A_{wetted} / M_{tank} \times CF_2 \quad (2.4)$$

with $G_{(H_2)}_{ORG}^{\beta/\gamma} = a_{rad} \times \exp^{(-E_{rad}/RT)} \times (r_f \times [TOC])$ (2.5a)

$$G_{(H_2)}_{ORG}^{\alpha} = 0.5 \times a_{rad} \times \exp^{(-E_{rad}/RT)} \times (r_f \times [TOC]) \quad (2.5b)$$

$$G_{(H_2)}_{H_2O}^{\beta/\gamma} = \frac{0.32}{1 + 2.4[NO_3^-] + 0.62[NO_2^-] + 0.31[Na^+]_{ex}^2} + \frac{0.13}{1 + 139[NO_3^-] + 54[NO_2^-]} \quad (2.6a)$$

$$G_{(H_2)}_{H_2O}^{\alpha} = \frac{1.05}{1 + 2.4[NO_3^-] + 0.63[NO_2^-]} + \frac{0.35}{1 + 3900[NO_3^-] + 1400[NO_2^-]} \quad (2.6b)$$

and:

E_{thm} = 89.6 kJ/mole, the activation energy for the thermal reaction

a_{thm} = 3.94 E+09 mole/kg-day, pre-exponential factor of the thermal rate

E_{rad} = 48.8 kJ/mole, activation energy in organic radiolysis, G

a_{rad} = 1.11 E+07 H₂/100 eV, the pre-exponential term in organic radiolysis, G

r_f = 0.6 for DSTs and 0.3 for SSTs (unitless), the total organic carbon reactivity coefficient

Note: The $r_f = 0.3$ for SSTs is an average value from the tanks tested. If tank has a high fraction of energetic organic compounds, the r_f can be adjusted to as high as 0.6.

R = 8.314 J/mole/K, gas constant

R_{corr} = 6.0 E-08 for DSTs and 2.4 E-07 for SSTs (ft³/min/ft²), corrosion coefficient

E_{H_2} = the hydrogen generation efficiency coefficient from corrosion is 20% if $[NO_3^-]$, $[NO_2^-]$, and $[OH^-] > 0.1$, otherwise it is 50%.

$[TOC]$ = total organic carbon concentration in the liquid waste (wt%)

Note: Insoluble energetic organic compounds (excluding oxalate) in the solid layer should be considered case by case when data is available

$[Al^{+3}]$ = aluminum concentration in liquid waste (wt%)

$[NO_3^-]$	=	nitrate ion concentration in the liquid waste (moles/L)
$[NO_2^-]$	=	nitrite ion concentration in the liquid waste (moles/L)
$[Na^+]_{ex}$	=	concentration of sodium minus nitrate and nitrite concentration in liquid waste (moles/L)
$H_{load}^{\beta/\gamma}$	=	total heat load of the tank from beta/gamma (watt/kg)
H_{load}^{α}	=	total heat load of the tank from total alpha (watt/kg)
L_f	=	liquid weight fraction in the waste (unitless)
T	=	temperature of waste (K)
A_{wetted}	=	area of steel exposed to moisture-containing waste (ft^2)
M_{tank}	=	total mass in the waste (kg)
CF_1	=	conversion factor from $(H_2/100 \text{ eV})(\text{watts/kg})$ to (mole/kg-day)
CF_2	=	conversion factor from $(\text{m}^3/\text{kg-min})$ to (mole/kg-day).

The analysis of all available gas generation data (Hu 2004a) shows that the temperature-dependent reaction, HGR_{thm} , in Eq. (2.2) follows Arrhenius behavior, which can result from a multistep degradation of organic compounds initiated by radiolysis and followed by thermal reactions. The portion of hydrogen generated by radiolysis, HGR_{rad} , Eq. (2.3), depends on the radiation dose. The hydrogen generated by corrosion, HGR_{corr} , Eq. (2.4), is a function of the interior surface area of the tank exposed to moisture-containing waste. The G-value for hydrogen generation from radiolysis of organic compounds, G_{org} , in Eq. (2.5a) from beta/gamma and Eq. (2.5b) from total alpha is also temperature dependent and follows Arrhenius behavior. The water radiolysis rate is temperature independent, with the G-value, G_{H2O} , in Eq. (2.6a) from beta/gamma and Eq. (2.6b) from total alpha reduced by scavenging by nitrate and nitrite ions.

The numerical values in the equations were established by the analysis of gas generation kinetic data from waste samples with the aid of tank field surveillance data and tank waste characterization data. The reactivity coefficient, r_f , was used to adjust for differences in reactivity of TOC among tanks. A detailed description of this model is given in Hu (2004a).

Hu (2004a) compared the predictions of the model with the observed hydrogen gas release rates in 28 tanks for which enough data were available to determine a release rate. Comparison of the calculated generation rates and the observed release rates for the 28 tanks indicated that the calculated generation rates generally were within a factor of 2 to 3 of the field observations. Hu (2004a) also found that the calculated rates of gas generation by radiolysis and thermolysis in moles of hydrogen per kg of waste per day spanned a large range for these 28 tanks in accord with the variations in the chemical and physical properties of the wastes.

2.2 Gas Bubble Retention Mechanisms

The principal mechanisms of gas bubble retention can be grouped into three categories: bubbles retained by direct attachment to particles (e.g., armored bubbles, attached bubbles, agglomerates), bubbles trapped between particles by capillary forces, and bubbles held within the waste by its strength. In sediment layers, bubble retention is dominated by waste strength and capillary forces, though it may be limited by the bulk buoyancy of the sediment. Pore-filling bubbles are held in the interstitial spaces or pores between solid particles by capillary forces.

Bubbles can also push or displace the solid particles apart to form more or less round bubbles that are trapped by the strength of the surrounding liquid-solid matrix.

Gas release occurs when the bubble retention mechanisms fail. Pore-filling bubbles are not a flammability concern because they move slowly through the porous media. Individual particle-displacing bubbles containing enough gas to present a flammability concern cannot exist, but small particle-displacing bubbles can disengage from the waste relatively rapidly and produce measurable gas releases in the aggregate.

This section explains the differences and limitations of the two major types of bubbles. Gas release behavior is discussed in Section 2.3.

2.2.1 Bubbles Retained by Capillary Forces

Bubbles can fill the interstitial spaces or pores between solid particles when the lithostatic load is sufficient to hold the particles in contact against the force of the bubble's internal pressure that acts to push them apart. The bubble pressure is controlled by capillary force in the throats between pores—the narrower the throat, the higher the pressure. This retention mechanism requires either relatively large particles whose larger pore throats reduce the internal bubble pressure or a deep waste column that increases the lithostatic load—or both. These bubbles assume an irregular, dendritic shape conforming to the passages between the particles. When the bubble's internal pressure is sufficient to overcome the capillary forces, it may still be restrained by the yield strength of the waste, as described in Section 2.2.2.

Whether a bubble is held by yield strength or capillary force depends on a Bond number criterion developed by Gauglitz et al. (1994, 1995, 1996). This dimensionless group is the sum of two parameters: a ratio of gravitational force to surface tension force and a ratio of waste strength force to surface tension force. If the Bond number exceeds unity, a bubble exists in the pore-filling configuration held by capillary forces between particles. The Bond number is expressed as

$$N_{Bo} = \frac{\Delta\rho g H_S D_p}{4\sigma} + \frac{\tau_y D_p}{4\sigma} A \quad (2.7)$$

where

- H_S = the height of the lithostatic column above the bubble
- D_p = the mean pore throat diameter through which a bubble must pass to escape retention; assumed to be represented by the particle diameter
- $\Delta\rho$ = the difference between solid particle and liquid density
- σ = the surface tension
- τ_y = the bulk sediment yield stress
- A = parameter related to how the yield stress resists bubble expansion; it was estimated at 2.8 by Gauglitz et al. (1995)
- g = gravitational constant.

The upper limit on the vertical size of a pore-filling bubble can be derived from the balance of capillary forces and the hydrostatic pressure difference between the top and bottom of the bubble (Stewart et al. 1996). As the bubble grows, the radii of curvature of the liquid-gas interfaces in the pore throats continually adjust in response to the uniform and increasing gas pressure inside the bubble. The bubble internal pressure eventually exceeds the restraining force of surface tension and liquid hydrostatic head in throats on the upper surface of the bubble. This allows gas to push out of the top of the bubble, reducing the pressure to allow liquid to flow into the bottom of the bubble. The bubble thus moves upward until a new equilibrium is established. The gas volume fraction at which this motion occurs is called the percolation threshold. The maximum height, Δh , that a pore-filling bubble can attain before percolating is expressed as

$$\Delta h = 4\sigma/\rho_L g D_p \quad (2.8)$$

where ρ_L is the liquid density.

Pore-filling bubbles, though they may represent a considerable gas volume, are not a flammability concern because there is no mechanism for large numbers of pore-filling bubbles to be released rapidly. For gas in a pore-filling bubble to move, liquid must be displaced from the pores the bubble enters and flow into the pores the bubble vacates. Flow of liquid through the porous media of Hanford waste is generally a very slow process. Also, before a pore-filling bubble can be released into a tank's headspace, it must rise into a decreasing lithostatic load and eventually become a particle-displacing bubble, whose characteristics are described below.

2.2.2 Gas Retained by Waste Strength

The shape of a particle-displacing bubble held by the strength of the waste (as determined by the Bond number in Eq. (2.7)) is determined by the relative effects of surface tension and waste strength. With relatively weak waste or small bubbles, surface tension pulls bubbles into an approximately spherical shape. If the effect of waste strength is greater than surface tension force, the bubble grows into the weakest area of the waste surrounding it and assumes an irregular, dendritic shape. Particle-displacing bubbles that are not dendritic are called "round" bubbles, even though they may be ellipsoidal or similarly distorted from a truly spherical shape. A criterion for bubble shape is derived from simple scaling of the relative importance of strength to surface tension. The maximum diameter of round bubbles before they begin to assume a dendritic shape is given by

$$D_b < \frac{\sigma}{\tau_y} \quad (2.9)$$

The ability of the material to resist the bubble's buoyancy limits the size of the bubble. A particle-displacing bubble can grow vertically only until the buoyant force exerted by the bubble exceeds the ability of the waste to hold it in place. A criterion for incipient motion of a solid sphere immersed in a Bingham fluid can be derived in terms of a critical-gravity yield number, Y_G . The resulting limiting diameter for upward motion of a bubble is expressed as

$$D_b < \frac{\tau_y}{\rho_s g Y_G} \quad (2.10)$$

where ρ_s is the degassed sediment density. The number Y_G is approximately 0.2 for motion of a sphere in an unsheared medium. Using typical waste properties, Eq. (2.9) and (2.10) predict that the maximum diameter of a round bubble is about 0.5 to 1.0 cm. This size is consistent with observations in waste and simulants (Rassat et al. 1997, 1998, Gauglitz et al. 1996). The important conclusion of this analysis is that a very large round bubble that could contain a hazardous amount of gas simply cannot exist.

Eq. (2.10) also gives the limiting height of dendritic particle-displacing bubbles. When they grow to the limiting size, the increased hydrostatic pressure difference pinches off the bottom of the bubble and pushes the top of the bubble upward in a manner similar to percolation of pore-filling bubbles. However, motion of particle-displacing bubbles is much faster because it does not require liquid flow through a porous medium. If the waste is moderately strong, the bubbles do not collapse completely, and dendritic bubble networks can form to provide continuous gas release (Gauglitz et al. 1996).

2.3 Gas Release Mechanisms and Behavior

The mechanisms of gas release and the behavior of a specific tank depend on the waste type and configuration. The three main Hanford tank waste types are liquid (no or little solids present), sludge (insoluble solids) and saltcake (solids precipitated from saturated liquid on cooling). Liquid, by definition, is a dilute or concentrated solution of dissolved solids, water, and a small fraction of soluble organics that contains at most a small fraction of suspended solid particles. Sludge consists of insoluble particles that are mostly metal oxides and hydroxides. Saltcake forms by the settling of sodium and aluminum salts precipitated from a cooling saturated solution. Solid particles in saltcake are generally expected to be larger than those in sludge. Some tanks contain a mixture of sludge and saltcake that can be considered a “mixed” waste type. However, the mixed waste appears to behave like saltcake in most cases.

Waste configuration is the arrangement of waste types in layers within a tank. The waste types are distributed within the tanks in only a few different ways, depending mainly on the amount of liquid. In order of decreasing amount of liquid, the main waste configurations are 1) “liquid,” containing almost entirely the liquid waste type; 2) “liquid-over-solid,” where a relatively deep layer of supernatant liquid overlies an equally deep layer of liquid-saturated sediment; 3) “wet solids,” where the sediment is saturated with liquid but there is little or no supernatant; and 4) “pumped,” where interstitial liquid has been removed by saltwell pumping or other means such that its level is well below the waste surface.

Two other waste configurations are also of interest: “crust,” a solid-over-liquid arrangement consisting of a floating layer of gas-bearing solids, and “mixed slurry,” in which most of the solid particles and small gas bubbles are kept in suspension mechanically. A crust exists only with the more concentrated saltcake wastes in a few DSTs as a modification of the liquid-over-solid configuration. The SSTs A-101 and AX-101 were unique in that, based on core sample

evidence, the entire mass of undissolved solids in the tank apparently floated in a very dense liquid layer. However, this cannot be treated as a true “crust.” Mixed slurry existed in SY-101 from 1993 to 2000, when the mixer pump was operating. The characteristics of gas retention and release in each of the four main waste configurations, liquid, liquid-over-solid, wet solids, and pumped, are discussed in more detail below.

2.3.1 Liquid Waste Configuration

Only a few of the DSTs currently contain essentially all liquid. In these tanks, gas generated in solution by radiolysis nucleates to form small bubbles on suspended microparticles or on the tank walls. Once these bubble-particles grow large enough to become buoyant, they rise immediately to the surface and release gas into the headspace to be removed by ventilation. This continuous background release mechanism makes retention of a large volume of gas bubbles physically impossible in this configuration.

2.3.2 Liquid-over-Solid Configuration

Except for DSTs with thick crust layers, gas is retained mainly in the sediment. The liquid-over-solid waste configuration in saltcake DSTs is subject to BDGReS that are the largest and fastest releases known to occur in Hanford tanks and the only releases that are known to have created flammable conditions in a tank’s headspace. Before the mixer pump was installed in SY-101, its buoyant displacements typically released over 5,000 scf of gas over a period of 10 to 20 minutes three times per year. At least two of these releases exceeded 10,000 scf, enough to make the tank headspace just flammable. Five other DSTs (AN-103, AN-104, AN-105, AW-101, and SY-103) have historically exhibited BDGReS that are much smaller, slower, and less frequent (Hedengren et al. 2000). None have reached flammability, and only two releases in AN-105 have exceeded 25% of the LFL.

The yield stress in the sediment of these five DSTs as measured by the ball rheometer increases roughly linearly from zero at the upper surface to 200–300 Pa where the ball rheometer stopped (Hedengren et al. 2000). This strength is sufficient to hold bubbles up to about a centimeter in diameter. The particle size is relatively small so that Eq. (2.7) predicts only particle-displacing bubbles. The material strength is also sufficiently low that the bubbles remain approximately round, according to Eq. (2.9). Measurements made by the void fraction instrument (VFI) and retained gas sampler (RGS) show that gas volume fraction in the sediment layer increases with depth from near zero at the top. Two distinct gas fraction profiles were observed: one approximately parabolic, with a peak at about the midpoint of the sediment layer; and another roughly linear, with the maximum at the bottom. The peak gas fraction ranges from 0.1 to 0.16 with averages from 0.04 to 0.11 (Hedengren et al. 2000).

BDGReS occur in saltcake tanks with a deep layer of supernatant liquid when a portion, or “gob,” of the sediment accumulates enough gas to become buoyant with respect to the liquid above it, breaks away, and rises through the liquid. The stored gas bubbles expand as the gob rises, failing the surrounding material so a portion of the gas can escape into the headspace. After releasing gas until it is no longer buoyant, the gob sinks back to rejoin the sediment. This

process can be described in terms of several criteria based on the waste properties and configuration (Meyer and Stewart 2001).

2.3.2.1 Buoyancy Ratio Criterion

Gas accumulation in the sediment is a balance between gas generation and steady background gas release. The background release is assumed to be a slow migration of bubbles that qualitatively obeys the form of Stokes' Law. Consistent with in situ measurements, the waste viscosity (determining bubble rise velocity) is assumed to increase linearly with waste depth from zero at the top of the layer. The gas volume fraction (also called void fraction) profile can be determined from the equations for conservation of bubble mass and number (Meyer and Wells 2000). Assuming a uniform bubble nucleation rate and gas generation rate, the solution is a parabolic distribution given by

$$\alpha(\eta) = \frac{16}{3} \frac{C_B}{\rho_s} \left(\frac{GT}{P_s} \right)^{\frac{1}{3}} H_s^2 \eta (1 - \eta) \quad (2.11)$$

where ρ_s is the degassed sediment density (measured), T is the average sediment temperature (measured), P_s is the average pressure in the retained gas (calculated from measured waste physical properties), G is the volumetric gas generation rate (estimated from measured retained and released gas concentrations and waste chemical and physical properties), H_s is the depth of the sediment layer (measured), and $\eta = z/H_s$, where z is the distance from the tank bottom. If the average gas fraction in the sediment layer, α_{avg} , predicted by Eq. (2.11) exceeds the neutral buoyancy gas fraction, a BDGRE will occur. The coefficient C_B is adjusted so that the tanks currently exhibiting buoyant displacements are predicted to exceed neutral buoyancy. The neutral buoyancy gas fraction is the gas volume fraction required for the sediment to become buoyant with respect to the liquid above it. Thus, the neutral buoyancy gas fraction is defined by the waste density as follows:

$$\alpha_{NB} = 1 - \frac{\rho_L}{\rho_s} \quad (2.12)$$

where ρ_L is the liquid density (measured). The criterion of Eq. (2.11) and (2.12) is usually expressed as the buoyancy ratio, the average gas fraction obtained by integrating Eq. (2.11) divided by the neutral buoyancy gas fraction, Eq. (2.12). The buoyancy ratio is expressed by

$$BR = \frac{\alpha_{avg}}{\alpha_{NB}} = \frac{C_B}{\rho_s - \rho_L} \left(\frac{GT}{P_s} \right)^{\frac{1}{3}} H_s^2 \quad (2.13)$$

If the buoyancy ratio is unity or greater, BDGReS can be expected.

2.3.2.2 Energy Ratio Criterion

The process of gas release from a gob undergoing buoyant displacement requires that sufficient energy be released to disrupt the waste surrounding the bubbles to allow them to escape as the gob reaches the waste surface. The amount of energy available is directly proportional to the depth of the supernatant through which the gob rises. The amount of work required to yield the gob is directly proportional to the yield stress of the material. In addition, a large fraction of the buoyant energy is dissipated in other processes so the required energy is much greater than that which would just yield the waste.

Stewart et al. (1996) developed an energy model to account for this. The ratio between the buoyant energy, E_b , and the energy required to yield the gas-bearing gob participating in the buoyant displacement, E_y , may be expressed as

$$\frac{E_b}{E_y} = \frac{\alpha_C \rho_L g h}{(1 - \alpha_C) \varepsilon_y \tau_y} \left[\left(1 + \frac{1}{\gamma} \right) \ln(1 + \gamma) - k \right] \quad (2.14)$$

where h is the distance from the top of the participating gob to the top of the liquid layer, τ_y is the yield stress in shear, and ε_y is the strain at failure, which is taken to be unity. The parameters γ , k , and α_C are determined from

$$\gamma = \frac{\rho_L g h}{P_A} \quad (2.15)$$

$$k = \frac{\alpha_{NB}(1 - \alpha_C)}{\alpha_C(1 - \alpha_{NB})} \quad (2.16)$$

$$\alpha_C = \alpha_{NB} + \frac{\beta \tau_y}{\rho_S g H_S} \quad (2.17)$$

where β is ratio of the yield stress in tension to the yield stress in pure shear, which is taken to be unity. For weak waste with τ_y much less than $\rho_S g H_S$, $\alpha_C = \alpha_{NB}$, and Eq. (2.14) can be written as

$$\frac{E_b}{E_y} = \frac{(\rho_S - \rho_L) g h}{\varepsilon_y \tau_y} \left[\left(1 + \frac{1}{\gamma} \right) \ln(1 + \gamma) - 1 \right] \quad (2.18)$$

Based on experimental observations and tank behavior, some gas can be released when the energy ratio exceeds 3, and releases of a large fraction of the stored gas can occur above energy ratios of 5.

2.3.2.3 Other BDGRE Criteria

Several empirical relationships have also been used that correlate groupings of physical parameters with BDGRE behavior. The single waste property found to most closely correlate with BDGRE behavior was tank average specific gravity or density.^(a) A criterion was established for waste transfers (Fowler 1995) based on the weighted mean specific gravity of transferred waste and the waste in the receiver tank. If the average specific gravity exceeded 1.41, BDGRE behavior was considered possible.

Estey and Guthrie (1996) found that the product of sediment depth and supernatant specific gravity gave a clearer separation than the average specific gravity between tanks that exhibited BDGREs and those that did not. If the product of sediment depth (in inches) and supernatant specific gravity exceeded 150 inches, BDGREs were considered possible.

2.3.3 Wet Sediment Configuration

Tanks which consist entirely of sediment that is saturated with liquid but little or no supernatant may retain a significant volume of gas. For the generally expected larger particles of saltcake waste, Eq. (2.7) states that the upper several meters will retain individual round bubbles while the bottom few meters will contain pore-filling bubbles due to the high lithostatic load. With the generally expected smaller particles of sludge, retained gas bubbles will exist as particle-displacing bubbles at any depth. The wet sediment tanks with saltcake waste type have the highest gas fraction and deepest waste, so they account for most of the total retained gas volume.

Unlike the liquid-over-sediment waste configuration, the gas retention characteristics of sludge and saltcake in wet sediment tanks appear to be quite different. The median gas fraction for sludge tanks is only 0.01 but is 0.12 for all saltcake tanks. An average gas fraction exceeding 0.25 has been measured in a saltcake tank; the sludge tanks only reach 0.12. A large group of the sludge tanks exhibit a barely detectable gas fraction (Barker et al. 1999). This difference likely results from the generally lower waste depth, lower gas generation rate per unit volume, and smaller particle size of the sludge.

The characteristics of gas releases from the wet sediment configuration depend to a great degree on how the gas is stored. For example, pore-filling bubbles require liquid to flow through the porous media in order to migrate. This makes gas release a very slow process. Gas release from particle-displacing bubbles may be somewhat faster but is believed to be limited to a small region of a tank. This is confirmed by headspace gas monitoring data that show these tanks typically releasing on the order of 10 to 100 scf of gas over a period of several days. These releases tend to occur when the barometric pressure falls rapidly during storm passage in the late fall and early spring.

(a) Reynolds DA. 1994. *Evaluation of Specific Gravity versus Gas Retention*. Internal memo 7E310-94-024, Westinghouse Hanford Company, Richland, Washington.

2.3.4 Pumped Configuration

In pumped tanks containing saltcake, drainable liquid has been removed and the interstitial liquid level is usually well below the waste level. The portion of the waste above the liquid level exerts a disproportionately high lithostatic load on the waste below. Applying Eq. (2.7) with the density difference set equal to the solid density, only one meter of unsaturated saltcake above the liquid level is required to force the entire column into the pore-filling bubble configuration.

In pumped sludge tanks, liquid removal results in consolidation of the liquid-solid matrix rather than draining of a fixed porous structure (Simmons 1996). Thus, pumping does not create an unsaturated overburden in sludge like it does in saltcake, and the lithostatic load does not change very much. Even with an unsaturated layer, Eq. (2.7) indicates that the expected smaller size of sludge particles prevents formation of pore-filling bubbles. Trapped gas bubbles remain in a particle-displacing configuration in any case and tend not to release during pumping.

Gas fraction information for pumped tanks is available for some sludge tanks (Barker et al. 1999). Only one pumped saltcake tank, BY-109, has gas fraction data (Mahoney 2000). Note that BY-109, while predominantly saltcake (253 kgal), does contain 24 kgal of sludge (Hanlon 2002). The limited data show there is essentially no reduction in gas fraction after pumping a sludge tank consistent with the theory above. The mean gas fraction below the interstitial liquid level in BY-109 measured with the retained gas sampler was 0.09, the lowest of all tanks measured with the RGS.

Gas release behavior during and after saltwell pumping of saltcake tanks is discussed in detail in Section 3.3.2. There is no known mechanism for large spontaneous gas release from a pumped tank.

2.4 Waste Classification

As discussed above, release of retained gas depends on the form and nature of the waste as well as the nature of any waste-disturbing activities. To develop and implement cost-effective flammable gas control strategies, tanks are grouped according to waste form and gas release behavior.

2.4.1 Waste Group Definition

Three “waste groups” are defined based on the understanding of gas retention and release behavior to identify the tanks that pose a flammable gas hazard and to differentiate tanks that are susceptible to BDGReS from those that are not (Barker and Hedengren 2004). Waste group classifications are presented in Barker and Hedengren (2004). The following is a summary of the waste group classifications:

- **Waste Group A:** Includes tanks that have the propensity for spontaneous BDGReS and have sufficient retained gas to exceed the LFL if all of it is released suddenly. BDGReS may also be induced in these tanks by large waste disturbances that suspend sediment in the supernatant or otherwise increase its density.

- **Waste Group B:** Includes tanks that do not exhibit spontaneous BDGReS but have sufficient retained gas to exceed the LFL if all of it were released suddenly. Given the level of retained gas, significant gas releases potentially can be induced by large disturbances in the sediment.
- **Waste Group C:** Includes all DSTs and SSTs that do not have sufficient retained gas to achieve the LFL if all of the retained gas were released suddenly, regardless of whether they exhibit spontaneous BDGReS.

2.4.2 Criteria for Waste Group Classification

The three criteria specified in Barker and Hedengren (2004) are used to place the tanks in the appropriate waste group. The first is a screening that evaluates the potential flammable gas hazard. Tanks with insufficient retained gas in the sediment layer to cause the tank headspace to become flammable if all of it were released at once are considered nonhazardous and are placed in Waste Group C. The sediment gas volume is the product of the waste volume calculated from the wet sediment depth and the average retained gas volume fraction. The gas volume fraction may be determined from measurements, assigned a conservative bounding value, or limited to the neutral buoyancy gas fraction for tanks with liquid-over-sediment waste configuration.

If a tank is shown to retain sufficient gas to make the headspace flammable, it is assigned to Waste Group A if it has the propensity for spontaneous BDGReS or to Waste Group B if it does not. The presence of a sufficient supernatant liquid layer is the first requirement for BDGReS. This requirement is evaluated by calculating the “energy ratio” defined by Eq. (2.18). Tanks with an energy ratio below the threshold value of 3 cannot release gas in a BDGRe and are assigned to Waste Group B.

Tanks exceeding the energy threshold are further evaluated to determine whether buoyancy can be achieved. The buoyancy ratio, Eq. (2.13), is used for this evaluation. If the buoyancy ratio exceeds unity, BDGRe behavior is indicated and the tank is assigned to Waste Group A.

3.0 Gas Releases Induced by Global Waste Disturbances

A global waste disturbance is one that affects all or a large fraction of the total waste volume. Global disturbances can range from a change in waste volume, with coincident change in hydrostatic pressure, to a complete redistribution of the waste mass as in mixer pump operation. The objective of this section is to evaluate the impacts of authorized globally waste-disturbing activities based on the dominant gas release mechanisms and tank waste group assignment described in Section 2.4. The need for gas monitoring, the potential long-term changes in waste group assignment, and increases in the steady-state hydrogen generation after the activity is completed are also included in the evaluation.

The globally waste-disturbing activities covered in this section are summarized in Table 3.1. Generally, the activities are specific to either DSTs or SSTs based on the context of the operation. For example, saltwell pumping is treated separately from waste removal because saltwell pumping is performed only in SSTs. Similar nuances are explained in the detailed discussion below. Besides the waste disturbance itself, any activity that adds material to the tank also decreases the headspace, which exacerbates the consequences of gas releases. Except for natural evaporation, global waste disturbances associated with natural phenomena are not included because they are either of low probability (e.g., severe earthquakes) or low consequence (Johnson et al. 2001).

Table 3.1. Summary of Globally Waste-Disturbing Activities

Activity (tank type)	Description	Waste Disturbance
Waste removal (DST)	Liquid or slurry pumped out	Reduced hydrostatic pressure and supernatant depth, sediment uncovering and interstitial liquid drainage
Waste addition (both)	Liquid or slurry transferred in	Change in supernatant density, increased supernatant and sediment depth and hydrostatic pressure; sediment dissolution
Saltwell pumping (SST)	Supernatant and interstitial liquid pumped out	Interstitial liquid drainage, increased lithostatic load, decreased wet sediment depth
Saltcake dissolution (SST)	Dissolution of soluble solids by water addition with brine removal	Destruction of sediment structure, waste volume reduction, decrease in lithostatic load
Water addition (both)	Water added without brine removal	Solids dissolution, increased hydrostatic head, decreased supernatant density
Mixer pump operation (DST)	High-energy jet mobilizes sediment	Destruction of sediment structure, solids suspended in supernatant
Airlift circulator operation (DST)	Low-energy circulation mobilizes sediment	Destruction of sediment structure, solids suspended in supernatant
Chemical addition (DST)	High-density liquid added to supernatant	Increased hydrostatic pressure, supernatant depth, and density; sediment dissolution
Natural evaporation (DST)	Water loss to atmosphere	Increased supernatant density, decreased supernatant depth

Saltcake dissolution, as discussed herein, pertains to water additions to SSTs where the primary mechanism of waste disturbance is dissolution. Although dissolution also occurs during modified sluicing retrieval of SSTs, the waste is initially mobilized hydraulically. Modified sluicing is not addressed in this document; analyses of gas releases during modified sluicing are presented in Huckaby and Wells (2004).

3.1 Waste Removal

Fluid waste is transferred out of a tank by pumping. The pumping process itself is considered a local disturbance whether the inlet is located in the supernatant or in the sediment. The fluid may all be liquid if the transfer pump inlet is located in the supernatant or “cavity” in the sediment or a slurry if the inlet is in the sediment or the waste is mixed before or during transfer. For supernatant removal which uncovers sediment, either due to an uneven sediment surface or the pump inlet being located in an excavated area or cavity in the sediment, gas release is induced by drainage of interstitial liquid as with saltwell pumping (Section 3.3). In the case of slurry transfer, the primary waste disturbance is the mixing operation, which is covered in Section 3.6. Essentially all of the retained gas would have been released by mixing, nullifying the potential consequences of waste removal. Removal of waste from a tank causes a global disturbance by reducing the hydrostatic pressure on the entire waste volume. The waste configuration is also altered by reducing the depth of supernatant, which also increases the tank headspace.

This method of waste removal is specific to DSTs. It includes the proposed supernatant decant operation, which was analyzed in detail by Wells et al. (2002) and whose results are summarized in Section 3.1.2. Waste removal from SSTs is considered under saltwell pumping in Section 3.3 or saltcake dissolution in Section 3.4.

3.1.1 Mechanisms for Gas Releases During Waste Removal

The primary mechanism of gas release from this activity is the reduction in hydrostatic pressure, as explained below. Gas release may also be caused by drainage of interstitial liquid from uncovered sediment. The expansion of retained gas bubbles resulting from the reduction in hydrostatic pressure can cause a sediment layer to become buoyant or allow the expanded bubbles to disengage from the sediment. If a buoyant displacement occurs, the suspended sediment increases the supernatant bulk density, reducing the neutral buoyancy gas fraction and potentially inducing secondary BDGReS. The increased density also increases the hydrostatic pressure, which decreases the gas fraction and reduces the potential for a BDGRE. However, this effect is much less than of the increase in buoyancy. If the entire sediment layer in AN-105 [waste conditions from Wells et al. (2002)] were suspended in the supernatant, for example, the neutral buoyancy gas fraction would decrease almost 60%, from 0.095 to 0.054, while the gas volume fraction of a gob initially just at neutral buoyancy would only decrease 2%, from 0.095 to 0.093 due to increased hydrostatic pressure.

Peak hydrogen concentrations resulting from BDGReS induced by waste removal are highest if supernatant is removed rather than sediment. If sediment is removed, the rate of hydrostatic pressure reduction on the gas in the undisturbed portion of the sediment layer is slower. This

tends to induce BDGReS later into a larger headspace. The local disturbance of sediment pumping would also release some of the retained gas that would otherwise have been released in BDGReS.

Waste removal from a Group A tank^(a) will induce BDGReS of a magnitude similar to its historical spontaneous releases. A BDGRE can theoretically also be induced by waste removal in a Group B tank if the retained gas volume fraction is high enough to achieve buoyancy by the hydrostatic pressure reduction. The potential for an induced BDGRE in a Group B tank can be evaluated prior to waste removal as follows: The change in sediment retained gas volume fraction, α , with the average sediment gas pressure, P_S , is given by (Wells et al. 2002)

$$\frac{d\alpha}{dP_S} = -\frac{\alpha(1-\alpha)}{P_S} \quad (3.1.1)$$

Integrating Eq. (3.1.1) from initial state 0 to the final state 1 following waste removal yields

$$\alpha_1 = \left[\frac{P_1}{P_0} \left(\frac{1-\alpha_0}{\alpha_0} \right) + 1 \right]^{-1} \quad (3.1.2)$$

We want to know the pressure reduction necessary to raise the initial gas fraction to the neutral buoyancy value, α_{NB} , defined by Eq. (2.12). Substituting α_{NB} for α_1 and solving for the pressure change gives

$$\frac{\Delta P}{P_0} = -\frac{\alpha_{NB} - \alpha_0}{(1 - \alpha_0)\alpha_{NB}} \quad (3.1.3)$$

The initial average sediment gas pressure can be approximated by the pressure at the midpoint of the sediment:

$$P_0 = P_A + g[\rho_L(H_{L0} + H_C) + \rho_S H_S / 2] \quad (3.1.4)$$

where P_A is barometric pressure, ρ_L and ρ_S are the liquid and sediment density, respectively, H_{L0} is initial supernatant depth, H_C is the thickness of the crust if one exists, and H_S is the sediment depth. Assuming only supernatant is removed, the liquid depth changes by ΔH_L and the corresponding pressure change is $\Delta P = \rho_L g \Delta H_L$. Substituting this for ΔP in Eq. (3.1.3) and using Eq. (3.1.4) for P_0 gives an expression for the reduction in supernatant depth necessary to raise the initial gas volume fraction to neutral buoyancy:

$$\Delta H_L = -\left(\frac{P_A}{\rho_L g} + H_{L0} + H_C + \frac{\rho_S H_S}{\rho_L / 2} \right) \frac{\alpha_{NB} - \alpha_0}{(1 - \alpha_0)\alpha_{NB}} \quad (3.1.5)$$

(a) As described in Section 2.4.1, Waste Group A and B tanks store sufficient gas to make the headspace flammable if all of it were suddenly released; Waste Group C tanks do not. Waste Group A tanks exhibit spontaneous BDGReS, while tanks in Waste Group B do not.

If the proposed waste removal is greater than or equal to this depth, a BDGRE can be expected during the operation. Note, however, the increased headspace from a decant could potentially move the tank into Waste Group C prior to occurrence of a BDGRE.

Drainage of interstitial liquid from uncovered sediment can release gas in a similar manner as observed during saltwell pumping in SSTs (Section 3.3). For DSTs, plausible scenarios whereby sediment may be uncovered and subsequently drain include: 1) an uneven sediment surface or 2) the pump inlet is located in an excavated area or cavity in the sediment. The ability of the strength of a particular sediment to support either of these configurations, i.e., the possibility of large-scale slumping or collapse of the sediment during pumping/drainage, is not considered in this analysis.

As a screening evaluation, Waste Group evaluation (as described in Barker and Hedengren 2004) of the tank with no supernatant will determine whether the tank stores sufficient gas to make the headspace flammable if all of it were suddenly released. For a Waste Group C result, no further evaluation is required. For a Waste Group B result, additional evaluations of the releasable gas volume may be performed.

Initially, consideration of the transfer pump inlet location may restrict potential gas release sufficiently so as to not require controls. That is, liquid will not be drained from sediment below the inlet of the transfer pump, thereby restricting potential gas release. Thus, Waste Group evaluation (Barker and Hedengren 2004) of the tank with no supernatant and a sediment depth limited on the bottom by the transfer pump inlet will determine whether the tank stores sufficient gas to make the headspace flammable if all of it were suddenly released. If not, no further evaluation is required.

The total gas release possible for a given transfer pump inlet elevation may be determined from the waste's configuration. As described in Huckaby and Wells (2004), liquid will drain from exposed, un-drained sediment until the interstitial liquid level has equilibrated in a liquid "pool". For the scenarios described above, this "pool" may either be a region or regions of the tank where the sediment is less deep than in other areas, or it may be an excavated area or cavity around the transfer pump. In practice, the volume available for liquid drainage may be difficult to determine. Thus, conservative assumptions are required to determine an allowable liquid transfer volume that will not threaten flammability due to additional gas released by possible subsequent draining.

The total gas release volume accounts for the gas released from the sediment by removal of the transfer volume (liquid removed in excess of any supernatant liquid present) and the potential drainage volume. Sediment surface irregularities and/or excavated areas or cavities are conservatively neglected. The gas release volume at standard conditions, V_{GR} , may be computed from the volume of liquid removed from below the sediment level (transfer volume), V_{LR} , as

$$V_{GR} = V_{LR} \left[\frac{\alpha}{(1 - \alpha)\phi_L} \right] \frac{P_0}{P_A} \quad (3.1.6)$$

where α is the retained gas fraction in the undrained waste and the liquid volume fraction in the de-gassed waste, ϕ_L , is defined by

$$\phi_L = \frac{\rho_{SS} - \rho_S}{\rho_{SS} - \rho_L} \quad (3.1.7)$$

where ρ_{SS} is the dry solid density.

To conservatively account for the potentially unknown waste configuration (waste surface irregularity, excavated area or cavern size, etc.), we assume that the total volume of sediment required to provide the transferred liquid is also available for drainage. That is, the solid volume remaining after the liquid is removed and gas is released is neglected, so that the drained liquid volume, V_{LD} , is equal to $V_{LR}/((1-\alpha)\phi_L)$. Any physical constraints which would preclude these assumptions are ignored. Thus, the gas release volume at standard conditions from the draining the remaining sediment, V_{GD} , may be determined from

$$V_{GD} = V_{LR} \left[\frac{\alpha}{[(1-\alpha)\phi_L]^2} \right] \frac{P_0}{P_A} \quad (3.1.8)$$

The amount liquid that may be transferred is thus limited such that the summation of the transfer and drainage induced gas release volumes, Eq. (3.1.6) and (3.1.8), is less than the minimum gas volume for flammable conditions in the headspace V_{Gmin} determined from the Retained Gas Volume Criterion from Barker and Hedengren (2004) or

$$V_{LR} < V_{Gmin} \left\{ \frac{\alpha}{(1-\alpha)\phi_L} \left[1 + \frac{1}{(1-\alpha)\phi_L} \right] \frac{P_0}{P_A} \right\}^{-1} \quad (3.1.9)$$

If a liquid volume in excess of that provided by Eq. (3.1.9) needs to be transferred, it may be possible to control the hazard by headspace ventilation and transfer rate. Cessation of a transfer will stop gas release due to liquid removal. However, drainage may continue, thereby increasing the headspace gas concentration. Thus, in addition to ensuring that the transfer operation itself will not cause the headspace to reach flammability, the additional gas release from drainage must also be accounted for.

As shown in Section 3.3, the rate of retained gas release is approximately proportional to the interstitial liquid drainage rate. Interstitial liquid may only drain as fast as the sediment is uncovered, or, similarly, as fast as liquid is removed from the tank. The rate of gas release, Q_{GAS} , as defined conservatively by the transfer rate, Q_D , is therefore

$$Q_{GAS} = \frac{Q_D}{(1-\alpha)\phi_L} \frac{P_0}{P_A} \alpha \quad (3.1.10)$$

During the transfer the potential gas release due to draining will be maximized when the available drainable liquid volume is equivalent to the volume into which it can drain. Under the conservative assumptions discussed above, this maximum gas release, $V_{GD\max}$, is given by

$$V_{GD\max} = V_{ST} \left[1 - \frac{1}{1 + \frac{1}{(1-\alpha)\phi_L}} \right] \frac{P_0}{P_A} \alpha \quad (3.1.11)$$

where V_{ST} is the total sediment volume. Gas release due to drainage will be minimized at the initiation of the transfer and as the liquid is completely removed. We conservatively allow for the release of this maximum gas volume due to draining throughout the transfer. Section 3.1.3 discusses the application of Eq. (3.1.10) to determine headspace concentrations as a function of ventilation and transfer rates and to define monitoring frequencies to ensure that the maximum potential drainage release from Eq. (3.1.11) will not threaten the LFL.

3.1.2 Expected Gas Releases During Waste Removal

Extensive modeling studies based on detailed simulations of the supernatant decant process show with a high degree of confidence that tanks that historically exhibited BDGReS would not reach LFL during waste removal. The pre-remediated state of SY-101 was not evaluated; spontaneous BDGReS in SY-101 did exceed the LFL (Hedengren et al. 2000). The simulation results described by Wells et al. (2002) for Tanks AN-103, -104, and -105 and AW-101 are summarized in Table 3.2. These results bound the consequences of supernatant liquid waste removal from Group B and C tanks.

While it is possible for the current Group B DSTs (Barker and Hedengren, 2004) to experience BDGReS during decant if the change in liquid depth is at least that given by Eq. (3.1.5), the result is expected to be inconsequential. Because the sediment in Group B tanks is, by definition, not near buoyancy, an induced BDGRe can occur only near the end of the transfer if at all. In fact, if the buoyancy ratio is 0.5 or less, no BDGReS can be induced by waste removal. The shallow supernatant depth near the end of transfer reduces the gas release volume, and the larger headspace provides maximum dilution. These factors combine to essentially eliminate the possibility that these late gas releases from current Group B tanks would achieve flammability.

It is possible that a Waste Group B DST with a deep supernatant layer and buoyancy ratio approaching unity may be created. In this event, a BDGRe may be triggered relatively early in a decant operation and may not be nullified by release volume and headspace effects. However, the behavior of the BDGRe tanks shows that release size is roughly proportional to the buoyancy ratio. In fact, the maximum expected hydrogen concentration at a buoyancy ratio of 1 is less than 15,000 ppm (Stewart et al. 2005). Thus, in a tank with a buoyancy ratio less than unity, the likelihood of large decant-induced BDGReS is very low. In addition, as stated above, the decant

Table 3.2. Summary of Supernatant Decant Analysis Results

Tank	Run	Quantity	Median	95% CL
AN-105	Base Case (200 gpm)	Peak Hydrogen (ppm)	2,800	8,300
		Fraction Gas Release	0.09	0.22
	30 gpm Decant	Peak Hydrogen (ppm)	2,600	8,000
		Fraction Gas Release	0.09	0.22
	Stop-Start Control	Peak Hydrogen (ppm)	2,800	8,200
		Fraction Gas Release	0.09	0.22
	Water Backfill	Peak Hydrogen (ppm)	23	7,600
		Fraction Gas Release	0.00	0.13
AN-104	Base Case	Peak Hydrogen (ppm)	2,800	7,100
		Fraction Gas Release	0.12	0.27
	30 gpm Decant	Peak Hydrogen (ppm)	2,500	6,500
		Fraction Gas Release	0.12	0.27
	Stop-Start Control	Peak Hydrogen (ppm)	2,800	7,000
		Fraction Gas Release	0.12	0.27
	Water Backfill	Peak Hydrogen (ppm)	34	6,400
		Fraction Gas Release	0.00	0.16
AN-103	Base Case	Peak Hydrogen (ppm)	10,500	21,100
		Fraction Gas Release	0.18	0.33
	30 gpm Decant	Peak Hydrogen (ppm)	8,200	19,100
		Fraction Gas Release	0.18	0.33
	Stop-Start Control	Peak Hydrogen (ppm)	9,000	19,400
		Fraction Gas Release	0.18	0.33
	Water Backfill	Peak Hydrogen (ppm)	1,500	21,600
		Fraction Gas Release	0.01	0.25
AW-101	Base Case	Peak Hydrogen (ppm)	400	1,500
		Fraction Gas Release	0.05	0.17
	30 gpm Decant	Peak Hydrogen (ppm)	400	1,500
		Fraction Gas Release	0.06	0.19
	Stop-Start Control	Not run	--	--
	Water Backfill	Peak Hydrogen (ppm)	27	1,100
		Fraction Gas Release	0.00	0.09

analyses of Wells et al. (2002) that demonstrated that the tanks that historically exhibited BDGReS would not reach LFL during waste removal bound the potential gas releases from Group B and C tanks.

Four cases were run for each of the four tanks studied (Wells et al. 2002). The base case applied a constant 200-gpm decant rate until all of the supernatant was removed. The headspace ventilation rate was 125 scfm in AW-101 and 100 scfm in the other tanks. The sensitivity to decant rate was tested with a second run using a 30-gpm decant rate. The effect of the normal control strategy was tested with a third run by stopping the decant when the headspace hydrogen concentration exceeded the action level of 6,250 ppm and restarting only when the hydrogen concentration fell below 500 ppm. Finally, the efficacy of a proposed mitigation strategy of

backfilling with water at the same flow rate as the decant was assessed with a fourth run. It was thought that maintaining the hydrostatic pressure with the backfill might prevent BDGReS. Potential gas release from the floating crust layer via dissolution was not addressed, given its relatively insignificant volume and release rate compared with BDGReS in these tanks.

Probability distributions of the results were created using a Monte Carlo simulation. The model was run 5,000 or 10,000 times^(a) for each case with input parameters for each run selected from probability distributions developed from data and theory. Each simulation represents a complete decant operation. The median and the value at the 95% confidence level (95th percentile of the cumulative frequency distribution) of the peak hydrogen concentration and the fraction of gas released during decant are given in Table 3.2

Table 3.2 shows that except for AN-103 the peak hydrogen concentrations from BDGReS induced by depressurization approximate those resulting from historic spontaneous events. Base case results for AN-104 and AN-105 were similar, showing a median peak hydrogen concentration of 2,800 ppm with the 95% confidence level below 25% of the LFL. The effect of the additional gas retained in AN-104 is counteracted by its larger headspace and lower retained hydrogen fraction. AN-103 contains about twice the gas volume of AN-105, and more of the waste is closer to neutral buoyancy. The result is the highest hydrogen concentrations of the four tanks investigated with a median of 10,500 ppm, just exceeding 25% of the LFL with the 95% confidence level at 21,100 ppm, about 50% of the LFL. The gas releases induced in AW-101 were inconsequential due to the shallow sediment layer and low retained hydrogen fraction. The median and 95% confidence level peak hydrogen concentrations were 400 and 1,500 ppm, respectively.

The gas releases induced in all four tanks were insensitive to decant rates from 30 to 200 gpm. Only AN-103 showed a measurable decrease in peak hydrogen concentration at the median from 10,500 ppm at 200 gpm to 8,200 ppm at 30 gpm. Similarly, the stop-start control strategy did not reduce the peak hydrogen concentrations measurably in any of the four tanks. The initial rise in hydrogen concentration above the action level always resulted from a BDGRE, and the gas release rate and duration of a BDGRE are not affected by decanting once initiated. Therefore, the peak hydrogen concentration was not affected. The main result of the stop-start control strategy was to increase the total time required for decant by a factor of two or three.

Although backfilling with water did not reduce the peak hydrogen concentration at the 95% confidence level, it caused an important reduction in the mean value. Water backfill cannot prevent a small reduction in hydrostatic pressure, so a few BDGReS were predicted in some of the runs. In runs where BDGReS occurred, the small headspace kept the peak hydrogen concentration higher than it would have been without the backfill. However, except for AN-103, the smaller reduction in hydrostatic pressure completely prevented BDGReS in over half of the runs. This allowed the median peak hydrogen concentration to remain at the background level of 23 to 37 ppm. In AN-103, BDGReS were prevented in only about 20% of the runs, so the

(a) All cases for AN-105 and base case runs for the other three tanks used 10,000 simulation runs. The sensitivity cases for tanks other than AN-105 used 5,000 simulation runs.

median peak hydrogen concentration was lowered to 1,500 ppm from 10,500 ppm but did not fall all the way to background. These results show that the water backfill strategy is probably beneficial. At worst, it does not increase the peak hydrogen concentration over that of the base case. At best, it creates a high probability of having no induced gas releases at all.

3.1.3 Gas Monitoring Considerations During Waste Removal

To track the important trends in the headspace hydrogen concentration due to induced BDGReS, gas monitoring must be essentially continuous. Peak hydrogen concentrations typically occur within about an hour of initiation of a BDGRE, and the concentration can exceed the action level in a much shorter time. As noted in Section 3.1.2, BDGReS can potentially be induced during waste removal in DSTs of any waste group. This is not a concern in Group C tanks or when supernatant removal is complete or nearly so, which is the only time BDGReS might occur in the current Waste Group B tanks. The analysis, which bounds Group B and C tanks, shows that the current Group A tanks remain below the LFL throughout the transfer but are expected to exceed the action level of 6,250 ppm at the 95% confidence limit.

It is also possible for delayed BDGReS to occur in Group A tanks after the transfer is shut down. Again, this is not a concern after transfer is complete or nearly so. Because Group A tanks exhibited spontaneous BDGReS at random times prior to waste removal, they should continue to do so during periods when transfer is temporarily halted (e.g., for repairs). In fact, the depressurization that occurred prior to the halt would make a BDGRE even more likely. However, the resulting hydrogen concentrations would be lower than historical norms because removing waste increases the tank headspace and lowers the fraction of gas released per event (see Eq. 3.5.7 and 3.5.8). Delayed releases would generally be smaller than historical norms because of the increased headspace and reduced gas pressure from the waste already removed.

Active ventilation on the order of 100 scfm dilutes the headspace back to near background concentrations within about one day. Because BDGReS occur over a few hours (rapid compared with the ventilation rate), the peak hydrogen concentration during a supernatant decant is not very sensitive to the ventilation rate. The ventilation rate was varied from 10 to 200 scfm in two of the Monte Carlo simulation runs (representing the same complete decant operation but with altered input parameters) for the base case in AN-105 that produced a peak hydrogen concentration of 10,000 ppm at 100 scfm. The peak hydrogen concentration results are plotted in Figure 3.1. The curve fit through the results indicates that even a very pessimistic passive ventilation rate would not cause the hydrogen concentration to exceed the LFL. On the other hand, increasing the ventilation rate to 500 scfm would only reduce the peak concentration to about 8,000 ppm, a 20% reduction from the base case.

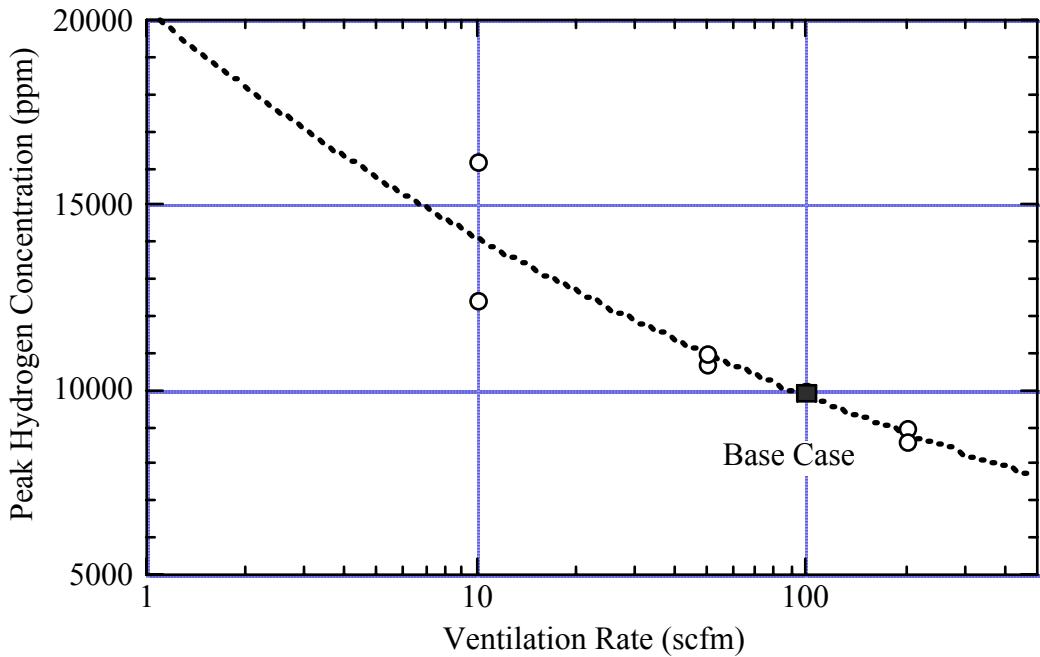


Figure 3.1. Sensitivity of Peak Hydrogen Concentration to Ventilation Rate in AN-105

The headspace hydrogen concentration due to liquid removal and drainage from uncovered sediment may be estimated with Eq. (3.1.10) and (3.1.11). The simple headspace mass conservation model developed to study gas releases during saltcake dissolution (Stewart 2001) and discussed in Section 3.4.3 is modified in Wells et al. (2002) to consider only liquid transfer out of a tank. The hydrogen concentration in the headspace at time i , C_{Hi} , may be determined from the following finite difference equation:

$$C_{Hi} = \frac{C_{Hi-1} + \frac{Q_{GAS}\chi_H}{V_{HSi}}\Delta t}{1 + \frac{(Q_{VIN} + Q_{GAS})}{V_{HSi}}\Delta t} \quad (3.1.12)$$

where V_{HSi} is the headspace volume at time i , C_{Hi-1} is the headspace hydrogen concentration at the previous time step, χ_H is the hydrogen fraction of the retained gas in the waste, Q_{VIN} is the headspace ventilation rate, and Δt is the time step.

AN-103, which has a large retained gas volume, is evaluated as an example. Note that, even with no supernatant, the headspace in AN-103 would reach flammability were all of its retained gas instantaneously released (gas is assumed to all be available at neutral buoyant conditions). For this analysis, it is conservatively assumed that the inlet to the transfer pump is at the bottom of the tank. An approximate dry solid density of $2,500 \text{ kg/m}^3$ is used, and the initial headspace hydrogen concentration is set to zero. All other waste input parameters are median values from Wells et al. (2002).

The maximum gas release volume from draining in AN-103 is computed from Eq. (3.1.11) to be approximately 208 scm resulting in a hydrogen concentration of approximately 34,000 ppm, or 0.84 LFL. Thus, to ensure the LFL cannot be reached at any point during the transfer, the transfer-induced headspace concentration must be kept below 6,000 ppm (0.15 LFL).

Figure 3.2 shows the predicted headspace hydrogen as a function of time for the case of liquid removal from the sediment layer in AN-103. Physical constraints on the liquid removal rate due to waste configuration are ignored. The tank headspace is conservatively assumed to remain constant.

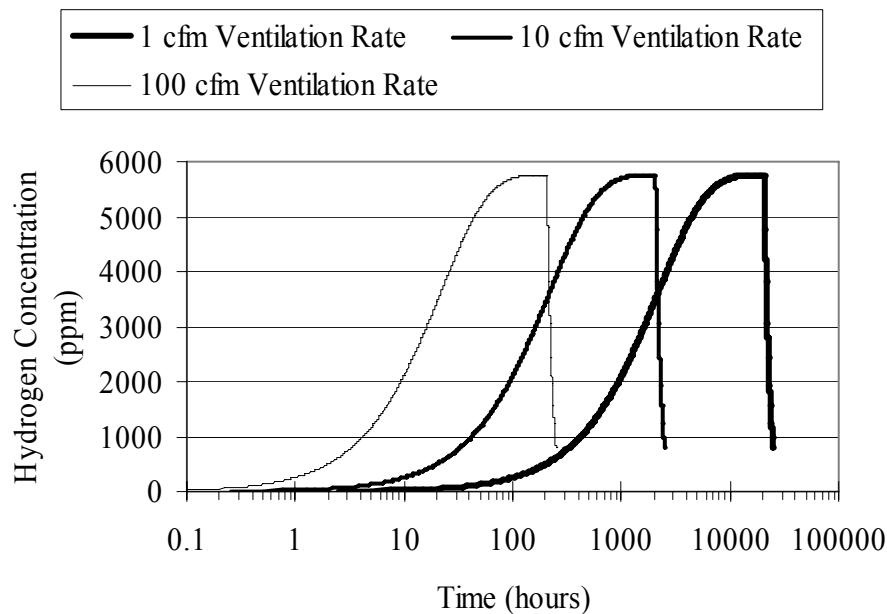


Figure 3.2. Predicted Hydrogen Concentration in AN-103 Resulting from Sediment Draining

Ventilation rates of 1, 10, and 100 cfm are used to approximate barometric, passive, and active ventilation. The maximum transfer rates that limit the headspace hydrogen concentration to 0.15 LFL (6,000 ppm) at complete drainage of the sediment are determined. The limiting transfer rate with 1 cfm of ventilation is approximately 0.2 gpm, with 0.15 LFL approached at approximately 15,000 hours (625 days). For a ventilation rate of 10 cfm, 0.15 LFL is approached at approximately 1,400 hours (58 days) with a transfer rate of 2 gpm. A transfer rate of 20 gpm causes the tank headspace to approach 0.15 LFL at approximately 140 hours with a ventilation rate of 100 cfm.

Loss of ventilation and/or transfer rate acceleration may potentially cause the headspace concentration to exceed 0.15 LFL (6,000 ppm) due to gas release from liquid transfer, thus allowing for LFL to be reached were the maximum drainage release (208 scm resulting in 34,000 ppm) to occur after transfer is stopped. The headspace ventilation must therefore be monitored at an adequate frequency to ensure that the transfer may be terminated such that 0.15 LFL is not reached. We conservatively assume that the gas release rate upon ventilation

failure is a combination of transfer and drainage. From Eq. (3.1.9), the time to reach the limiting gas volume, V_{Glim} , is therefore

$$t = \frac{V_{Glim}}{Q_D \left\{ \frac{\alpha}{(1-\alpha)\phi_L} \left[1 + \frac{1}{(1-\alpha)\phi_L} \right] \frac{P_0}{P_A} \right\}} \quad (3.1.13)$$

Experience from saltwell pumping in SSTs indicates that even the transfer/drainage rate of 2 gpm may be significantly greater than actual drainage rate (Section 3.3), thus the rate of gas release is probably exaggerated in the model. Further, as discussed in Section 2.3.4, sludge tanks apparently did not release a significant fraction of their retained gas during saltwell pumping, thus sediment drainage in a sludge tank may be expected to have significantly lower gas release rates than a saltcake tank.

In the AN-103 example with $Q_D = 2$ gpm and a ventilation rate of 100 cfm, the maximum headspace hydrogen concentration is approximately 600 ppm. From Eq. (3.1.13), the time required to reach the limiting gas volume of 33 scm (from $6,000 - 600 = 5,400$ ppm) is approximately 78 hours. Therefore, to protect against reaching LFL due to potential gas release from drainage, the headspace ventilation should be confirmed at intervals of less than 78 hours.

With an increased transfer rate of $Q_D = 20$ gpm at the ventilation rate of 100 cfm, the maximum headspace hydrogen concentration is approximately 5,800 ppm. Thus the limiting gas volume is 1 scm (200 ppm), requiring headspace ventilation monitoring at less than approximately 3 hour intervals.

3.1.4 Potential Changes in Waste Group Classification After Waste Removal

Complete supernatant removal prevents future BDGREs, which would move a tank from Group A to Group B at a minimum. It may also be possible to move a tank to Group C for partial supernatant removal if the resulting increase in headspace were sufficient. Partial supernatant removal in a Group B tank may potentially create a Group A tank due to the pressure reduction effect on the buoyancy ratio (see Eq. 2.13). A Waste Group A tank can be created by waste removal only if: 1) the supernatant removal is sufficient that the resulting pressure reduction in the sediment causes the buoyancy ratio to exceed unity, 2) the supernatant removal is not sufficient to result in an energy ratio less than three, and 3) the supernatant removal is not sufficient to result in a tank classification of Waste Group C. Waste Group evaluation is therefore required for waste removal.

3.1.5 Potential Changes in Hydrogen Generation After Waste Removal

The total hydrogen generation rate in the tank is reduced by the removal of hydrogen generating material. In addition, the higher surface-to-volume ratio of the remaining waste will cool and reduce the thermal portion of the total hydrogen generation rate. Thus the future steady-state flammability hazard is bounded by the analysis performed on the initial waste condition.

3.2 Waste Addition

Fluid waste (liquid waste or liquid-solid slurry) is pumped into a tank so that its actual introduction causes only a local disturbance. Fluid waste may be dumped directly onto the existing waste from a riser or injected under the waste surface through a drop-leg. If a future transfer system were designed to produce high-velocity jets near the tank bottom, mixing would be a concurrent global waste disturbance that should be treated as discussed in Section 3.6. For the analysis in this section, the incoming waste is considered sufficiently concentrated that solids dissolution can be ignored. Dissolution in SSTs by dilute waste is bounded by saltcake dissolution, Section 3.4, and, where hydraulic mixing may also be significant, by Huckaby and Wells (2004). Dissolution due to the addition of dilute waste in DSTs is bounded by water addition in Section 3.5.

Transfer of waste into a tank causes two kinds of global disturbances (dissolution is addressed in Section 3.5). It increases the hydrostatic pressure on the entire waste volume and potentially changes the density of the liquid layer. It also increases the depth of the supernatant and, if slurry is added or precipitation occurs, the sediment layer. Waste additions to SSTs that result in an energy ratio greater than 3 (see Section 2.3) are prohibited. Waste addition considered here is specific to DSTs.

3.2.1 Mechanisms for Gas Releases During Waste Addition

Increasing hydrostatic pressure prevents gas releases by compressing retained gas bubbles and reducing the gas volume fraction. Also, if the density of the incoming fluid is less than the density of the existing supernatant, the neutral buoyancy gas fraction is increased (see Eq. 2.12), which reduces the potential for a BDGRE. However, if the transfer involves a concentrated liquid or a liquid-solid slurry that increases the density of the existing supernatant, the neutral buoyancy gas fraction decreases and a BDGRE may be induced during the addition. As in supernatant decant, the first buoyant displacement can also suspend solids, further reducing the neutral buoyancy gas fraction and possibly inducing additional gas releases.

Other than induced BDGREs, the only mechanisms for gas release during waste addition are ammonia evaporation and release of “hitch-hiker bubbles” from incoming slurry. The rate of ammonia evaporation from a free liquid surface is greatly accelerated by any stirring action induced by waste addition. This effect is intensified if the incoming waste has a high concentration of dissolved ammonia. The hydrogen control limit for flammability is set, however, so that the possible presence of other gases is accounted for (Kripps 2005).

Transfer of a liquid-solid slurry from a gas-retaining tank by mixing, sluicing, or similar method is likely to create small bubbles that attach to particles or are so small they move with the slurry and do not separate until reaching the receiver tank. These “hitchhiker bubbles” create a chronic additional gas release in the receiver tank that is proportional to the transfer rate. However, the total volume of gas and the release rate are not enough to be of concern, as is shown in the next section.

Induced BDGREs are the primary mechanism for consequential gas releases during waste additions. If the incoming fluid is denser than the initial supernatant, the reduction in neutral buoyancy gas fraction quickly overcomes the opposing compression effect and could induce a BDGRE. The general principle is shown for a hypothetical waste addition to a tank in Figure 3.3, where the average gas volume fraction in the sediment and the neutral buoyancy gas fraction are plotted versus the volume of waste added. The added waste is assumed to have a density of 1,700 kg/m³ and the initial supernatant a density of 1,370 kg/m³. Waste addition compresses the gas in the sediment and lowers the gas volume fraction. However, the increasing supernatant density decreases the neutral buoyancy gas fraction much more rapidly. The two curves converge after an addition of about 50,000 gallons—when a BDGRE would occur.

Waste Group A tanks should be most susceptible to BDGREs induced by this mechanism.^(a) However, because the initial supernatant density is already quite high in these tanks and most are nearly full, a sizeable reduction in neutral buoyancy void fraction due to waste addition is unlikely. Only additions of slurry with a high solids loading (e.g., if the waste were mixed before or during transfer) could increase the supernatant density and lower the neutral buoyancy void fraction measurably in a receiving Group A tank.

Adding high-density waste to a borderline Group B tank might induce a BDGRE in extreme cases. However, this possibility can be determined beforehand and the added waste volume and density adjusted to prevent it. Because the neutral buoyancy gas fraction is very sensitive to supernatant density, calculating the buoyancy ratio (Eq. 2.13) to find out whether the waste addition could move the tank into Waste Group A in the subsequent steady state will also determine whether a BDGRE will occur during the operation, as is shown in the following analysis.

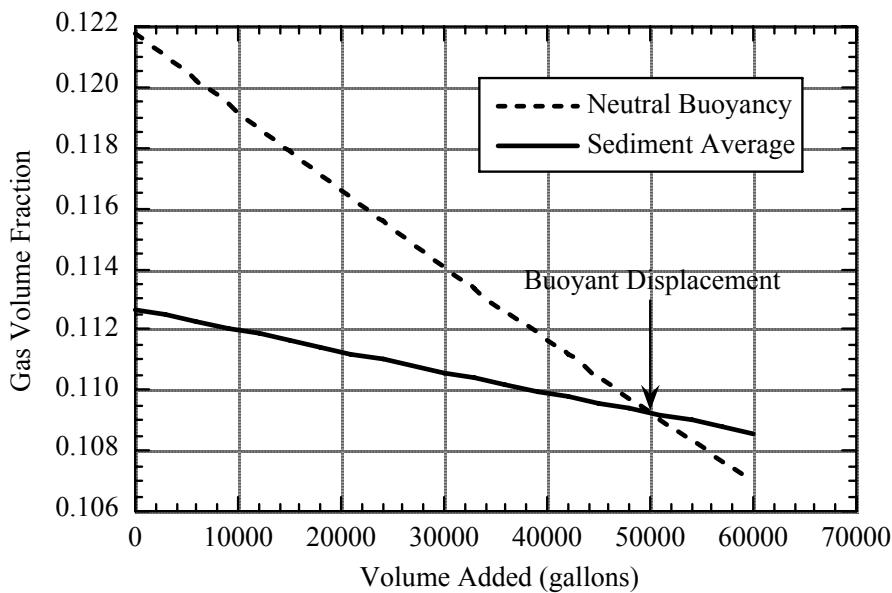


Figure 3.3. Mechanism for Inducing BDGREs During Waste Addition

(a) As described in Section 2.4.1, Waste Group A and B tanks store enough gas to make the headspace flammable if all of it were suddenly released; Waste Group C tanks do not. Waste Group A tanks exhibit spontaneous BDGREs, while tanks in Waste Group B do not.

Figure 3.4 illustrates the waste configuration during and after a waste addition. We assume a volume V_T of waste with density ρ_T is added to the supernatant, which has an initial density ρ_{L1} and thickness H_{L1} . After mixing, the supernatant will have an increased thickness, H_{L2} , and a new density, ρ_{L2} . The thickness of the gas-retaining sediment layer will decrease from H_{S1} to H_{S2} due to the compression of the retained gas from increased hydrostatic pressure. The gas volume fraction will be reduced from α_1 to α_2 in the process.

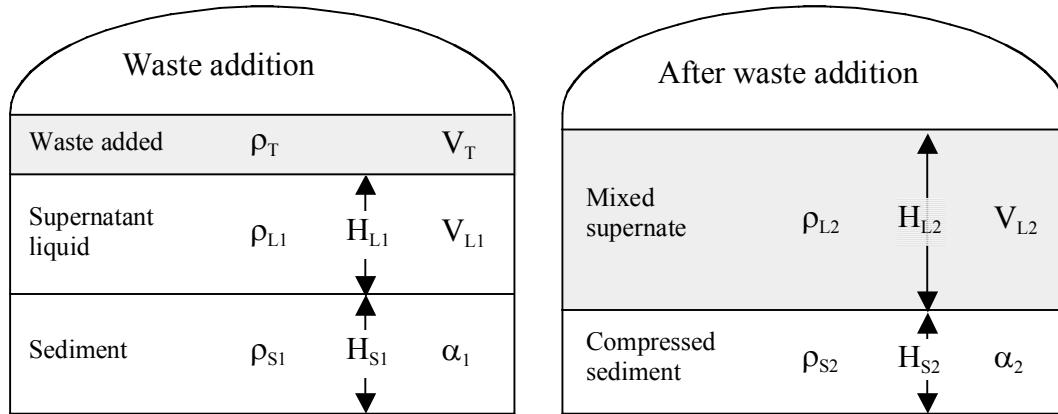


Figure 3.4. Tank Waste Configuration During and After Waste Transfer

Mass conservation for the mixed supernatant liquid after the transfer dictates that

$$\rho_{L2}V_{L2} = \rho_{L1}V_{L1} + \rho_T V_T \quad (3.2.1)$$

If we assume constant volume mixing (generally correct to first order, neglecting dissolution or precipitation), then

$$V_{L2} = V_{L1} + V_T \text{ and } H_{L2} = H_{L1} + V_T / A \quad (3.2.2)$$

where A is the tank cross-sectional area. Combining Eq. (3.2.1) and (3.2.2) results in an expression for the mixture density in terms of the transferred quantities:

$$\rho_{L2} = \frac{\rho_{L1}V_{L1} + \rho_T V_T}{V_{L1} + V_T} \quad (3.2.3)$$

The gas volume fraction in the sediment decreases in response to the increase in hydrostatic pressure in accordance with Eq. (3.1.2):

$$\alpha_2 = \frac{1}{1 + \frac{P_2(1-\alpha_1)}{P_1 \alpha_1}}$$

The pressure ratio can be determined by applying Eq. (3.1.4) for the average pressure in the sediment. It is useful to ignore the small change in height of the sediment layer so that $H_{S1} \approx H_{S2} = H_S$. With this simplification, the pressure ratio is found to be

$$\frac{P_2}{P_1} = \frac{1 + \frac{g}{P_A} [\rho_{L2}(H_{L2} + H_C) + \rho_S H_S / 2]}{1 + \frac{g}{P_A} [\rho_{L1}(H_{L1} + H_C) + \rho_S H_S / 2]} \quad (3.2.4)$$

Eq. (3.2.4), along with Eq. (3.2.3) and (3.1.2), gives the final gas volume fraction in terms of the initial conditions and the properties of the added waste.

If the final gas volume fraction is greater than the final neutral buoyancy gas fraction, a BDGRE will occur during the waste addition. That is, if the ratio of the average gas volume fraction to the neutral buoyancy value is greater than unity, a BDGRE will occur. The gas volume fraction at neutral buoyancy is given by Eq. (2.12). Combining this with Eq. (3.1.2) for final gas volume fraction provides an expression for final buoyancy ratio at the end of transfer:

$$BR_2 = \frac{\alpha_2}{\alpha_{NB2}} = \frac{\alpha_1 \rho_S / (\rho_S - \rho_{L2})}{\alpha_1 + (1 - \alpha_1) P_2 / P_1} \quad (3.2.5)$$

The buoyancy ratio long after the transfer is complete and the tank reaches its new steady state can also be evaluated directly from Eq. (2.13) to determine whether the result of the waste addition will eventually cause BDGREGs and move the tank classification into Waste Group A. Denoting this as BR_3 ,

$$BR_3 = \frac{\alpha_2}{\alpha_{NB2}} = \frac{C_B}{\rho_S - \rho_{L2}} \left(\frac{GT}{P_2} \right)^{1/3} H_S^2 \quad (3.2.6)$$

The final density is calculated with Eq. (3.2.3) and the final pressure with Eq. (3.1.4). The small change in sediment depth is ignored, and the temperature and gas generation rate in the sediment are assumed unaffected by the waste addition. Because the long-term steady-state buoyancy ratio, BR_3 , via Eq. (3.2.6) is always greater than that determined from the process model of Eq. (3.2.5), a determination that there will be no change in waste group (i.e., $BR_3 > 1$) also ensures that no BDGREGs will occur during the process (i.e., $\alpha_2 < \alpha_{NB2}$). To demonstrate this, we recast Eq. (3.2.5) for the buoyancy ratio at the end of transfer in terms of the initial buoyancy ratio prior to transfer by substituting for the initial gas volume fraction, α_1 , in the numerator:

$$BR_2 = \frac{1}{\alpha_1 + (1 - \alpha_1) P_2 / P_1} BR_1 \frac{\rho_S - \rho_{L1}}{\rho_S - \rho_{L2}} \quad (3.2.7)$$

The analog to Eq. (3.2.7) for the buoyancy ratio long after transfer is computed via Eq. (3.2.6) as

$$BR_3 = BR_1 \frac{\rho_s - \rho_{L1}}{\rho_s - \rho_{L2}} \left(\frac{P_1}{P_2} \right)^{1/3} \quad (3.2.8)$$

The ratio of long-term to short-term buoyancy ratios is greater than unity except for a very large initial gas fraction, as shown by

$$BR_3 / BR_2 = \left[\alpha_1 + (1 - \alpha_1) P_2 / P_1 \right] \left(\frac{P_1}{P_2} \right)^{1/3} > 1 \text{ for } \alpha_1 > 2/3 \quad (3.2.9)$$

Temperature and gas volume are also related. Adding higher-temperature waste would increase the sediment gas volume fraction and could potentially cause it to become buoyant. Assuming a constant pressure and instantaneous heat transfer from the mixed liquid layer into the sediment, the change in gas volume fraction due to a change in temperature is expressed analogously to Eq. (3.1.2) by

$$\alpha_2 = \frac{1}{1 + \frac{T_1}{T_2} \frac{(1 - \alpha_1)}{\alpha_1}} \quad (3.2.10)$$

where T_1 and T_2 are the absolute temperatures before and after waste addition, respectively. With an initial gas volume fraction of 0.1 and an initial temperature of 100°F, an increase in the sediment temperature of 65°F would be required to raise the gas fraction 10%. This means that the initial buoyancy ratio would have to be 0.9 or greater to cause a BDGRE. We conclude that temperature changes are not an important issue in gas releases during waste addition.

3.2.2 Expected Gas Releases During Waste Addition

No specific analysis has been performed on gas release rates and volumes of BDGRe induced in Group A tanks during waste addition. Because the initiating mechanism is sediment buoyancy, as it is in spontaneous releases, BDGRE gob size would be expected to follow historic norms. However, the peak hydrogen concentration would be higher because waste addition decreases the headspace and increases the gas release fraction by raising the hydrostatic pressure (see Eq. 3.5.7 and 3.5.8).

There has been no experience or analysis on the size of BDGRe induced in Group B or C tanks that have not exhibited them in the past. In the absence of historic BDGRe the waste should be more uniform, so the first induced BDGRE might be larger than indicated by the behavior of the current Group A tanks. However, if the waste addition is adjusted to prohibit BDGRe, no appreciable gas release is expected from any other mechanism.

Waste addition tends to increase ammonia evaporation by disturbing the supernatant and refreshing the liquid surface. Ammonia concentrations as high as 7,000 ppm were observed in SY-102 while it was receiving waste from SY-101 (Mahoney et al. 2000). Ammonia was also observed to increase to about 1,500 ppm in SY-102 while it received saltwell liquor.

The “hitchhiker” bubble phenomenon apparently caused hydrogen concentrations to rise in AY-102 when it was receiving sludge sluiced from C-106 (Cuta et al. 1999). The hydrogen concentration in SY-102 increased to about 240 ppm when receiving mixed slurry from SY-101 (Mahoney et al. 2000). However, the gas release rates in both these cases were relatively low, and there was insufficient gas in the entire transfer to raise the hydrogen concentration to the action level, even with zero ventilation. Hitchhiker bubbles were determined to be a negligible effect in the brine transferred from tanks being saltwell pumped (Peurrung et al. 1998). No hydrogen elevation was detected in SY-102 while it received saltwell liquor.

3.2.3 Gas Monitoring Considerations During Waste Addition

BDGReS induced by any operation occur over a relatively short period that requires essentially continuous monitoring to capture the approach to peak hydrogen concentration. However, in the case of waste addition, the calculations outlined in Section 3.2.1 can be applied to adjust a given transfer to avoid creating a tank exhibiting spontaneous BDGReS, which also prevents BDGReS from occurring during the transfer.

Transfers of waste from tanks in which the waste has been mixed or otherwise agitated can carry hitchhiker gas bubbles that are released in the receiver tank. If the ventilation system failed while the transfer continued, the hydrogen concentration could exceed the action level, though it is doubtful that the LFL would be reached. Because the mechanics of saltwell pumping separates the retained gas bubbles from the brine, and because of the very low transfer rate, hitchhiker bubbles are not an issue in tanks receiving saltwell liquor, even without active ventilation. To quantify these assertions, the simple headspace mass conservation model developed to study gas releases during saltcake dissolution (Stewart 2001) and discussed in Section 3.4.3 is simplified to consider only inflow of waste to a tank. The headspace hydrogen concentration is expressed as a function of time as

$$C_H(t) = \frac{Q_{GAS}\chi_H}{Q_{VIN} + Q_{GAS}} + \left(C_0 - \frac{Q_{GAS}\chi_H}{Q_{VIN} + Q_{GAS}} \right) \left(1 - \frac{Q_{TIN}}{V_{HS0}} t \right)^{\frac{Q_{VIN} + Q_{GAS}}{Q_{TIN}}} \quad (3.2.11)$$

where

- C_0 = initial headspace hydrogen concentration
- V_{HS0} = initial headspace volume
- χ_H = hydrogen fraction of the retained gas in the waste
- Q_{GAS} = gas release rate from bubbles in the transfer stream
- Q_{VIN} = estimated headspace passive (or measured active) ventilation rate
- Q_{TIN} = incoming transfer flow rate.

As a bounding case, consider an empty DST being filled with slurry containing a gas volume fraction of 0.03. This was the maximum value attributed to the SY-101 mixed slurry transferred to SY-102 during remediation and is likely to be the bounding value for any transfer. Assume the gas contains 60% hydrogen, twice the concentration in SY-101 waste but representative of other DSTs that might be mixed prior to transfer. The peak hydrogen concentration occurs when the tank is filled to the maximum level, assumed to be 422 inches. At a transfer rate of 100 gpm

with no in-line dilution, the tank would be filled in eight days, and Eq. (3.2.11) predicts a peak hydrogen concentration of 2,400 ppm at a 100-scfm ventilation rate. Using a passive ventilation rate of 2.5 scfm, conservatively representative of a failed active ventilation system (Hu 2004b), the peak hydrogen concentration would be 26,000 ppm and the action level of 6,250 ppm would have been exceeded in 70 hours (about three days). With zero ventilation, the peak hydrogen concentration would be 30,000 ppm, still under the LFL for hydrogen alone. A concurrent ammonia concentration of over 50,000 ppm would be necessary to reach 100% of the LFL.

For the lower transfer rates of saltwell pumping, a much longer time is required, and the hydrogen concentration is much more sensitive to ventilation. Still assuming the bounding gas loading in the transfer stream but a 10 gpm transfer rate (might represent three or four SSTs being saltwell pumped simultaneously), the peak hydrogen concentration is only 240 ppm for a 100 scfm ventilation rate. For 2.5-scfm passive ventilation, the peak is 9,000 ppm after 81 days, exceeding the action level of 6,250 ppm in 43 days. Zero ventilation gives the same 30,000-ppm peak as the 100-gpm transfer, while the action level is exceeded in 29 days. If completely filling a sealed (zero ventilation) empty tank with a highly gassed slurry reaches only 75% of the LFL, smaller transfers with lower gas loading pose no hazard with or without ventilation.

3.2.4 Potential Changes in Waste Classification After Waste Addition

As described in Section 3.2.1, it is possible to reduce the neutral buoyancy gas fraction enough that a Group B or C tank could eventually experience BDGReS and become a Group A tank. However, a very large waste addition at high density would be required. Figures 3.5 and 3.6 show the results of buoyancy ratio calculations using Eq. (2.13) for various transfer volumes and densities into Waste Group B and C tanks. Table 3.3 shows the inputs used in the calculations for each tank. The added waste volume in the calculations was limited to that required to raise the waste level to about 10 m.

The Waste Group C Tank has a very low initial supernatant density, so adding high-density waste has a strong effect. About 70,000 gallons of additional waste at a density of 1,500 kg/m³ has the potential to make this tank exhibit BDGReS. However, the headspace reduction of this small addition would probably not be sufficient to move the tank out of Waste Group C. If the incoming density is reduced to 1,300 kg/m³, the buoyancy ratio reaches unity at about 280,000 gallons of waste addition. At this point, the headspace is about 75% of the initial volume, so it might now become flammable if all of the retained gas were released, and the tank would move to Waste Group A. With a density of 1,200 kg/m³, the buoyancy ratio does not approach unity with even the maximum possible addition, though the greatly reduced headspace after a very large transfer would likely place the tank in Waste Group B.

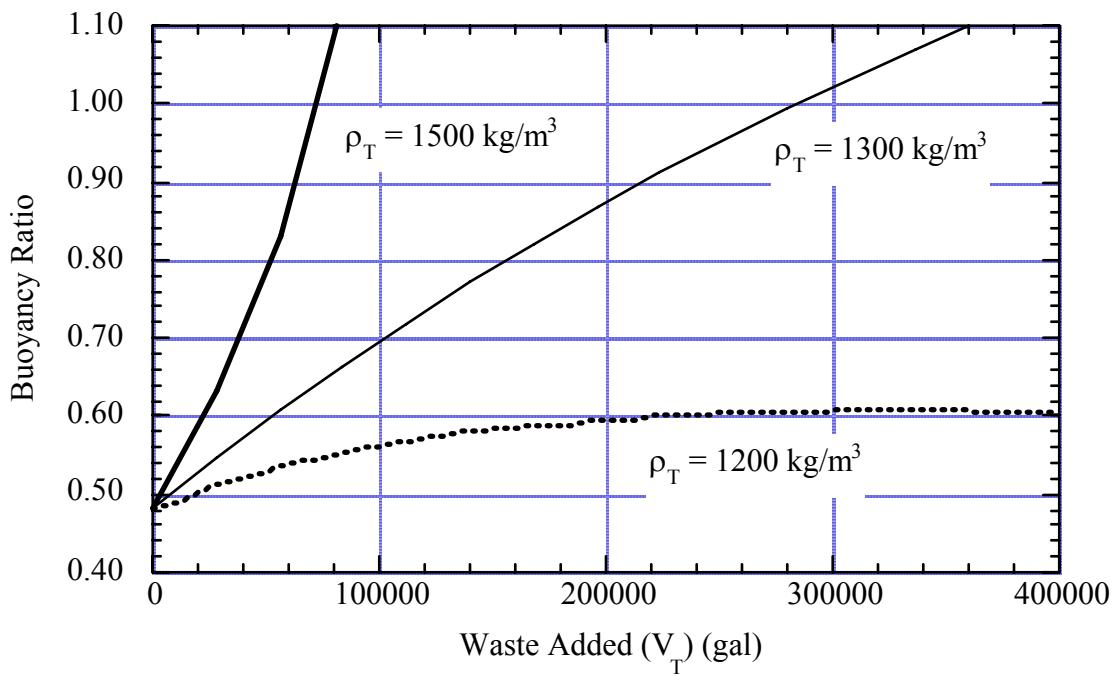


Figure 3.5. Effects of Waste Addition on Buoyancy Ratio for Waste Group C Tank

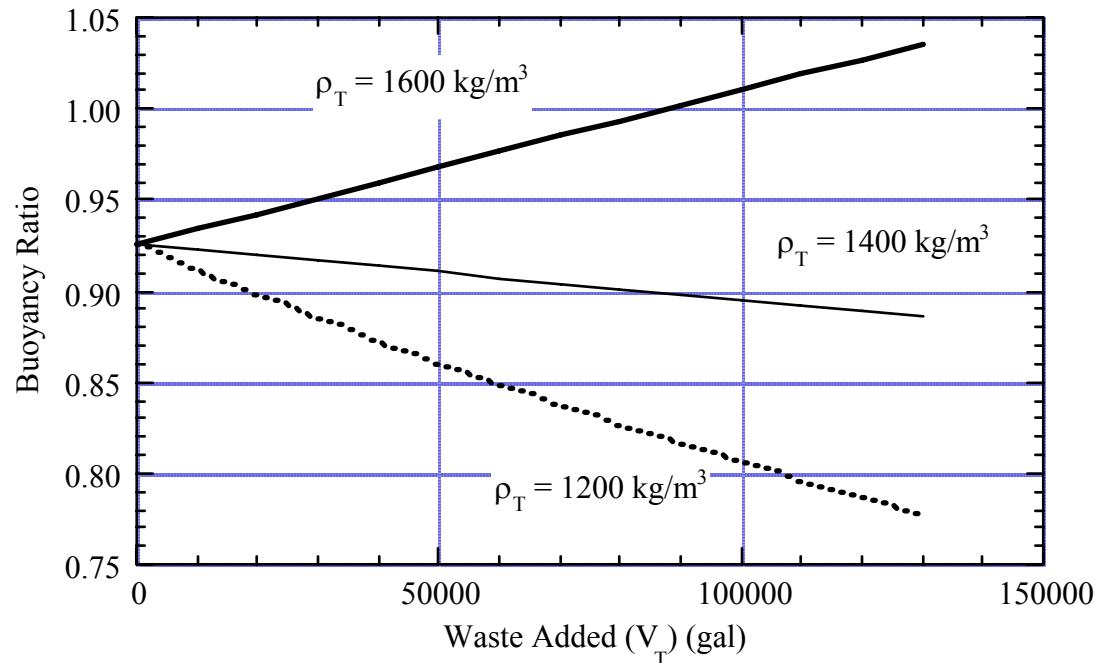


Figure 3.6. Effects of Waste Addition on Buoyancy Ratio for Waste Group B Tank

Table 3.3. Inputs Used for Waste Transfer Analysis

Property	Waste Group C Tank	Waste Group B Tank
H_{L1} (m)	1.6	7.3
H_S (m)	2.4	2.3
ρ_{L1} (kg/m ³)	1070	1370
ρ_S (kg/m ³)	1306	1560
α_{NB1}	0.18	0.12
G (mol/m ³ /day)	0.0017	0.0116
T (K)	293	305
BR ₁	0.48	0.93

The waste level in the Waste Group B tank is higher than in the Waste Group C tank, so only about 120,000 gallons may be added. The supernatant is much less dilute, so only the densest waste addition, 1,600 kg/m³, can decrease the neutral buoyancy gas fraction and, because the sediment density is higher than in the Waste Group C tank, the effect is less pronounced. The buoyancy ratio will exceed unity in this tank if the maximum of 120,000 gallons of waste with a density of 1,600 kg/m³ is added.

These examples show that it is possible for a waste addition to elevate a tank's classification from Waste Group C or B to A, where the tank may exhibit BDGReS and have a sufficiently small headspace that a sudden release of the entire gas inventory could make it flammable, even without the potential increase in the sediment depth. This points out the need for tank classification analyses prior to any waste addition to ensure the proper controls are in place prior to moving a Group C tank into Waste Group B and adjusting the operation to prevent moving a tank into Waste Group A.

While additions to Waste Group A tanks cannot raise their classification, the reduced headspace will increase the hydrogen concentration resulting from continuing spontaneous BDGReS. Also, the increased liquid depth and consequent increased hydrostatic pressure will increase the volume of gas released in a BDGRE. Any solids that are transferred in add to the sediment depth, providing a larger total retained gas volume. All these effects combine to exacerbate the consequences of existing BDGReS.

3.2.5 Potential Changes in Hydrogen Generation After Waste Addition

By definition, adding waste to a tank reduces its headspace and may increase its total hydrogen generation rate. Both will exacerbate steady-state hydrogen concentration and time to flammability. Therefore, the steady-state hydrogen generation rate must be reassessed for any waste addition.

3.3 Saltwell Pumping

Saltwell pumping is the primary method for removing drainable liquids from the SSTs to prevent or reduce the consequences of a leak. Saltwell pumping at Hanford has been completed as part of Interim Stabilization.

For saltwell pumping, a long cylindrical metal screen (the saltwell screen) is installed as a well casing near the center of the tank and extends virtually to the bottom of the tank. A jet pump located inside and at the bottom of the saltwell screen pumps liquid out as it drains into the saltwell screen.

Saltwell pumping is a gradual process. At the onset of pumping, supernatant flows freely into the well and the pumping rate is limited by pump capacity. After the supernatant is exhausted, the rate at which brine enters the saltwell screen slows as brine drains through the waste. The pumping rate is then reduced to approximately match the liquid drainage rate. The pumping campaign is complete when the pumping rate falls below some predetermined value (less than 1/2 gpm).

Campaigns have been as short as three months (Tank S-103, from which less than 24,000 gallons of liquid were removed) and as long as a year (Tank SX-104, from which about 117,000 gallons of liquid were removed). The duration of saltwell pumping in any given tank is a function of the amount of drainable liquid, the drainage rate (which depends on permeability), and the actual times the system is available for pumping (stoppages for corrective and preventive maintenance are common).

Besides the removal of liquid, the most important global waste disturbance of saltwell pumping in saltcake waste is the increased lithostatic load in the waste column as the buoyant force of the interstitial liquid is removed. This creates a new waste configuration with an unsaturated layer of waste overlying a saturated zone. Only the saturated regions of the sediment are capable of retaining flammable gas and, as described in Section 2, the increased lithostatic load changes the retained gas configuration to a form with a different release behavior. The increased lithostatic load also leads to compaction and subsidence of the central portion of the tank that may eventually create a broad crater or depression. Sludge gives up liquid by consolidation rather than draining, and neither the waste configuration nor the gas retention mechanism changes during pumping. The rest of this section concerns pumping of saltcake tanks.

3.3.1 Mechanisms for Gas Release During Saltwell Pumping

Removing supernatant and interstitial liquid by saltwell pumping can induce gas releases by several mechanisms. The hydrostatic pressure within the waste decreases as liquid is removed, which causes trapped gas bubbles to expand and dissolved gases to evolve into the bubbles. Bubble growth from both mechanisms can cause bubble disengagement or percolation gas release (see Section 2.2). Particle-displacing bubbles that were not released by these mechanisms "pop" when the interstitial liquid drains away from around them (Peurrung et al. 1997). However, some liquid remains after the bulk of it has been drained away, and some gas will remain trapped as small, pore-filling bubbles surrounded by this undrainable liquid. Bubbles in this condition are eventually released when the entrapping liquid evaporates or gradually drains away.

The evaporation of moisture that is held up in the unsaturated waste after initial draining is a source of ammonia and other soluble gases (Peurrung et al. 1997). Because this liquid is trapped in small crevices between particles, its surface area is huge and the evaporation rate can be high.

However, the evolution of these gases is accounted for in the hydrogen control limit for flammability (Kripps 2005).

3.3.2 Expected Gas Releases During Saltwell Pumping

Gas is released from the waste as it becomes unsaturated so the gas release rate roughly follows the pumping rate. Hedengren et al. (2001) give a summary of observed gas release behavior during saltwell pumping based on headspace gas monitoring data. The correlation between saltwell pumping and gas release is illustrated in Figure 3.7, where the hydrogen release rate in Tank U-105 is plotted along with the daily volume of liquid pumped. Headspace ventilation rates were determined from a coincident tracer gas study and used with headspace hydrogen and nitrous oxide concentrations to estimate their release rates.^(a) In Figure 3.7, we see about a one-day lag between pumping activity and gas release. But the lag is greater in several cases, and pumping activities are not always accompanied by increased gas release rates. This suggests that the timing and magnitude of gas releases are also subject to factors other than pumping.

Similar correlations between pumping and gas release are also observed in headspace gas monitoring data in other tanks (Huckaby et al. 1999).^(a) The highest gas release rates occur shortly after the onset of pumping, when the waste drains rapidly and pumping rates tend to be high. This is illustrated in Figure 3.8, where the estimated hydrogen release rate^(b) has been

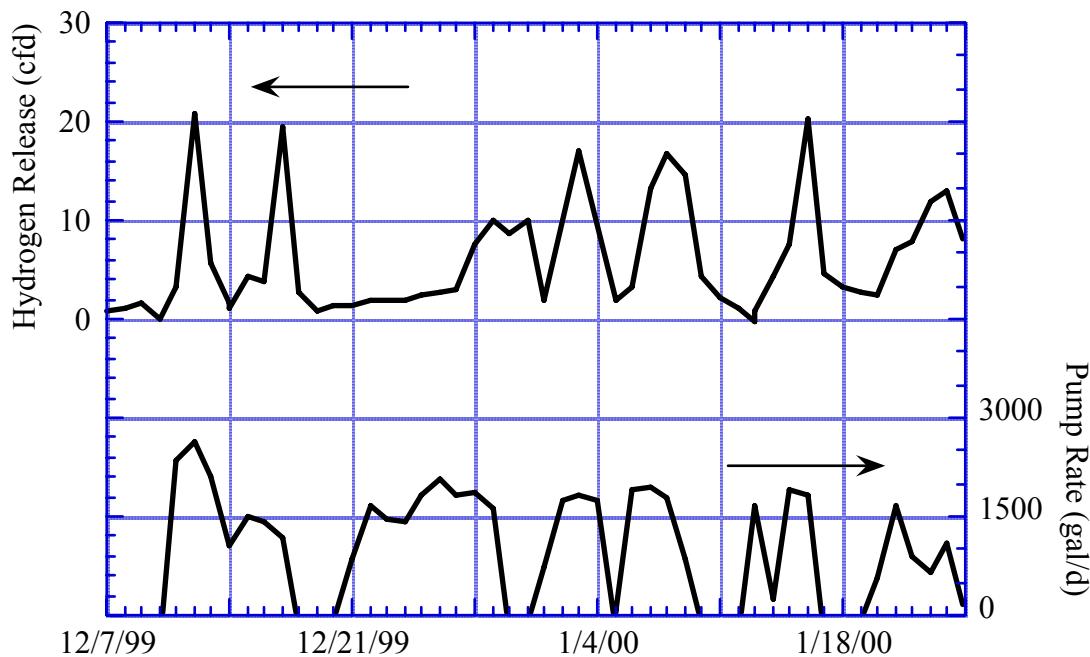


Figure 3.7. U-105 Hydrogen Release and Saltwell Pumping Rates

(a) Peurrung LM and JL Huckaby. March 10, 2000. *Gas Release Behavior During Saltwell Pumping*. Letter Report TWS00.39, Pacific Northwest National Laboratory, Richland, WA.

(b) The negative gas release rates in Figure 3.7 are not physically reasonable and indicate that the actual ventilation rate at those times was considerably higher than the 9.1 ft³/min used in calculations.

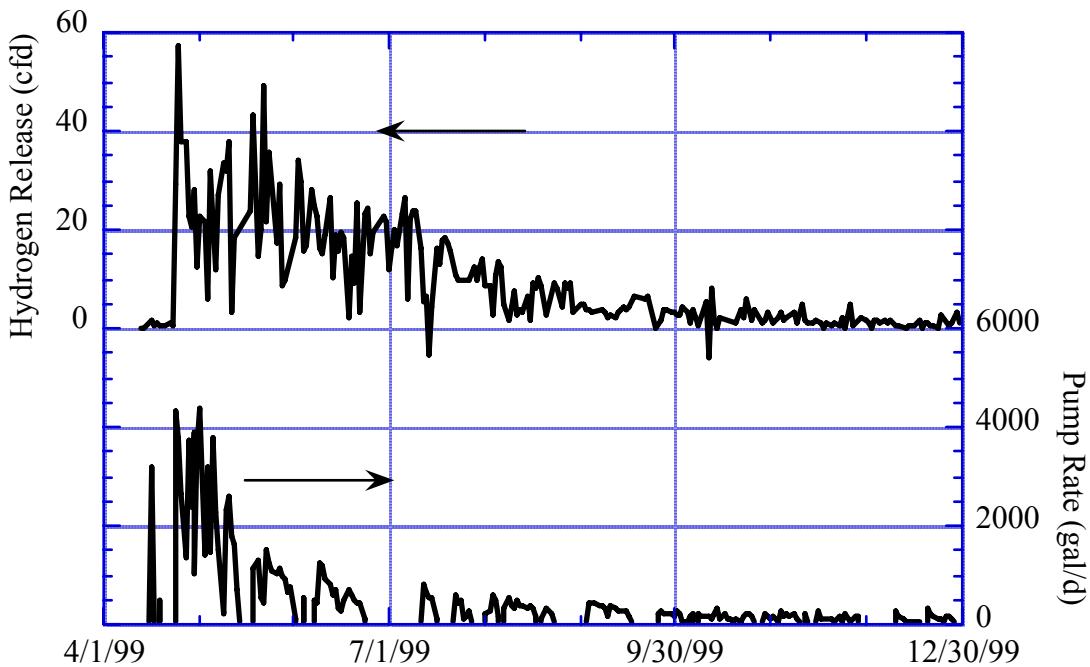


Figure 3.8. Tank 241-S-106 Estimated Hydrogen Release Rate and Pumping Rate

plotted above the pumping rate for S-106. Much of the liquid being pumped in this initial period consists of supernatant and interstitial brine immediately surrounding the saltwell screen that drains quickly. During this phase, starting and stopping the saltwell pump can cause gas release rates to rise and fall rapidly, as shown in Figure 3.7 for U-105.

Continued pumping exhausts the supernatant and depletes the region near the saltwell screen of liquid; thus the drainage rate decreases. Gas release rates decline correspondingly, as indicated in Figure 3.8. In this second phase of pumping, the liquid saturation interface (or interstitial liquid level) rises gradually from a low point at the saltwell to a higher level near the tank wall. The majority of liquid being pumped during this phase comes from the regions away from the saltwell screen. A temporary cessation in pumping has little immediate effect on drainage rate far from the saltwell, and the gas release rate is thus more independent of pumping rate.

After pumping ceases, drainage continues from the outer regions of the waste toward the saltwell screen as well as from the upper regions of the waste that had not drained completely. This continues to release trapped gases, albeit slowly. The headspace hydrogen and nitrous oxide concentrations in Tank S-106, for example, appear to have remained high for months after saltwell pumping was stopped.^(a)

(a) Peurrung LM and JL Huckaby. March 10, 2000. *Gas Release Behavior During Saltwell Pumping*. Letter Report TWS00.39, Pacific Northwest National Laboratory, Richland, WA.

Because ammonia is highly soluble in liquid wastes, headspace ammonia concentrations are not expected to rise and fall as pumping starts and stops (Peurrung et al. 1997). A reservoir of dissolved ammonia exists where the waste itself remains wet; it releases ammonia vapor to the headspace as mass transport allows. Saltwell pumping apparently did not affect the ammonia concentration in Tank U-105 until the supernatant was removed; then the increased wetted surface area (associated with the exposed porous waste surface) caused a rapid rise in ammonia concentration (Figure 3.9).^(a)

The overall result of the saltwell pumping process in saltcake tanks is the release of a substantial fraction of the retained gas, though it is impossible to quantify it accurately. Approximate integration of the hydrogen release rate over the pumping period in S-106 indicates a total gas release of $160 \pm 40 \text{ m}^3$ (Hedengren et al. 2001). This volume ranges from 30 to 70% of the total retained gas volume estimated by several methods. Approximately half of the retained gas volume in U-105 was released by saltwell pumping.^(a) Similar fractions could be expected in other saltcake tanks. As discussed in Section 2.3.4, sludge tanks are not expected to release a significant fraction of their retained gas during saltwell pumping.

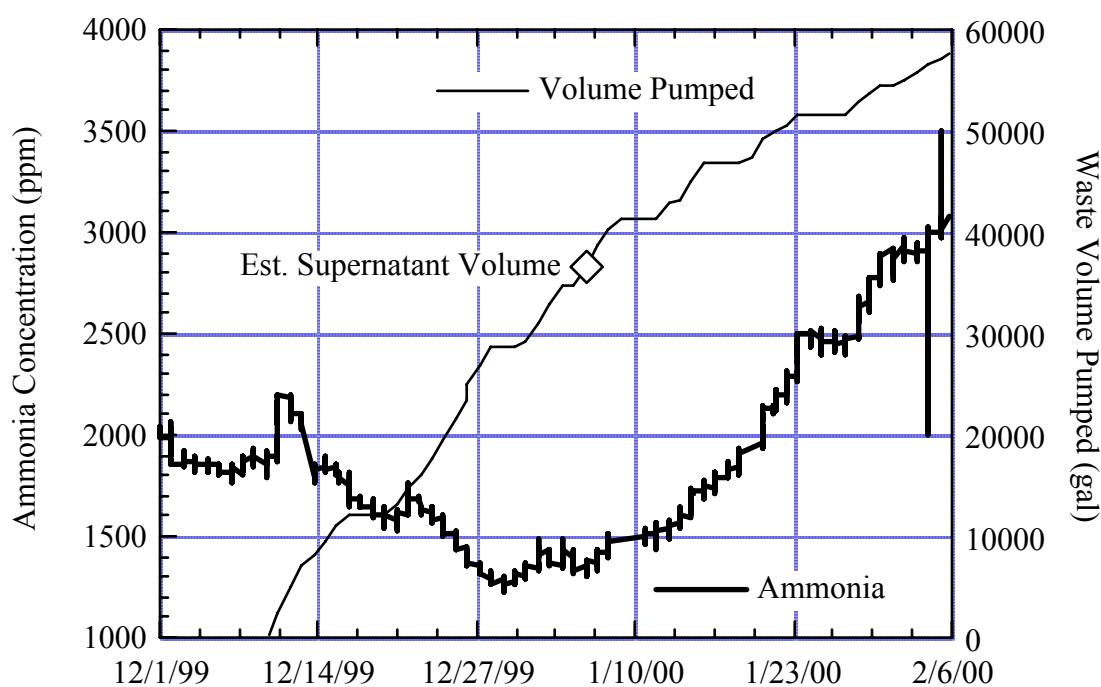


Figure 3.9. U-105 Ammonia Release and Waste Volume Pumped

(a) Huckaby JL, LA Mahoney, and ER Siciliano. 2001. *Waste Gas Releases Associated with Recent Saltwell Pumping of Single-Shell Tanks*. TWS01.036, Pacific Northwest National Laboratory, Richland, WA.

3.3.3 Gas Monitoring Considerations During Saltwell Pumping

Hydrogen concentration measurements in SSTs from 1995 to 2001 (McCain 2001) show that gas releases induced by saltwell pumping are unimportant in terms of flammability. The highest hydrogen gas concentration ever measured by standard hydrogen monitoring systems (SHMS) in an SST was 7,200 ppm, detected in BY-106 during saltwell pumping in 1995 (Watrous et al. 2000). U-103, U-105, and U-109 had hydrogen concentrations approaching 5,000 ppm during pumping in 1999 and early 2000. Saltwell pumping in S-111 was shut down on February 7, 2002 when the hydrogen concentration passed 5,500 ppm; it eventually reached a peak of 6,600 ppm on February 11. The headspace hydrogen concentration and waste level for the period of interest are plotted in Figure 3.10.

Saltwell pumping gas release rates are quite low but can become an issue in SSTs because the low passive ventilation rates can allow released gas to accumulate in the headspace. However, hydrogen concentrations recorded in several tanks during pumping have been well below 25% of the LFL. Gas releases during saltwell pumping can be controlled, with some delay, by shutting down the pump (Watrous et al. 2000). This ability to control gas releases is also demonstrated by the behavior in S-111 after pumping was shut down (Figure 3.10).

During saltwell pumping, headspace sampling is required periodically (e.g., weekly) to detect whether the hydrogen concentration is approaching the action level (Watrous et al. 2000). This low frequency is consistent with very low gas release rates, but monitoring is needed because the potentially low ventilation rates allow gas to accumulate. Experience with S-111 was a good test of this requirement because it already had continuous gas monitoring. Applying a weekly schedule to the hydrogen concentration transient in Figure 3.10 and assuming the first

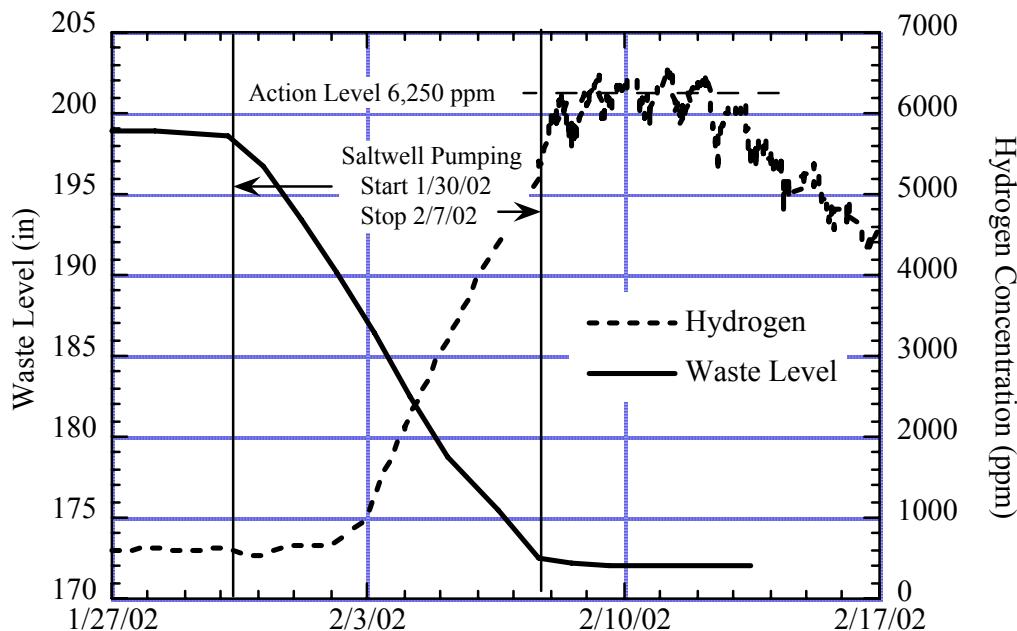


Figure 3.10. Hydrogen Concentration and Waste Level History in S-111

measurement was made when pumping began January 30, a measurement would have been made on February 6, when the concentration was over 4,000 ppm. If this relatively high reading did not trigger more frequent monitoring, a third measurement would have occurred February 13. Extrapolating the slope of the hydrogen concentration curve from February 6 indicates it could have reached as high as 10,000 ppm at the third sample. This implies that strict weekly monitoring would not have prevented exceeding the action level but would have permitted pumping to shut down well before reaching the LFL.

3.3.4 Potential Changes in Waste Classification After Saltwell Pumping

While saltwell pumping produces a marked change in the waste configuration of SSTs, its main effects are to release up to about half of the gas inventory in a tank and to reduce the amount of liquid-saturated waste in which new gas can be generated and stored. The absence of supernatant and the presence of a high lithostatic load also prevent any possibility of a BDGRE. At the same time, removal of supernatant and subsidence of the drained sediment increases the headspace, providing more dilution volume for future releases. All these changes would tend to move a tank's classification toward Waste Group C. Therefore, while reevaluation of the tank classification may be operationally beneficial, it is not required for saltwell pumping.

3.3.5 Potential Changes in Hydrogen Generation After Saltwell Pumping

Saltwell pumping reduces the potential for steady-state flammability by removing liquid waste from the tank and increasing the headspace. Therefore, the steady-state gas generation rate is bounded by the prepumping calculation and need not be reevaluated.

3.4 Saltcake Dissolution

Saltcake dissolution is a method of waste retrieval from SSTs containing saltcake (Estey et al. 2001). Water is added to the waste to dissolve the soluble salts, and the resulting brine is pumped out of the tank and transferred to a DST at about the same rate that it is created. Thus, saltcake dissolution is distinctly differentiated from water addition, Section 3.5; water is added only at the rate at which liquid is removed. Though the water addition and pumping system may also remove insoluble solids, the primary product delivered to the receiving DST is a concentrated liquid. Saltcake dissolution does not include small water additions associated with line flushes, lancing in instruments or saltwell screens, and the like, which are covered in Section 3.5. The addition of brine to the receiving DST is covered under Waste Addition in Section 3.2.

Batch transfer followed by dissolution was used to dissolve the thick crust layer and remediate gas retention in SY-101 (Johnson et al. 2000). A similar batch process is also planned as a step in delivering waste feed from the DSTs to the vitrification plant. A large volume of water would be added to the tank following removal of most of the existing supernatant. After dissolution, which may be aided by mixing, the brine would be transferred to a staging tank. This operation is only a conceptual plan and is not yet described in sufficient detail to permit analysis.

The global waste disturbance of saltcake dissolution is the destruction of the waste microstructure by dissolution of solid particles and the major reduction in waste volume.

3.4.1 Mechanisms for Gas Release During Saltcake Dissolution

Stewart (2001) studied gas release during saltcake dissolution. Any gas retained in the region of the waste in which soluble solids are dissolved is assumed to be released. Waste in tanks with mostly insoluble waste would not be disturbed by water addition, and little gas would be released. Because the solvent fluid is less dense than the saturated interstitial brine, the rate of dissolution and therefore the rate of gas release are limited by the rate at which the brine can drain from the waste so the solvent can contact the solids. The gas release rate thus depends on the brine pumping rate and is not accelerated by adding excess water above that which can contact the waste. If the solid matrix around the bubbles is not liquid saturated, pathways exist to the tank headspace and the gas generated in the waste is dissipated rapidly by diffusion (Stewart et al. 1996). Therefore, potentially hazardous gas releases can only occur by disrupting liquid-saturated solids.

Ammonia evaporation is not a factor during saltcake dissolution. Dilution accompanying dissolution reduces the ammonia evaporation rate significantly because the solubility of ammonia increases greatly as the dissolved salt concentration decreases with dilution. A water spray is also a very effective way to scavenge ammonia vapor from the headspace. This was clearly demonstrated during back-dilution in SY-101, where the headspace ammonia concentration decreased from around 400 ppm to less than 100 ppm in a matter of minutes after back-dilution began (Mahoney et al. 2000). Spraying water on the waste surface for retrieval inhibits ammonia release and probably reduces ammonia concentration below the historic baseline values.

3.4.2 Expected Gas Releases During Saltcake Dissolution

Based on the fact that dissolution and therefore gas release during dissolution are limited by the rate at which brine drains from the waste, a bounding gas release can be computed given the brine pumping rate. Water is assumed to be added at a rate Q_{WIN} such that the resulting brine production rate is equal to the brine capacity of the saltwell pump, Q_{BOUT} . Assuming that dissolution immediately releases all the gas stored in the waste being dissolved, the gas release rate, Q_{GAS} , is the product of the gas volume fraction, α , and the dissolution rate as follows:

$$Q_{GAS} = \alpha [F_{DSLW}/F_{BRINE}] Q_{BOUT} \quad (3.4.1)$$

where F_{DSLW} is the volume of original waste in which soluble solids are dissolved per unit volume of water added, and F_{BRINE} is the volume of total brine produced (water added plus dissolved solids plus interstitial liquid liberated) per unit volume of water added. Note that Eq. (3.4.1) describes the gas release rate from sediment in which the gas is retained at atmospheric pressure; i.e., there is no significant supernatant layer over the sediment.

The resulting release rates from typical tanks are quite low, especially when the brine is removed by the saltwell pumping system (see Section 3.3). For example, the solubility of U-107

waste is such that $F_{DSL} = 1.8$ and $F_{BRINE} = 2.15$ (Estey et al. 2001), and the gas volume fraction averages 0.17 at the 95% confidence level, as estimated with the BPE calculation and neutron log integration (Hedengren et al. 2001). The U-107 proof-of-concept demonstration planned to apply a maximum of 2,400 gallons of water per day during one eight-hour shift while the saltwell pump operated continuously. At the maximum water addition and assuming all of it dissolves waste, the water will create F_{BRINE} times its volume in brine, which requires an average pumping rate of 5,160 gal-brine/day, or 3.58 gpm. Eq. (3.4.1) indicates a corresponding gas release of 879 gal/day or 117 scfd.

The maximum daily water addition during the U-107 proof-of-concept demonstration was 1,757 gallons on February 1-2, 2003 (Baide et al. 2003). Unfortunately, the hydrogen sensor was considered to have failed as of January 12. After that date, industrial hygiene technicians took supplementary twice-daily readings with a combustible gas meter at times outside the water addition intervals. Thus, as reported in Baide et al. (2003), it was not possible to estimate the retained gas releases caused by dissolution during or after the proof-of-concept test.

The self-limiting behavior of gas release by dissolution makes the process controllable. Though some lag may be expected, the dissolution rate and gas release rate can theoretically be reduced relatively quickly by terminating pumping (accumulating brine forms a barrier to fresh solvent) and by shutting off the water spray (terminating the supply of fresh solvent). The most effective control is achieved if dissolution is local and the inventory of excess solvent is kept to a minimum. Though the actual response time of the gas release to changes in pumping and water addition rates are not known, saltwell pumping experience (see Section 3.3.2) indicates that it would be reasonable to expect gas release to lag cessation of pumping and spraying by about a day. However, the more local the dissolution, the shorter the lag time. These limitations indicate that saltcake dissolution should not produce much larger gas releases than those expected during saltwell pumping without concurrent dissolution.

The gas releases and waste behavior during the dissolution of the crust layer in SY-101 qualitatively validated our understanding of the physics of solvent flow and saltcake dissolution. The initial 120-inch crust layer in this tank was comparable to the non-supernatant waste in U-107 in thickness but contained almost twice the gas. The crust was dissolved in a series of three waste transfers and back-dilutions that eventually added 525,000 gallons of water. The bulk of the crust dissolution and gas release occurred in the second back-dilution, when the headspace hydrogen concentration peaked at about 3,000 ppm. This concentration represents a sudden release of about 200 scf of hydrogen, much more than can be expected during dissolution of SST waste.

3.4.3 Gas Monitoring Considerations During Saltcake Dissolution

The potential hazard for saltcake dissolution is not the gas release rate so much as the low passive ventilation rates that allow the gas to accumulate. The headspace hydrogen concentration at a given time, t , during saltcake dissolution can be computed by (Stewart 2001):

$$C_H(t) = \frac{Q_{GAS}\chi_H}{Q_A} + \left(C_0 - \frac{Q_{GAS}\chi_H}{Q_A} \right) \left(1 - \frac{Q_B}{V_{HS0}} t \right)^{\frac{Q_A}{Q_B}} \quad (3.4.2)$$

where

- C_0 = initial headspace hydrogen concentration
- V_{HS0} = initial headspace volume
- χ_H = hydrogen fraction of the retained gas in the waste

and

$$\begin{aligned} Q_A &= Q_{VIN} + Q_{GAS} \\ Q_B &= Q_{WIN} - Q_{BOUT} - Q_{GAS} \frac{P_A}{P_S} \end{aligned}$$

where

- P_A = pressure in the headspace atmosphere (~ 1 atm)
- P_S = pressure at which the gas is retained in the sediment
- Q_{GAS} = gas release rate defined by Eq. (3.4.1)
- Q_{VIN} = estimated headspace passive (or measured active) ventilation rate
- Q_{WIN} = solvent water flow rate
- Q_{BOUT} = brine pumping rate.

Figure 3.11 shows the headspace hydrogen concentration as a function of time obtained by applying Eq. (3.4.2) to the U-107 proof-of-concept example with higher brine pumping rates as a parameter, assuming the released gas is 50% hydrogen and the headspace ventilation rate is 2.5 cfm. The figure shows that the hydrogen concentration increases monotonically, even for the lowest pumping rates. The rate of rise is proportional to the pumping rate. At a 10-gpm pumping rate, it takes about six days to reach 25% of the LFL and almost 30 days at 3 gpm. However, at higher pumping rates, hydrogen concentration rises rapidly. At 100 gpm it takes 12 hours to reach 25% LFL and at 300 gpm less than 2 hours.

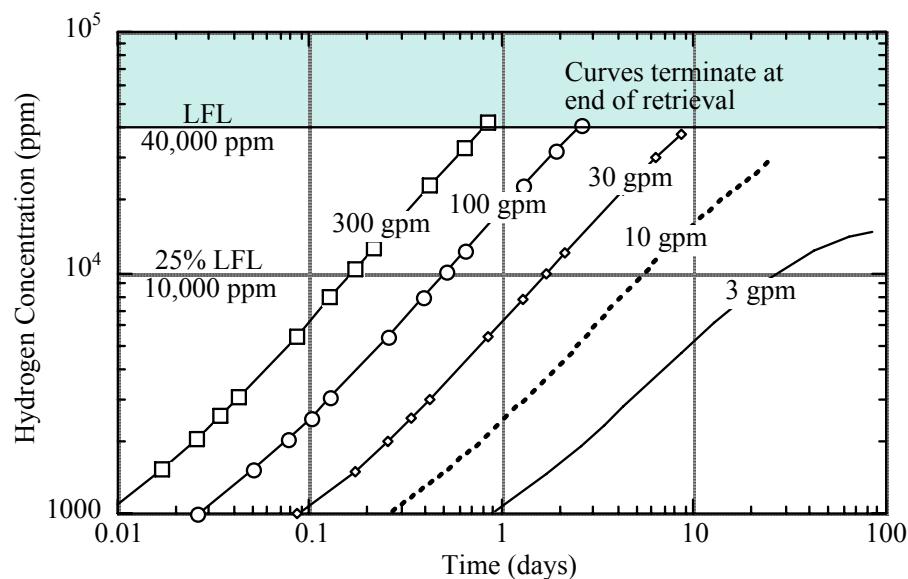


Figure 3.11. Hydrogen Concentration Versus Time for Saltcake Dissolution; Group B Tank

Because gas release during saltcake dissolution depends on having the brine continually removed, the associated gas release can be halted after some lag time by shutting down the pump and water spray system. To ensure that the system can be shut down well before reaching the LFL, the headspace hydrogen concentration must be monitored at an adequate frequency. If it is desired to shut down at 25% of the LFL, for example, the period between headspace samples should be no more than half the estimated time to reach 25% of the LFL. The required gas monitoring frequency must therefore be at least $2/t$ samples per unit time. The time required to reach a given hydrogen concentration, C_H , during saltcake dissolution can be calculated by solving Eq. (3.4.2) for t to give

$$t = \frac{V_{HS0}}{Q_B} \left[1 - \left(\frac{C_H - \frac{Q_{GAS}\chi_H}{Q_A}}{C_0 - \frac{Q_{GAS}\chi_H}{Q_A}} \right)^{\frac{Q_B}{Q_A}} \right] \quad (3.4.3)$$

For the saltcake dissolution example described above, the required gas monitoring frequencies defined as $f = 2/t$, with t defined by Eq. (3.4.3) for an action level of 10,000 ppm hydrogen, are plotted as a function of pumping and ventilation rates in Figure 3.12. The figure shows that the weekly monitoring required for saltwell pumping is adequate for the initial saltcake dissolution retrieval demonstration operations. With higher pumping rates, daily monitoring may be necessary.

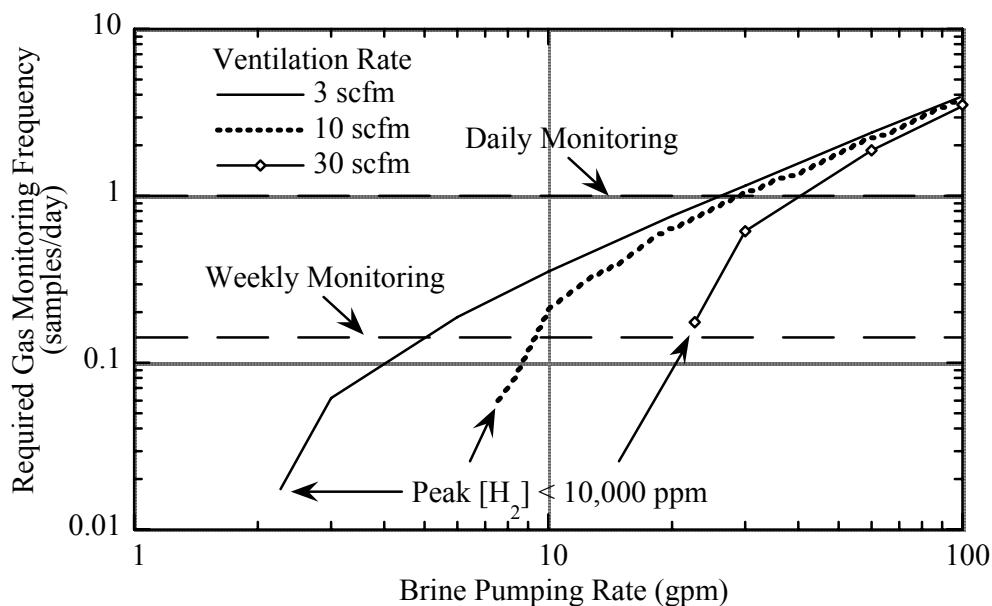


Figure 3.12. Required Gas Monitoring Frequency Versus Brine Pumping Rate

These monitoring requirements are based on the assumptions that dissolution (and therefore gas release during dissolution) is limited by the rate at which brine drains from the waste. Further, that the rate at which brine drains from the waste is quite low, and water is assumed to be added at a rate such that the resulting brine production rate is equal to the brine capacity of the transfer pump. For instances where rapid transfer rates may result in large drainage volumes, or when a large unsaturated water inventory may be result in dissolution, the analyses presented in Huckaby and Wells (2004) should be used.

3.4.4 Potential Changes in Waste Classification After Saltcake Dissolution

All SSTs that are successfully retrieved will move to Waste Group C because little waste will be left to generate and retain gas. Saltcake dissolution increases the headspace and reduces the consequences of gas releases even if not carried all the way to retrieval. Therefore, because the process cannot adversely affect a tank's gas retention and release behavior, no reevaluation of waste group is necessary.

3.4.5 Potential Changes in Hydrogen Generation After Saltcake Dissolution

Because saltcake dissolution removes waste from a tank, the post-retrieval hydrogen generation rate will be less than the initial rate. Therefore, the steady-state flammability status is bounded by the initial calculation, and no reevaluation is necessary.

3.5 Water Addition

Adding small volumes of water to both DSTs and SSTs is necessary from time to time to flush transfer lines, lance instruments into the waste, decontaminate hardware removed from the waste, or install retrieval systems. Larger water additions are possible and may be expected during retrieval operations. The global waste disturbance from water addition occurs via an increase in hydrostatic pressure and potentially by dissolution of soluble solids. Some water additions such as lancing may also be a local waste disturbance that releases a small additional gas volume. Large water additions can decrease the tank headspace, which adversely affects gas releases and may change the tank classification. Large water additions with concurrent or subsequent brine removal should be treated as SST saltcake dissolution, as described in Section 3.4. Gas release resulting from dissolution during water addition in SSTs where hydraulic mixing may be significant is bounded by Huckaby and Wells (2004). Water additions to SSTs are restricted to preclude BDGReS by prohibiting additions that result in an energy ratio greater than three (Kripps 2005).

3.5.1 Mechanisms for Gas Release During Water Addition

The dominant mechanism for gas release by water addition is dissolution of solids retaining gas. Gas release due to water addition is thus an issue only in tanks with some fraction of soluble solids. In saltcake SSTs, water added to the waste gradually invades interstitial liquid via diffusion and convection. As it contacts soluble salts, the water will eventually dissolve waste

that retains gas. The same gas release mechanism as saltcake dissolution would prevail but would act slower because brine is not removed to allow water to contact unsaturated waste.

In saltcake DSTs, added water would tend to first dissolve the floating crust layer, if present, releasing the gas contained therein in a manner similar to the SY-101 surface level rise remediation. If dissolving the crust does not saturate the liquid, convection would mix the dilute liquid throughout the supernatant, where it would eventually penetrate the sediment layer, dissolve solids, and release gas. However, this process (mixing and dissolution) is relatively slow and it is possible to add a large volume of water before significant dissolution can occur.

Water addition in the absence of significant dissolution lowers the average supernatant density, thereby raising the neutral buoyancy gas fraction. This and compression of the retained gas would significantly reduce the likelihood of BDGReS during the addition, even in Waste Group A tanks. Conversely, however, in addition to the gas release caused by dissolution of the gas-retaining matrix, buoyant displacement gas release events (BDGReS) may be triggered as solids are dissolved as explained below.

3.5.2 Expected Gas Releases During Water Addition

Water or dilute waste additions dissolve soluble solids in the sediment, releasing gas retained therein. The only additions that can be modeled conservatively in regards to dissolution without considering competing solubility equilibria (see Section 3.8) are additions of water or of waste that is much more dilute than the supernatant. In any event, water addition bounds the dissolution and thus the gas release that may occur.

Neglecting potential BDGReS, a maximum water addition volume can be calculated as the volume required to raise the headspace hydrogen concentration to a specified flammability limit (e.g. $[H_2]_{LFL} = 4 \text{ vol\%}$) with zero ventilation. The volume of water added to a tank that would release sufficient gas to bring the headspace to the LFL with zero ventilation can be expressed as

$$V_{max} = \frac{V_{HS}[H_2]_{LFL}}{\alpha X_H F_{DSLN}} \quad (3.5.1)$$

where V_{HS} is the headspace volume, α is the retained gas fraction, and X_H is the hydrogen fraction in the retained gas. The volume of waste in which soluble solids are dissolved and gas is released per volume of water added is denoted by F_{DSLN} . Noting that the post-addition headspace volume (neglecting changes due to dissolution and gas release) is given by the tank volume (V_T) minus the initial waste (V_{W0}) and added water (V_{max}) volumes, Eq. (3.5.1) may be written as

$$V_{max} = \frac{V_T - V_{W0}}{1 + \frac{\alpha X_H F_{DSLN}}{[H_2]_{LFL}}} \quad (3.5.2)$$

In a DST with a waste level of 422 inches, the waste volume is 1,160,094 gallons and the total tank volume 1,406,390 gallons, leaving a headspace of 246,296 gallons (Barker 2003). Assuming a water addition fills the tank to this level, for the conservative conditions of an average retained gas volume fraction of 0.2 at 100% hydrogen ($X_H = 1$), $[H_2]_{LFL} = 0.04$, and $F_{DSL} = 2$ gallons of waste releasing gas per gallon of water added, the maximum water addition to remain below the LFL is 24,630 gallons, according to Eq. (3.5.1). The initial waste level in this case would have been 413 inches.

For a DST with a waste level of 460 inches, the waste volume increases to 1,264,611 gallons, which reduces the headspace to 141,779 gallons (Barker 2003). With all other parameters as above, Eq. (3.5.1) gives the maximum water addition to remain below the LFL as 14,178 gallons.

The limiting water addition to 100 series SSTs to remain below the LFL from Eq. (3.5.1) is approximately 29,958 gallons with the waste parameters as specified above. The total tank volume and waste volume [design limit waste level is 281 inches (Roberts 2005)] for this limiting case are determined from Barker (2003). For the 200 series tanks with the same waste parameters and a design limit level of 285 inches (Roberts 2005), the maximum water addition is approximately 529 gallons.^(a) Again, total tank and waste volumes are determined from Barker (2003).

The DST upper-bound limit may not be sufficient if the dissolution triggers a BDGRE as discussed below. Depending on the circumstances, these potential BDGREGs may not be bounded by historical BDGRE release volumes. It is possible for a tank with a buoyancy ratio (BR) less than but approaching unity to experience a BDGRE due to dissolution. This possibility may be determined directly from the buoyancy ratio.

As presented in Section 2.3.2, Hanford data from DSTs that have historically exhibited BDGREGs indicates that gas fraction profiles in a sediment layer appear to approximate either a linear or parabolic profile over the sediment depth (Hedengren et al. 2000). A linear profile is expected to occur only in those tanks which regularly experience BDGREGs. In either the parabolic or linear profile case, dissolution of the sediment from the top down will remove the overburden of low gas sediment, thereby raising the average gas fraction in the remaining sediment. Depending on supernatant liquid density and hydrostatic pressure effects, buoyant conditions may be reached as a result of this overburden removal, thereby triggering a BDGRE. However, in order for any portion of the sediment to become buoyant as a result of the removal of an overburden, the maximum gas fraction in the layer must be at least equal to neutral buoyancy. For tanks with $BR \geq 1$, i.e., those that exhibit BDGRE behavior, this condition exists. The following analysis therefore applies to tanks with $BR < 1$, indicating that the expected gas fraction profile is parabolic.

(a) The reported water addition limit for the 200 series SSTs is lower than may be determined from Barker (2003) due to correction of the 200 series total tank volume. Personal communication from SA Barker, CH2M HILL, Richland, WA.

Differentiating the parabolic gas fraction profile of Eq. (2.11) with respect to non-dimensional elevation in the sediment, η , shows that the maximum gas fraction occurs at $\eta = 0.5$, or

$$\alpha_{\max} = \frac{4}{3} \frac{C}{\rho_s} \left(\frac{GT}{p} \right)^{\frac{1}{3}} H_s^2 \quad (3.5.3)$$

where ρ_s is the sediment density, G is the molar gas generation rate per unit volume, and p is the average hydrostatic pressure in the sediment layer. The coefficient C is adjusted so that the minimum ratio of the average steady-state gas fraction to the neutral buoyant gas fraction in the historic BDGRE tanks is unity. Substituting the buoyancy ratio, Eq. (2.13), Eq. (3.5.3) becomes

$$BR = \frac{\frac{3}{4} \alpha_{\max}}{\alpha_{NB}} \quad (3.5.4)$$

where α_{NB} is the neutral buoyant gas fraction. Thus, for $1 > BR \geq 0.75$, the maximum gas volume fraction will exceed neutral buoyancy and a portion of the sediment will be buoyant. Also, as noted above, buoyant conditions may always be reached in a tank with $BR \geq 1$. Therefore, the maximum water addition value determined from Eq. (3.5.2) applies to any tank with $BR < 0.75$.

When computing the buoyancy ratio for this criterion, it is most conservative to set the liquid density to the maximum of either the original liquid or the added liquid. Further, a BDGRE is precluded in tanks with an energy ratio (computed for tank conditions after the addition) less than three ($ER < 3$). Note also that, at steady state after water addition/dissolution, a tank with a $BR < 1$ will not experience BDGREs.

In tanks where BDGREs are already occurring ($BR \geq 1$), a gas release may occur at any time. A BDGRE may also be triggered by water addition as the non-buoyant overburden dissolves as in a non-BDGRE tank as described above. These induced BDGREs would likely be of lower volume than their historic behavior for the following reasons: 1) some of the gas will have been released in dissolving the overburden, 2) the probability of inducing secondary releases is lower because the liquid density is reduced by the water and dissolving the overburden leaves less sediment to be suspended, and 3) the secondary releases would be smaller via reason (1).

If we assume that a BDGRE of the maximum historical volume V_{GRE} (at 1 atm) occurs just after a water addition, Eq. (3.5.2) can be modified to include the effect of induced (or spontaneous) BDGREs on the maximum water volume to remain below the LFL as follows:

$$V_{\max} = \frac{V_T - V_{W0} - V_{GRE} \frac{X_H}{[H_2]_{LFL}}}{1 + \frac{\alpha_{NB} X_H F_{DSLN}}{[H_2]_{LFL}}} \quad (3.5.5)$$

Eq. (3.5.5) ignores the small increase in headspace volume resulting from the gas release (maximum of 1.4% in AN-105; conditions from Wells et al. (2002)). The increase in α_{NB} caused by the water addition is accounted for by using a conservatively high value of 0.2 (the actual maximum α_{NB} after water addition is 0.17 in AN-103 at conditions from Wells et al. (2002); 0.2 is used for direct comparison of the results to Eq. (3.5.1)). Table 3.4 lists the maximum BDGRE volume and tank conditions from Hedengren et al. (2000) and the maximum water addition to approach LFL from Eq. (3.5.5), assuming that F_{DSL_N} remains at 2 and that $[H_2]_{LFL}$ is 0.04. The limiting water addition is much lower in AN-105 due to its relatively large maximum BDGRE volume and small initial headspace.

Table 3.4. Peak Hydrogen Concentrations with Dissolution After Water and Caustic Addition

Event	AN-103	AN-104	AN-105	AW-101	SY-103
Maximum BDGRE volume (m^3)	4.6	18.7	28.9	19.2	20.8
Initial headspace (m^3)	1,697	1,380	1,051	1,061	2,490
Retained gas hydrogen fraction	0.61	0.45	0.60	0.32	0.54
Water volume added for LFL (gal)	12,800	11,200	4,900	12,100	19,300

Though tanks with a buoyancy ratio between 0.75 and 1.0 can theoretically experience an induced BDGRE by dissolution as discussed above, it is not possible to apply Eq. (3.5.5) because these tanks have no gas release history to define a release volume. Nevertheless, the behavior of the BDGRE tanks shows that release size is roughly proportional to the buoyancy ratio. The upper bound of the 95% confidence interval of the best fit of the buoyancy ratio to the 95th percentile of all historic BDGREs indicates that the maximum expected hydrogen concentration at a buoyancy ratio of one is less than 15,000 ppm (Stewart et al. 2005). Thus the likelihood of large dissolution-induced BDGREs is very low.

3.5.3 Gas Monitoring Considerations During Water Addition

Water additions less than the upper-bound volume (Eq. 3.5.2) in tanks with $BR < 0.75$ or $ER < 3$ should not require monitoring. In Waste Group-B DSTs that have $BR \geq 0.75$ and $ER \geq 3$, the likelihood of a dissolution-induced BDGRE threatening flammability is low. For water additions that exceed the upper-bound volume, active ventilation and flammable gas monitoring will be required; however, no monitoring frequency or duration can be provided. Water additions to Waste Group-A DSTs should be limited to prevent flammable conditions if a large BDGRE were to occur simultaneously with the dissolution-induced gas release. Additions larger than the threshold volume in SSTs with passive ventilation should be treated as saltcake dissolution with the monitoring frequency defined by Eq. (3.4.2).

3.5.4 Potential Changes in Waste Classification After Water Addition

Water addition decreases the headspace in a tank and could potentially move a tank from Group C into Group B or Group A, if it were already exhibiting spontaneous BDGREs, and requires a waste group evaluation. If dissolution is neglected or minimal, water additions reduce the mixed supernatant density and increase the neutral buoyancy gas fraction, decreasing the buoyancy ratio. With dissolution of sediment, the buoyancy ratio will be reduced by the

decreased sediment depth. Thus it is not possible for water addition to cause a tank with a buoyancy ratio initially less than one to have $BR \geq 1$.

In SSTs, however, a large water addition could create a supernatant layer sufficiently deep to make the energy ratio defined by Eq. (2.18) exceed the threshold value. In DSTs, this would require a Group B tank to be evaluated for BDGRE potential. However, the buoyancy ratio criterion (see Section 2.3.2.1) cannot be applied to SSTs because their waste properties and configuration differ considerably from those on which the model is based. To avoid this situation, water additions to Group B SSTs should be limited such that the energy ratio defined by Eq. (2.18) remains less than the threshold. Solids dissolution should be estimated in this calculation to account for all the liquid available.

By decreasing the headspace, raising the neutral buoyancy gas fraction, and increasing the hydrostatic pressure on the retained gas, water additions may raise the peak hydrogen concentration from ongoing spontaneous BDGREs in Group A DSTs. Though water also dissolves solids and thus tends to reduce BDGRE potential and size at steady-state, the former effects dominate unless a very large volume of water is used.^(a) However, the increase in peak hydrogen concentration resulting from spontaneous BDGREs after a water addition with no dissolution can be estimated easily.

The headspace hydrogen concentration, C_H , following a BDGRE caused by water addition with no dissolution can be conservatively calculated from the gas release volume, V_R , the fraction of hydrogen in the released gas, χ_H , the headspace volume, V_{HS} , and the initial concentration, C_0 , assuming an instantaneous release (Meyer and Wells 2000):

$$C_H = \chi_H \frac{V_R}{V_{HS}} + C_0 \quad (3.5.6)$$

For a given gob volume, V_0 , in the sediment (see Section 2.3.2), the volume of gas released into the headspace when the gob goes buoyant can be expressed as

$$V_R = \alpha_{NB} \Phi \frac{P_S}{P_A} V_0 \quad (3.5.7)$$

where

- α_{NB} = neutral buoyancy gas fraction
- Φ = gas release fraction
- P_S = hydrostatic pressure on the gas in the sediment layer
- P_A = ambient pressure in the tank headspace.

The neutral buoyancy model for the gas release fraction states that gas is released until the gob returns to neutral buoyancy at the waste surface. This results in the expression

(a) A calculation of the water dilution volume required to dissolve sufficient solids to remediate SY-101 and the results of this operation are summarized by Johnson et al. (2000).

$$\Phi = 1 - \frac{P_A}{P_S} \quad (3.5.8)$$

Combining Eq. (3.5.6), (3.5.7), and (3.5.8) and ignoring the low initial hydrogen concentration gives the overall expression for the post-GRE hydrogen concentration:

$$C_H = \alpha_{NB} \chi_H \left(\frac{P_S}{P_A} - 1 \right) \frac{V_0}{V_{HS}} \quad (3.5.9)$$

Assuming that the hydrogen fraction and gob volume do not change after water addition, the peak hydrogen concentration due to a spontaneous BDGRE after water addition is related to the historic value by the ratio for the post- and pre-addition states using Eq. (3.5.9); that is,

$$C_{H2} = C_{H1} \frac{\alpha_{NB2}}{\alpha_{NB1}} \frac{\left(\frac{P_{S2}}{P_A} - 1 \right)}{\left(\frac{P_{S1}}{P_A} - 1 \right)} \frac{V_{HS1}}{V_{HS2}} \quad (3.5.10)$$

Table 3.5 shows the results of applying Eq. (3.5.10) to five DSTs that historically experienced spontaneous BDGREGs (tank waste conditions from Hedengren et al. 2000, Wells et al. 2002). The table lists the maximum headspace hydrogen concentration for any BDGRE recorded on a SHMS (Hedengren et al. 2000) and the estimated concentration had the tank been filled to 422 inches by adding water at a density of 1,000 kg/m³ with no dissolution or adding caustic (discussed in Section 3.8) at a density of 1,500 kg/m³. No dissolution or precipitation is considered from the caustic addition. It is hypothetically possible to approach the LFL in SY-103 only after adding about 380,000 gallons of water. However, because it has been calculated that about 250,000 gallons of water would dissolve enough solids to prevent BDGREGs altogether in SY-103, this result is plausible only under the conservative assumption of no dissolution.^(a)

Table 3.5. Peak Hydrogen Concentrations Without Dissolution After Water and Caustic Addition

Event	AN-103	AN-104	AN-105	AW-101	SY-103
Maximum historic BDGRE [H ₂] (ppm)	1,600	6,100	17,000	6,200	5,100
Volume added for 422 in. (gal)	203,000	108,000	33,000	35,000	380,000
Peak [H ₂] after water addition (ppm)	5,000	14,000	23,000	8,400	49,000
Peak [H ₂] after caustic addition (ppm)	3,500	9,200	20,000	7,100	19,000

(a) Stewart CW and BE Wells. September 2000. *A Strategy for Remediation and Return to Service for Tank 241-SY-103*. Letter Report TWS01.01, Pacific Northwest National Laboratory, Richland, WA.

Further, as shown in Section 3.2.4, the addition of liquid at a lower specific gravity than the existing supernatant will reduce the buoyancy ratio. For water addition to SY-103 with no dissolution, the buoyancy ratio drops below unity with approximately 220,000 gallons added. AN-103 is a special case (it is the BDGRE tank used to set the buoyancy ratio coefficient, as described in subsection 2.3.2.1), and any water addition with no dissolution is sufficient to reduce the buoyancy ratio to below unity. It is not possible, however, to determine the specific volume of water needed to prevent BDGREGs though some larger volume would clearly do so.

With no dissolution or precipitation, adding water increases the peak hydrogen concentration more than adding caustic in all tanks. We conclude that, though adding water exacerbates the consequences of BDGREGs in Group A tanks, the total effect is not enough to approach the LFL based on historical behavior and, in sufficient quantities, may mitigate the tank.

3.5.5 Potential Changes in Hydrogen Generation After Water Addition

Water additions tend to dilute the total organic carbon concentration in the liquid but increase the liquid volume, essentially maintaining the total generation rate. However, water additions increase the contribution from water radiolysis due to reduction of the salt concentration and salt-scavenger effect. Therefore, water addition can increase the total generation rate.

3.6 Mixer Pump Operation

Mixer pump operation is capable of producing the largest gas releases of all the activities discussed in this report. Mixing the waste can release its entire gas inventory and make the headspace flammable in Group A and B tanks. Because of this, the initial period of mixing must be planned specifically to limit the volume of waste disturbed and control the rate of gas release. Only after this degassing period can the mixer pumps be operated to their design performance. The rest of this section is devoted mainly to the degassing period.

Mixer pumps are planned to be installed as part of DST waste retrieval and staging to the vitrification plant. A mixer pump degassed the waste in SY-101 in 1993, and a pair of powerful pumps was operated more recently in AZ-101. Mixer pump operation is specific to DSTs.

The global waste disturbance of a mixer pump is the hydraulic erosion and suspension of previously settled sediment by the pump jets. The rate of disturbance depends on the pump power, jet diameter and configuration, and the duration of the operation. Thus a single run of a small pump for a short time might be classified a local waste disturbance. But mixer pump operation is considered a global disturbance because it is designed to generally mix the tank.

3.6.1 Mechanisms for Gas Releases During Mixer Pump Operation

Gas release during mixer pump operation is caused mainly by the disruption of the waste structure retaining the gas bubbles. However, a major secondary release mechanism is the increase in effective liquid density by suspension of solid particles. This reduces the neutral buoyancy gas fraction, which may induce BDGREGs. Planning a degassing campaign thus

involves limiting the amount of waste disturbed in each pump run and including the effects of induced BDGReS. After degassing is completed, periodic mixing will continue to release gas that has accumulated between pump runs by detaching bubbles from sediment particles on which they nucleated and grew.

If a liquid waste surface with a high concentration of dissolved ammonia is stirred or agitated, the evaporative mass-transfer rate of ammonia can be greatly increased. While mixing generally stirs the waste, it does not tend to produce significant ammonia evaporation. Most of the mixing action affects the region of the waste near the tank bottom and does not appear to stir the waste surface enough to boost evaporation very much (Mahoney and Stewart 2002). Headspace ammonia concentrations in SY-101 remained below 200 ppm during the intense mixing period in March 2000 and increased only slightly during pump runs (Mahoney et al. 2000). The two 300-hp mixer pumps planned for DST waste feed delivery potentially can disturb waste at four times the rate of the single 150-hp pump in SY-101, with a correspondingly greater potential for increased ammonia evaporation. Nevertheless, ammonia releases remain more a toxicological concern than a flammability hazard.

3.6.2 Expected Gas Releases During Mixer Pump Operation

The rate and amount of gas released depends on the rate and volume of waste disturbed and the gas volume fraction retained. It is theoretically possible for multiple mixer pumps to release enough gas to bring the headspace hydrogen concentration to the LFL if they are started up and run at full power in Waste Group A or B tanks.^(a) A specific degassing plan is needed in these tanks that disturbs the waste a little at a time until most of the retained gas has been released.

After degassing is complete, however, periodic mixing can only release gas that has accumulated between pump runs. This is of no concern unless the pump has been idle for a long time. Table 3.6 shows the time required for DSTs AN-103, AN-104, AN-105, and AW-101, which currently experience BDGReS, to accumulate enough gas to make the headspace flammable (if possible) and to achieve neutral buoyancy. This assumes that mixing is terminated after the tank is initially completely degassed while the volumetric gas generation from Barker and Hedengren (2004) is 100% retained in the sediment. The table shows that these tanks would reach Waste Group B status in 1.1 to 3 years, and would reach neutral buoyant conditions after 3 to 14 years.

Table 3.6. Time to Return to BDGRE Behavior After Degassing

Tank	Years to Reach Waste Group B	Years for Neutral Buoyancy
AN-103	3	14
AN-104	1.1	3
AN-105	1.3	7
AW-101	2	3

(a) As described in Section 2.4.1, Waste Group A and B tanks store sufficient gas to make the headspace flammable if all of it were suddenly released; Waste Group C tanks do not. Waste Group A tanks exhibit spontaneous BDGReS, while tanks in Waste Group B do not.

In Waste Group A tanks, the sediment suspended in the supernatant by mixing increases the effective supernatant density and the neutral buoyancy gas fraction, inducing BDGReS. These induced gas releases are superimposed on the gas released as a result of mixing. Such gas releases will also be induced in Group C tanks that experience BDGReS, but they are of minor consequence because these tanks do not contain sufficient gas to make a well-mixed headspace flammable.

Borderline Group B tanks may also experience induced BDGReS during mixing. The analysis derived in Section 3.2 for BDGReS induced by waste addition can be applied to determine whether mixing will induce a BDGRE. A bounding liquid density can be calculated by assuming that the entire sediment layer is mixed with the initial supernatant as follows:

$$\rho_{L2} = \frac{\rho_{L1}H_L + \rho_S H_S}{H_L + H_S} \quad (3.6.1)$$

where ρ_{L1} is the initial supernatant density, ρ_S and H_S are the sediment density and depth, respectively, and H_L is the supernatant depth. This assumption about the density makes the analysis truly bounding. If the entire sediment layer is suspended in the supernatant, mixing is complete and no sediment layer remains to experience a BDGRE.

The final buoyancy ratio is calculated from the gas volume fraction and the neutral buoyancy gas fraction, including the effects of the increased density. If the final buoyancy ratio is greater than 1, a BDGRE may occur during mixing. The final gas volume fraction, accounting for compression by the increased density, is given by Eq. (3.1.2). Combining this with Eq. (2.12) for the final neutral buoyancy gas volume fraction provides an expression for the final buoyancy ratio:

$$BR_2 = \frac{\alpha_2}{\alpha_{NB2}} = \frac{\alpha_1 \rho_S / (\rho_S - \rho_{L2})}{\alpha_1 + (1 - \alpha_1) P_2 / P_1} \quad (3.6.2)$$

where the pressure ratio is calculated by Eq. (3.2.4) using the new liquid density from Eq. (3.6.1) and a constant liquid depth. Substituting the product of the initial buoyancy ratio (Eq. 2.13) and the initial neutral buoyancy gas fraction (Eq. 2.12) for the initial gas fraction yields

$$BR_2 = \frac{\rho_S / (\rho_S - \rho_{L2})}{1 + \left[\frac{\rho_S}{BR_1(\rho_S - \rho_{L1})} - 1 \right] P_2 / P_1} \quad (3.6.3)$$

Applying Eq. (3.6.3) to a hypothetical Group B DST implies BDGReS could be induced during mixing. While very conservative, this result suggests that induced gas releases larger than those caused by the waste disturbance alone are possible in some Group B tanks. Though BDGReS can also be induced in some Group C tanks, they are not of concern because these tanks contain insufficient gas to reach flammability.

Wells et al. (2002) performed detailed analyses of these effects for DSTs AN-103, AN-104, AN-105, and AW-101. The analyses were based on a disturbance of a specified fraction of the sediment each pump run with one run every 8 or 24 hours, continuing until the entire sediment layer was mixed. Gas releases from induced BDGReS are also included via Monte Carlo simulation using essentially the same model described for waste removal in Section 3.1. The base case assumes that 5% of the sediment volume is disturbed every 24 hours. Other combinations were a 5% disturbance every 8 hours to investigate the sensitivity to pump schedule and a 20% disturbance every 8 or 24 hours to test the effect of increasing disturbance. The headspace ventilation rate was 120 scfm in AW-101 and 100 scfm in the other tanks.

The results are summarized in Table 3.7, which lists peak hydrogen concentrations predicted during the mixing process and the fraction of gas that was released by induced BDGReS. The results are consistent with the supernatant decant analyses. If the volume of waste disturbed per mixer pump run is limited to no more than 5% of the sediment every 24 hours, the peak hydrogen concentrations remain well below the LFL, including the effect of induced BDGReS.

At the 95% confidence level, the peak hydrogen concentration during mixing was predicted to be 15,100 ppm for AN-105 and 10,600 ppm for AN-104. In both of these tanks, about 25% of the gas release was due to BDGReS at the 95% confidence level. Increasing the disturbance to 20% every 24 hours raised the peak hydrogen concentration at the 95% confidence level to 28,400 ppm and 18,500 ppm in AN-105 and AN-104, respectively. However, the fraction of gas released by BDGReS remained at about 25% in both tanks. Increasing the frequency to a 5% disturbance every 8 hours in AN-105 did not change the result appreciably.

Table 3.7. Summary of Mixing Gas Release Analysis Results

Tank	Run	Quantity	Median	95% CL
AN-105	5% every 24 hours	Peak hydrogen concentration (ppm)	7,500	15,100
		Fraction gas release by BDGReS	0.09	0.23
	20% every 24 hours	Peak hydrogen concentration (ppm)	20,300	28,400
		Fraction gas release by BDGReS	0.08	0.22
	5% every 8 hours	Peak hydrogen concentration (ppm)	9,800	17,100
		Fraction gas release by BDGReS	0.09	0.23
AN-104	5% every 24 hours	Peak hydrogen concentration (ppm)	5,500	10,600
		Fraction gas release by BDGReS	0.12	0.26
	20% every 24 hours	Peak hydrogen concentration (ppm)	12,900	18,500
		Fraction gas release by BDGReS	0.10	0.25
AN-103	5% every 24 hours	Peak hydrogen concentration (ppm)	15,300	25,800
		Fraction gas release by BDGReS	0.20	0.33
AW-101	5% every 24 hours	Peak hydrogen concentration (ppm)	1,800	2,700
		Fraction gas release by BDGReS	0.00	0.14
	20% every 8 hours	Peak hydrogen concentration (ppm)	6,900	8,800
		Fraction gas release by BDGReS	0.00	0.13

AW-101 showed a very low hydrogen concentration for the base case, reaching only 2,700 ppm at the 95% confidence level with only 14% released by BDGReEs. The median result showed no BDGReEs at all. A more aggressive simulation was performed with 20% of the sediment disturbed every eight hours. This raised the peak hydrogen concentration to 8,800 ppm at the 95% confidence level with 13% of the gas released by BDGReEs. Using the proposed pumps, operation of one mixer pump in AW-101 at a fixed azimuth for 14 minutes creates approximately a 5% disturbance (Wells et al. 2002).

About the same result was predicted for AN-103 with a 5% disturbance every 24 hours and AN-105 with a 20% disturbance every 24 hours. At the 95% confidence level, peak hydrogen concentration was 25,800 ppm with 33% of the gas resulting from induced BDGReEs. The median for hydrogen concentration in AN-103 was 15,300 ppm; 20% of the gas was released by BDGReEs. This matches predictions at the 95% confidence level for the same case in AN-105.

That none of the four tanks analyzed exceeded the LFL during degassing should not be taken to imply that no limitations are needed on mixer pump operations. The degassing schedules studied were chosen specifically as possible plans that could be accomplished safely. More aggressive mixer pump operation might, and unlimited mixer pump operation would, cause the tank headspace to become flammable. A carefully considered degassing plan is absolutely necessary in these tanks or any DST in Waste Group A or B.

3.6.3 Gas Monitoring Considerations During Mixer Pump Operation

The results in Table 3.7 show that, though BDGReEs may be induced, most of the total gas released from a tank during mixing is produced by the action of the mixer pump. The gas release rates from both processes are high, and continuous monitoring is required to follow the hydrogen concentration transient.

No “stop-start” simulation was run for the mixer pump simulations as it was for supernatant decant. However, because there is a considerable lag (up to an hour or more) between mixer pump startup and peak gas release, a nominal mixer pump run might be completed before the hydrogen concentration rises significantly. Therefore, terminating mixer pump operation at the action level is not expected to be very effective at controlling gas release. On the other hand, delaying pump restart until after the hydrogen concentration falls back below a threshold value (500 ppm was used during mixing in SY-101) could have a measurable mitigating effect. Also, because of the delay and the potential for secondary induced BDGReEs occurring even later, gas monitoring should be continued for some time after each pump run during degassing to ensure that the peak hydrogen concentration has been observed.

Active ventilation of 100 scfm was specified for the simulations described in Section 3.6.2. Because mixing releases essentially all of the retained gas in a tank compared with 25 to 33% released by supernatant decant (see Section 3.1.2), the peak hydrogen concentration is more sensitive to the ventilation rate. Figure 3.13 shows the effect of varying the ventilation rate from 10 to 200 scfm in two of the Monte Carlo simulations (representing the same degassing operation but with altered input parameters) of the base mixing case in AN-103 (5% disturbance

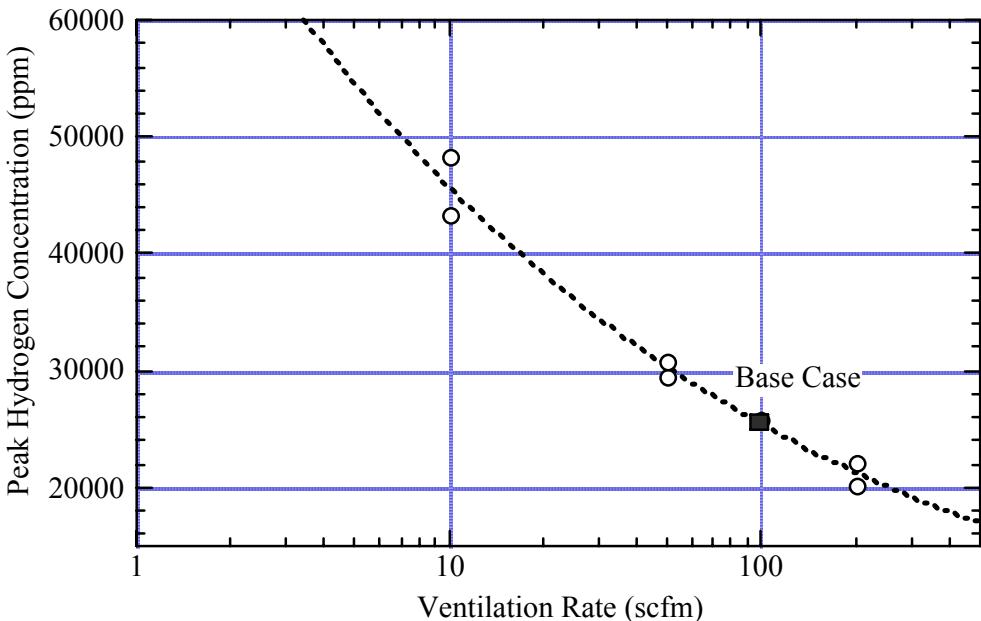


Figure 3.13. Sensitivity of Hydrogen Concentration to Ventilation in AN-103 During Mixing

every 24 hours) that produced a peak hydrogen concentration of 25,800 ppm at the 95% confidence level. The curve fit through the results shows that the LFL would have been exceeded for ventilation rates less than about 20 scfm. Increasing the ventilation rate to 500 scfm would have reduced the peak hydrogen concentration to about 18,000 ppm, a 30% reduction.

After the tank is mixed, essentially all of the retained gas will have been released. Further mixer pump runs will not release large gas volumes, and no further BDGReS will be induced. Hydrogen concentrations will remain within a few hundred ppm of the baseline, and monitoring is not necessary as long as the mixer pump runs often enough to prevent gas retention. As presented in Section 3.6.2, it would take at least one year for a tank to accumulate enough gas to be classified in Waste Group B, which would require a degassing plan for mixing.

Gas monitoring is necessary in these tanks and in any DST in Waste Group A or B. No formal basis exists to calculate how long the monitoring period after pump operation should be, but 24 hours is a reasonable time based on the experience with SY-101 mixing and the mixing analysis presented in Section 3.6.2 (Alleman et al. 1994; Wells et al. 2002).

3.6.4 Potential Changes in Tank Classification After Mixing

In the long term, assuming that mixing is eventually terminated and the tank allowed to return to a steady state, no changes in waste configuration result from mixer pump operation, and tank classification will not change.^(a) However, if mixer pump operation continues frequently

(a) The effect of “fluffing,” where the sediment layer is expanded while the previously mixed sediment gradually compacts to its former density, is discounted as a transient phenomenon compared with the minimum of three years before gas could accumulate to buoyancy (see Section 3.6.2).

enough to keep gas from accumulating, a tank can be maintained indefinitely in Waste Group C, as was done with SY-101 for many years. Because mixing cannot move a tank into a higher waste group, classification need not be reevaluated due to mixer pump operation. However, if a tank is reclassified as Waste Group C on the basis that mixing prevents gas accumulation, continued mixer pump operation is required to keep the tank in Group C. See Table 3.6 for the effects of terminating mixing in current BDGRE tanks.

3.6.5 Potential Changes in Hydrogen Generation After Mixing

Because no waste is added or removed by mixing, the total hydrogen generation does not change, and no reanalysis of steady-state flammability is needed. However, because hydrogen generation is sensitive to the waste temperature, the effect of mixer pump power deposition in the waste should be considered for long-term operation.

3.7 Airlift Circulator Operation

Airlift circulators (ALCs) were designed into the construction of DSTs AY-101, AY-102, AZ-101, AZ-102, AN-107, and AW-102 as well as SSTs in several tank farms. ALCs mix the waste by introducing a stream of air bubbles into 30-inch-diameter cylindrical tubes that extend from near the tank bottom well up into the supernatant. The bubble stream reduces the average density inside the circulator tubes, causing an upward flow.

In the AY and AZ aging waste tank farms, each tank has 22 ALCs installed to prevent temperature excursions caused by settling of heat-generating fission products. AN-107 has 21 ALCs of the same design as those in the aging waste tanks designed to “gently agitate the waste.” However, they were never connected to an air supply, and there are no plans to operate them in the future. AW-102 has only two ALCs that are much smaller (16 and 24 inch diameter) than in the other tanks that were intended to provide uniform feed for the 242-A Evaporator.

The global disturbance induced by ALC operation is similar to that of mixing except that the intense hydraulic action of a high-velocity jet is absent. However, the simultaneous operation of many large-diameter ALCs, albeit at a relatively low velocity, tends to make up in volume for their lack of intensity.

3.7.1 Mechanisms for Gas Releases During ALC Operation

The highest gas release potential occurs during startup of ALCs. Establishing flow into ALCs with their bases buried in the waste requires mobilizing a large volume of gas-bearing sediment. This could release a large fraction of the retained gas volume. However, as with mixer pump operation, after the initial degassing period, ALC operation will not release gas unless the ALCs have been idle for a long period, allowing gas to re-accumulate. As with mixing, the ALCs are designed to suspend sediment in the liquid, which reduces the neutral buoyancy gas fraction and could induce BDGREs.

Stirring or agitating a liquid waste surface with a high concentration of dissolved ammonia greatly increases the evaporative ammonia mass transfer rate. The action of gas bubbles rising continuously in the ALCs is an ideal process for stripping ammonia from the liquid into the bubbles. The large ammonia releases during transfers of SY-101 waste into SY-102, which created the highest ammonia concentrations recorded by gas monitoring equipment, apparently involved gas bubble sparging of the waste stream in a vacuum break (Mahoney et al. 2000). This indicates that high headspace ammonia concentrations can be expected during ALC operation in waste with a high concentration of dissolved ammonia.

Figure 3.14 is a schematic of a typical ALC in the AY and AZ farms. The left panel of the figure depicts ALC operation within a sediment layer, and the right panel gives details of the design. The ALC diameter, D_o , is 30 inches, and the base sits at height H_o , 30 inches above the tank bottom. The length, L , of 15 ALCs is 22 ft, but five are 17 ft long. The air line has a diameter, D_i , of 6 in. that tapers down to a 1-inch pipe, terminating in a sparger that consists of four triangular notches cut into the end of the pipe. Each of these triangular openings has a hydraulic diameter, d_e , of approximately 0.33 in. Airflow can be adjusted from 0 to 20 scfm in each ALC. Figure 3.15 shows the arrangement of ALCs in the AY and AZ tanks. One ALC is placed at the center, seven on a radius of 14.5 ft, and 14 at a 27-ft radius, for a total of 22 in each tank.

To estimate the potential for gas releases induced by ALC operation, simple models were developed for airlift pump theory and mobilization of sediments. These models are applied to conditions representing Tank AZ-101 during the late 1999 and mid-2000 ALC operation and

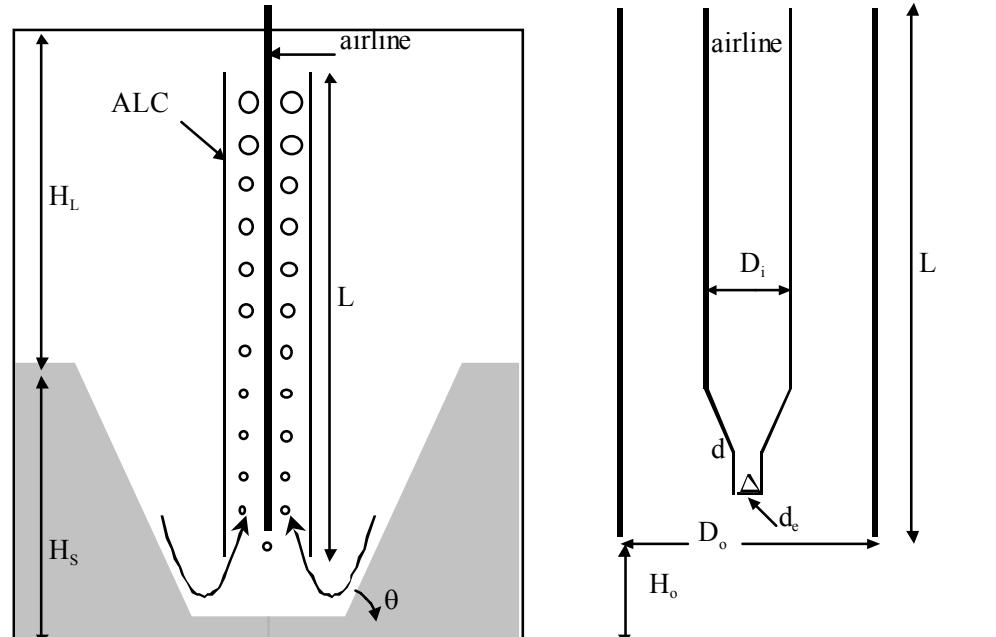


Figure 3.14. Schematic of Typical Airlift Circulator Operation

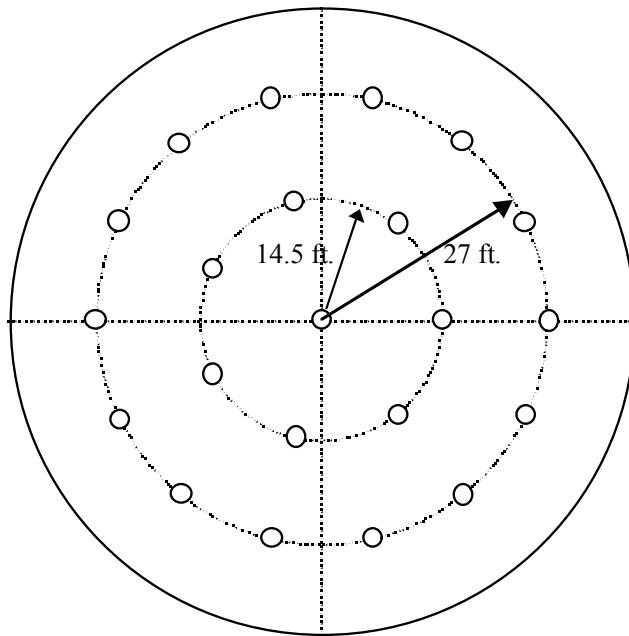


Figure 3.15. Location of ALCs in AY and AZ DSTs

hypothetical Group A and Group B tanks.^(a) None of the tanks that have historically exhibited BDGReS have ALCs. Thus, a Group A type tank is only considered should one of the tanks with operable ALCs be placed in Waste Group A.

3.7.2 Expected Gas Releases During ALC Operation

A sediment layer cannot be mixed and retained gas will not be released unless flow is established in the ALC. However, a sediment layer around the base of the ALC prevents flow unless the lithostatic pressure can be overcome and the material can be yielded by the pressure difference across the ALC tube. Anecdotal evidence is that the majority of the air lines are plugged in both AY and AZ tank farms, and water pressure in excess of 150 psi could not remove the plug. Even if the air lines are clear, the sediment may be too strong to allow circulation. Gas bubbles may escape around the outside of the ALC, or bubbles rising through the sediment inside the ALC may fail to mobilize the sediment therein. Seventeen ALCs were run during mixer pump operations in AZ-101.

Assuming that the sediment strength is a few hundred Pa, the startup pressure difference produced by 20 cfm of airflow should be more than adequate to yield the sediment and initiate flow. Also, bubbles generated at the sparger should preferentially rise through any sediment inside the ALC rather than bypassing around the outside.

(a) As described in Section 2.4.1, Waste Group A and B tanks store sufficient gas to make the headspace flammable if all of it were suddenly released; Waste Group C tanks do not. Waste Group A tanks exhibit spontaneous BDGReS, while tanks in Waste Group B do not.

If ALC pumping can be initiated, the resulting flow should mobilize the sediment in the region of the ALC and release the stored gas. The likely evolution of waste disturbance due to a single ALC in operation is illustrated in Figure 3.16. When the ALC first starts up, it is likely that supernatant will be drawn from around its base and sides (frame 1 in Figure 3.16). Agitation from the initial startup process would likely create a loose slurry inside the ALC. Additionally, air bubbles that escaped and rose to the surface along the outside of the ALC would create paths to the supernatant layer for liquid flow. This situation would be unstable, however, and the weak sediment would begin to collapse around the outside. During this process (frame 2), entrained sediment would be pumped upward and discharge at the outlet, releasing gas in the process.

Sediment would continue to slough off and be entrained toward the inlet, and a stable configuration could eventually be attained (frame 3). A maximum angle of repose, θ , exists for a given sediment strength that is stable. The maximum angle of repose is discussed further below. If the fluid velocities are high enough, fluid shear will continue to erode the sediment and the mobilized volume will increase. Finally (frame 4), the volume will become large enough and flow velocities reduced enough that fluid shear cannot erode any more material. This represents the maximum volume that could be disturbed by a single ALC. The size of the disturbed volume will depend on the sediment properties (strength, repose angle, critical shear stress for erosion, etc.) and the fluid velocities.

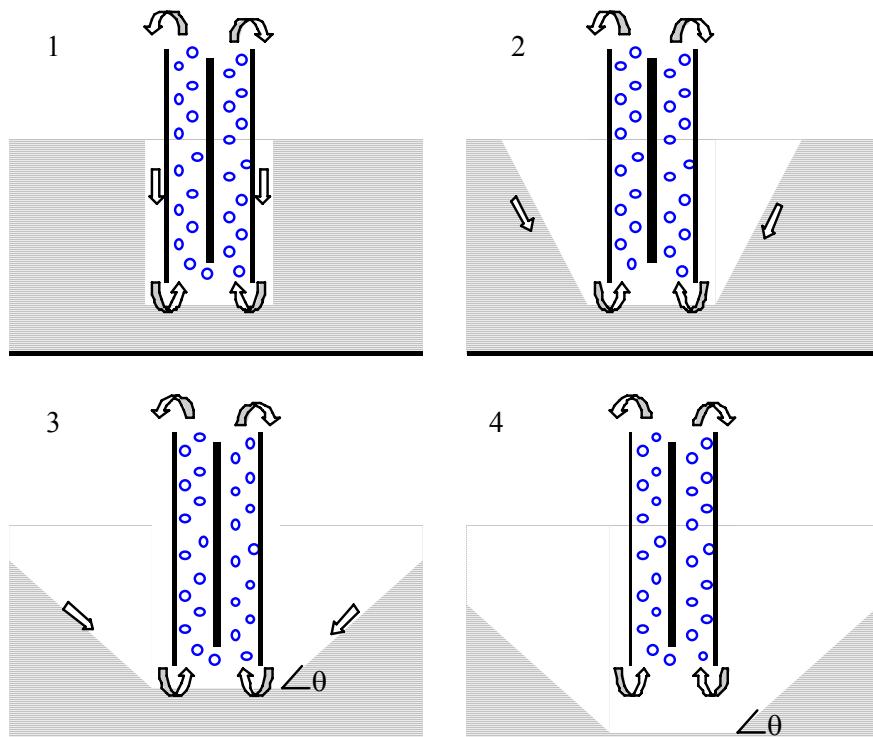


Figure 3.16. Stages of Waste Disturbance During ALC Operation

An expression for the maximum angle of repose, θ , for a weak granular sediment has been derived by Terrones^(a) and is given here:

$$\tan \theta = \frac{8\tau \Delta \rho g H}{(\Delta \rho g H)^2 - 16\tau^2} \quad (3.7.1)$$

where

τ = sediment shear strength.

H = distance from the base of the erosion to the top of the sediment layer. If the sediment is eroded all the way to the tank floor, then $H = H_S$.

$\Delta \rho$ = density difference between the sediment and the liquid above it.

During mixing, the supernatant density increases due to suspended solids. As this happens, the angle of repose increases, decreasing the rate of solids suspension. Thus, because suspension and resettling changes the preferred angle, which in turn changes the suspension rate, the system may oscillate and never reach a steady state. On the other hand, it is also possible that the system is highly damped and could quickly reach a steady state. Only further analysis and testing can determine which case applies.

The suction velocities and sediment properties determine the cleared base of the eroded conical volume around the ALC. To estimate the potential magnitude of the disturbed volume, assume the cleared base is equal to the diameter of the ALC. Hence, the eroded volume is a function of an inverted right circular cone with height H_S , minor base radius $D_0/2$, and major base diameter given by

$$D = D_0 + 2H_S \tan(\pi / 2 - \theta) \quad (3.7.2)$$

The volume is

$$V = \frac{\pi}{12} H_S (D^2 + D_0^2 + DD_0) \quad (3.7.3)$$

As noted above, mixing of the sediment into the supernatant will increase the angle of repose. In addition, the sediment height will reduce as the tanks walls are reached, further increasing the angle of repose. Thus, the model is conservative and is expected to overestimate the sediment volume disturbed especially for deep sediments that are relatively weak.

The pertinent waste properties for the tanks analyzed are shown in Table 3.8. Assuming $H = H_S$ and $\Delta \rho = \rho_s - \rho_L$, the angle of repose of the region disturbed by one ALC is shown in Table 3.9 for the DSTs analyzed. Also shown is the predicted gas release for 22 ALCs (17 in

(a) Meyer PA, CW Stewart, SD Rassat, RT Allemann, G Terrones, and DP Mendoza. May 1999. *Potential Gas Release by Bubble Slurry Flow Through a Hole in the Crust Layer in SY-101*. Letter Report TWS99.27 Rev. 1, Pacific Northwest National Laboratory, Richland, WA.

Table 3.8. Waste Properties for ALC Operation Analysis

Property	AZ-101 ^(a)	Group A	Group B
Supernatant Depth H_L (m)	8.49	5.91	4.73
Sediment Depth H_S (m)	0.48	4.5	1.57
Supernatant Density ρ_L (kg/m ³)	1240	1430	1150
Bulk Sediment Density ρ_S (kg/m ³)	1620	1572	1397
Sediment Shear Strength $\tau^{(b)}$ (Pa)	830	144	144
Sediment Gas Volume Fraction $\alpha^{(c)}$	0.07	0.04	.07
Gas pressure p (atm)	2.1	2.2	1.6

(a) Approximate AZ-101 waste conditions at time of ALC operation.
(b) BE Wells and SA Barker. 1999. *Summary of Yield Stress in Shear Data for Hanford Waste*. PNNL Letter Report TWS03.044. Pacific Northwest National Laboratory, Richland, WA.
(c) Barker et al. (1999)

Table 3.9. Disturbed Sediment Volumes

Property	AZ-101	Group A	Group B
Angle of repose (degrees)	90	10	17
Disturbed volume (m ³)	4	1847	646
Total gas released (m ³)	1	160	74
[H ₂] (ppm)	200	74800	13400

AZ-101) at standard pressure, computed from the disturbed volumes, gas volume fractions, and pressure ratios. Values for the disturbed volume vary widely because (per Eq. 3.7.1) a shallow sediment layer with relatively high shear strength such as that in AZ-101 has a high angle of repose, whereas deep sediments have a shallower angle, hence a larger volume. The small angles of repose in the hypothetical Group A and B tanks result in the potential for the entire sediment volume to be disturbed by ALC operation. This releases all of the stored gas. The Group A and B tanks are predicted to reach LFL and exceed 25% of the LFL respectively. Noting that the model is conservative for relatively deep and weak sediment, this potential for gas release indicates that ALC operation in Waste Group A or B tanks should be treated with the same care as mixer pump operation. The model hydrogen concentration results for AZ-101 are within an approximate factor of two of the recorded values for ALC operation in AZ-101 in May 2000 (430 ppm maximum hydrogen concentration) (Carlson et al. 2000).

Equation (3.6.3) applied to these example tanks demonstrates that the hypothetical Group A and Group B tanks could experience BDGRe during ALC operation. However, as the model conservatively predicts that the entire sediment will be mobilized, triggered BDGRe would not increase the headspace concentrations over those in Table 3.9 for these specific tanks. The gas release rates discussed below may be affected, however.

The actual cleared area at the base of the ALCs will be strongly affected by the magnitude of the induced flow velocities in the region. Models for predicting the flow rate in the ALC exist in the literature. One model presented in (Hetsroni 1982) gives the nondimensional flow rate as

$$\frac{Q_m^2}{2A^2 g(L - L_s)} = \frac{\alpha_m - (1 - \rho_L / \rho_S) + (1 + L_e / L_s)}{(4f / D_h)[L_e + L_s(1 - \alpha_m)^{-1.75}] + 2[1 / (1 - \alpha_m) - 0.75]} \quad (3.7.4)$$

where Q_m is the flow rate of the slurry mixture in the ALC, A is the ALC cross-sectional area, α_m is the average gas volume fraction in the rising column, ρ_m is the average density of the mixture, D_h is the hydraulic diameter, and f is the friction factor. The length, L_e , is the distance between the inlet and the air discharge nozzle, and the term L_s is the remaining length $L_s = L - L_e$. The mixture gas fraction depends on the flow rate and is given by

$$\alpha_m = \frac{Q_g}{Q_g + Q_m + sA} \quad (3.7.5)$$

where Q_g is the gas flow rate and s is the relative slip between the rising air bubbles and the fluid. The evaluation of Eq. (3.7.4) and (3.7.5) is iterative, with the mixture gas volume fraction being adjusted until the equation balances.

Eq. (3.7.4) was evaluated for in-tank conditions using a slip factor of 0.3 and a friction factor of 0.005. Flow rate and average rise velocity (Q_m/A) are plotted versus mixture density in Figure 3.17. The maximum flow rate occurs when no solids are entrained, so while it is not possible to quantify the disturbance size due to ALC operation exactly, we can make some useful observations regarding potential gas release rate. Because the flow rate is related to mixture

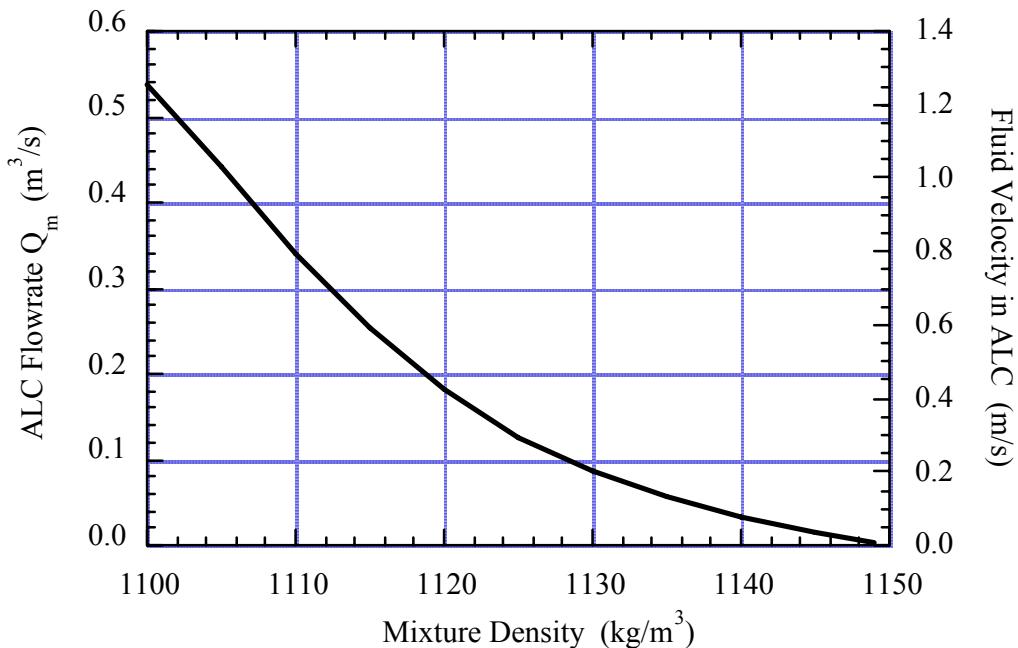


Figure 3.17. ALC Flow Rate and Velocity

density and mixture density is related to solids loading, we can relate ALC flow rate to gas release rate. The mixture is made up of convective liquid and sediment. Hence we can write

$$Q_m = Q_L + Q_S \quad (3.7.6)$$

where Q_L is the supernatant flow rate and Q_S is the flow rate of sediment. Mass conservation is expressed as

$$\rho_m Q_m = \rho_L Q_L + \rho_S Q_S \quad (3.7.7)$$

Combining Eq. (3.7.6) and (3.7.7) we obtain the entrainment rate of sediment:

$$Q_S = \frac{\rho_m - \rho_L}{\rho_S - \rho_L} Q_m \quad (3.7.8)$$

The corresponding total gas release rate at standard conditions, Q_{REL} , is given by

$$Q_{REL} = \gamma \frac{\rho_S}{\rho_A} Q_S \quad (3.7.9)$$

The predicted gas release rate for one ALC operating in a tank is shown in Figure 3.18. The gas release rises with increasing mixture density until a maximum of about 2 cfm is reached. The gas release then falls rapidly to zero as the ALC begins to stall.

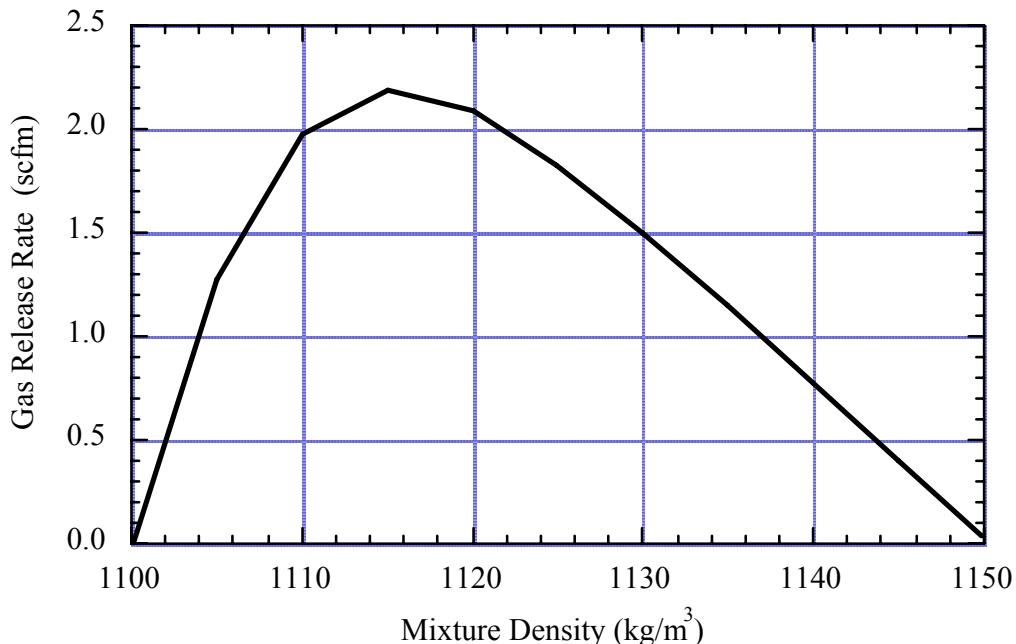


Figure 3.18. Gas Release Rate from a Single ALC

The results of this analysis suggest that ALCs may or may not be able to start due to the strength of the sediment and the modest lift pressures that can be achieved by 20 scfm of air flow. If started, fairly large waste disturbances are possible because the weak sediment can slough toward the inlet of the ALCs. However, analysis suggests that the flow velocities in the ALCs are relatively small, and even small solids loadings will tend to stall the ALCs. This result allows us to conclude that the gas release rate must be relatively low because the ALCs can only accommodate very low solids loadings.

3.7.3 Gas Monitoring Considerations During ALC Operation

Based on the analysis above, the gas release rate during ALC operation is likely to be limited. However, if one of the tanks with operable ALCs was placed in Waste Group A or B at some time in the future, the total gas release could bring the headspace to the LFL. Besides the gas released directly by the waste disturbance, it is also possible to induce BDGRe as the suspended sediment reduces the neutral buoyancy gas fraction. These events would add a rapid release rate component to that caused by the ALCs that could quickly drive the headspace hydrogen concentration above the action level. Ammonia would also be released at a potentially high rate depending on the airflow, whether the ALC started or not (see Section 3.7.1). This suggests that gas monitoring is necessary for ALC operation in Waste Group A and B tanks, that the monitoring should be continuous, and that some kind of degassing plan should be applied.

3.7.4 Potential Changes in Tank Classification After ALC Operation

Because, like mixing, ALC operation does not add or remove waste from the tank or change the waste configuration in the long term, no waste group change is possible.^(a) Hence reevaluation of waste group placement is not required. After initial degassing, periodic ALC operation would prevent gas accumulation and maintain the tank in Waste Group C if it were reclassified.

3.7.5 Potential Changes in Hydrogen Generation After ALC Operation

Hydrogen generation rates will not change as a result of ALC operation. Also, unlike mixer pump operation, the ALCs tend to be self-cooling and do not heat the waste appreciably. Reevaluation of hydrogen generation is not required.

3.8 Chemical Addition

Concentrated sodium hydroxide and possibly sodium nitrite solutions need to be added to DSTs occasionally to keep the waste chemistry within corrosion control limits. Additions are typically on the order of 50,000 gallons of solution with a specific gravity on the order of 1.5 (Fort 2001). No chemical additions are anticipated in SSTs except possibly during retrieval.

(a) The effect of “fluffing,” where the sediment layer expands and the previously mixed sediment gradually compacts to its former density, is discounted as a transient phenomenon compared to the minimum of three years before gas could accumulate to buoyancy (see Section 3.6.2).

Addition of chemicals other than for corrosion protection is not permitted. This section thus considers only chemical additions to DSTs.

The global disturbances from chemical addition may include possible solid dissolution as well as relatively small increase in the hydrostatic pressure, the same as a waste addition, along with an increase in the supernatant depth. However, the high density of the solution also increases the bulk density of the supernatant, thereby reducing the neutral buoyant gas fraction. In addition, depending on waste chemistry, solids may be precipitated.

3.8.1 Mechanisms for Gas Release During Chemical Addition

Mechanisms for appreciable gas release are solids dissolution and the decrease in neutral buoyancy gas fraction resulting from mixing the relatively heavy caustic solution in the supernatant. The reduction in neutral buoyancy gas fraction would likely induce a BDGRE in a Waste Group A tank^(a) and possibly in a borderline Group B tank for large-volume additions. BDGREs may also be triggered in any Waste Group by solids dissolution in the sediment.

The interactions between sodium, the anions in salt solids, hydroxide, and aluminum hydroxide are complex enough to make it impossible to guarantee conservatism in modeling the dissolution caused by caustic unless the chemistry of the waste is accounted for. A more detailed discussion of the modeling concerns for precipitation as well as dissolution can be found in the Appendix. The only additions that can be modeled conservatively in regards to gas release by dissolution without considering competing solubility equilibria are additions of water or of caustic that is much more dilute than the supernatant (as gauged by the relative specific gravities). Water addition, as analyzed in Section 3.5, bounds the dissolution that may occur from any caustic addition.

The analysis for the decrease in neutral buoyancy from a chemical addition is exactly the same as that described for waste addition in Section 3.2. With no dissolution in Group A tanks, induced BDGREs can be avoided by ensuring that the density of the solution added is less than that of the existing supernatant. With no dissolution in Groups B and C tanks, induced BDGREs can be prevented by adjusting the addition to prevent spontaneous BDGREs in the future based on an evaluation of post-operation tank classification via Eq. (2.13). Sediment dissolution may trigger a BDGRE in any Waste Group (Section 3.5). The likelihood of enhancing ammonia evaporation is less for chemical addition because added waste usually contains ammonia.

The temperature of chemical additions may be higher than that of the waste, which causes retained gas in the sediment to expand so that it could theoretically become buoyant. However, as discussed in Section 3.2, a very large temperature change would be needed to cause buoyancy, and temperature change is not considered a serious gas release mechanism.

(a) As described in Section 2.4.1, Waste Group A and B tanks store sufficient gas to make the headspace flammable if all of it were suddenly released; Waste Group C tanks do not. Waste Group A tanks exhibit spontaneous BDGREs, while tanks in Waste Group B do not.

3.8.2 Expected Gas Releases During Chemical Addition

The gas released due to potential dissolution from a chemical addition is bounded by the analyses of Section 3.5 for water addition. No specific analysis has been performed on gas release rates and volumes of BDGReS induced in Group A tanks during chemical addition. As with waste addition (Section 3.2), the initial gob size of BDGReS induced during chemical addition should be similar to that of spontaneous releases. The peak hydrogen concentration would be increased above historical norms by the smaller headspace volume and higher gas release fractions. However, the increase would not be sufficient to make a historic Waste Group A tank exceed the LFL, as shown in Section 3.5.4.

There have been no analyses or experience on the size of BDGReS induced in Group B or C tanks that have not experienced them in the past. Therefore, it is not possible to apply Eq. (3.5.5) because these tanks have no gas release history to define a release volume. However, the behavior of the BDGRE tanks shows that release size is roughly proportional to the buoyancy ratio (Stewart et al. 2005). Thus the likelihood of large dissolution-induced BDGReS in Group B DSTs is very low. No large gas release is expected from any other mechanism if the chemical addition is adjusted to prohibit BDGReS.

3.8.3 Gas Monitoring Considerations During Chemical Addition

With dissolution, chemical additions of less than the upper-bound water volume (Eq. 3.5.2) in tanks with $BR < 0.75$ or $ER < 3$ (Section 3.5) should not require monitoring. In Waste Group-B DSTs that have $BR \geq 0.75$ and $ER \geq 3$, the likelihood of a dissolution-induced BDGRE threatening flammability is low. For chemical additions that exceed the upper-bound water volume, active ventilation and flammable gas monitoring will be required; however, no monitoring frequency or duration can be provided. Chemical additions to Waste Group-A DSTs should be limited to prevent flammable conditions if a large BDGRE were to occur simultaneously with the dissolution-induced gas release (as analyzed in Section 3.5 for water addition).

Without dissolution, no monitoring is required in Group B tanks as long as the addition is adjusted to prevent creating a new Waste Group A tank, which also prevents gas releases during the addition. Spontaneous BDGReS can occur at any time in Waste Group A tanks, and chemical addition without dissolution would exacerbate the resulting hydrogen concentration somewhat, though not near the LFL, as presented in Section 3.5.4.^(a) However, there is no effective control for spontaneous BDGReS, even if one were detected during the operation.

3.8.4 Potential Changes in Waste Classification After Chemical Addition

Caustic additions to DSTs for corrosion control are likely to be relatively large. Table 3.10 shows the change in buoyancy ratio resulting from a chemical addition of 100,000 gallons with a

(a) No potential precipitation is accounted for in the analysis for BDGRE size in Section 3.5.4. Any new sediment formation, however, would only potentially affect BDGRE size after steady state conditions had been reached.

Table 3.10. Change in Buoyancy Ratio for Addition of 100 kgal at SpG = 1.5

Group	Buoyancy Ratio	
	Before	After
B	0.93	0.95
B	0.56	0.60
C	0.48	1.3
C	0.97	1.2

specific gravity of 1.5 in four hypothetical DSTs representing Waste Groups B and C. The buoyancy ratios before and after the operation are calculated using Eq. (2.13). The tanks chosen represent the maximum buoyancy ratio values for their respective waste classification prior to caustic addition. This caustic addition causes the buoyancy ratio to exceed unity only in the hypothetical C tanks, but these tanks would remain in Waste Group C because of their large headspace. The buoyancy ratio increases slightly in the B tanks. Note that no change is considered in sediment depth due to solid dissolution and/or precipitation.

This exercise shows that it is potentially possible to move a tank into or to the borderline of Waste Group A with caustic addition even without a potential increase in sediment depth due to solids precipitation. As evident from Eq. (2.13), growth of the sediment layer due to precipitation, with all other parameters constant, will further increase the buoyancy ratio. Therefore, an evaluation of the end-state tank classification is required for caustic additions to Group B and C tanks.

The effect on BDGREs beyond that presented in Section 3.5.4 due to a potential increased sediment depth from precipitation in Group A tanks is uncertain. Further, as discussed in Section 3.8.1, the interactions between the waste's constituents are complex enough to make it impossible to accurately predict precipitation caused by caustic addition unless the chemistry of the waste is accounted for. Thus, evaluation of the possible precipitation and the subsequent effect of an increased sediment depth on BDGRE size is required for caustic additions to Group A tanks.

3.8.5 Potential Changes in Hydrogen Generation After Chemical Addition

A chemical addition to a tank adds to the total liquid inventory and adds to the amount of water available for radiolysis. It is not known whether increasing the hydroxide concentration would increase gas generation by itself or by dissolving solids, which would increase the aluminum and salt concentrations in the liquid. To accommodate the potential for changes, the hydrogen generation rate must be reevaluated for chemical additions.

3.9 Natural Evaporation

Evaporation of water from the waste historically has not been considered a flammable gas issue but is treated here for completeness. It occurs naturally in all tanks but is important only in tanks with a relatively dilute supernatant and a high heat loading in the sediment. Significant evaporation has not been observed in concentrated saltcake tanks regardless of heat loading with

either passive or active headspace ventilation. The presence of a floating crust layer further inhibits evaporation in the historic BDGRE tanks (see Hedengren et al. 2000). In fact, psychometric data showed that the DST SY-101 actually absorbed water for long periods before it was remediated by transfer and dilution in early 2000. Evaporation is a potential issue only for DSTs whose waste configuration would make them susceptible to BDGREs. Reduction of the liquid content by evaporation is considered beneficial in SSTs with waste in the wet sediment or pumped configuration.

Evaporation causes two kinds of global disturbance. Slowly removing water from the supernatant and lowering the waste level decreases the hydrostatic pressure on the sediment and increases the density of the supernatant.

3.9.1 Mechanisms for Gas Releases During Evaporation

Evaporation of water from the waste does not itself release retained gas. However, it increases the concentration of dissolved solids in the supernatant, thereby raising its density. Carried to an extreme, the increasing supernatant density may reduce the neutral buoyancy gas fraction sufficiently to allow BDGREs. However, the reduction in supernatant depth would also tend to reduce the size of gas releases or prevent them altogether. At the same time, the increased salt concentration in the liquid reduces evaporation, so the process tends to be self-limiting.

If a BDGRE occurs due to a reduction in the neutral buoyancy gas fraction, it will suspend sediment temporally in the supernatant, further reducing the neutral buoyancy gas fraction to potentially induce additional BDGREs. Decreasing hydrostatic pressure, which expands retained gas, would also act to promote gas release. However, the process of evaporation is so slow that the expanded bubbles would be released long before this could have an effect.

Though a reduction in neutral buoyancy void fraction resulting from evaporation would most likely induce a BDGRE in a Group A tank, the tanks that historically exhibit BDGREs have a very high salt concentration along with a floating crust layer, and their waste level history shows no evidence of significant evaporation (Hedengren et al. 2000).^(a) Evaporation might eventually induce a BDGRE in a borderline Group B tank; however, the tanks most likely to be borderline tend to have the most concentrated supernatant and therefore the least evaporation. The tanks with a high evaporation rate have a very dilute supernatant and are not likely to become an issue.

In either case, the analysis performed in Section 3.2 for waste addition also applies to evaporation, which is simply the reverse of addition. The possibility of BDGREs induced by evaporation can be predicted and prevented if the effects of evaporation are evaluated and halted before the tank classification moves into Group A. The evaluation method is developed following the derivation given in Section 3.2.1. Figure 3.19 shows a sketch of the waste configuration before and after evaporation assuming no precipitation of solids. We assume a

(a) As described in Section 2.4.1, Waste Group A and B tanks store sufficient gas to make the headspace flammable if all of it were suddenly released; Waste Group C tanks do not. Waste Group A tanks exhibit spontaneous BDGREs, while tanks in Waste Group B do not.

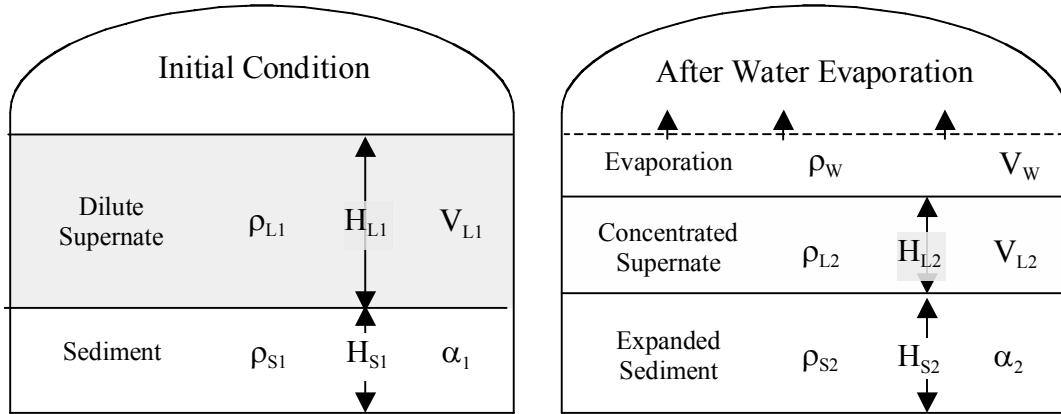


Figure 3.19. Tank Waste Configuration Before and After Evaporation

volume V_W of water with density ρ_W leaves the supernatant, which has an initial density ρ_{L1} and thickness H_{L1} . After evaporation, the concentrated supernatant will have decreased thickness H_{L2} and increased density ρ_{L2} . The thickness of the gas-retaining sediment layer will increase slightly from H_{S1} to H_{S2} due to the expansion of the retained gas from decreased hydrostatic pressure. The gas volume fraction will increase from α_1 to α_2 in the process.

The mixture density is expressed in terms of the evaporated water volume and initial quantities:

$$\rho_{L2} = \frac{\rho_{L1}V_{L1} - \rho_WV_W}{V_{L1} - V_W}, \quad V_W < V_{L1} \quad (3.9.1)$$

Combining Eq. (2.12) for the neutral buoyancy gas fraction with Eq. (3.1.2) for the final gas volume fraction provides an expression for the buoyancy ratio after evaporation:

$$BR_2 = \frac{\alpha_2}{\alpha_{NB2}} = \frac{\alpha_1\rho_S/(\rho_S - \rho_{L2})}{\alpha_1 + (1 - \alpha_1)\bar{P}_2/P_1} \quad (3.9.2)$$

The pressure ratio is given by Eq. (3.2.4) using the final density from Eq. (3.9.1) and the final supernatant depth computed by Eq. (3.2.2) with V_T replaced by V_W . Substituting the product of the initial buoyancy ratio, Eq. (2.13), and the initial neutral buoyancy gas fraction, Eq. (2.12), for the initial gas fraction yields

$$BR_2 = \frac{\frac{\rho_S}{\rho_S - \rho_{L2}}}{1 + \left[\frac{\rho_S}{BR_1(\rho_S - \rho_{L1})} - 1 \right] \frac{\bar{P}_2}{P_1}} \quad (3.9.3)$$

If the final buoyancy ratio is greater than unity, a BDGRE is likely.

For evaporation, using Eq. (3.9.1) and (3.9.3) yields a larger final buoyancy ratio than that calculated directly with Eq. (2.13), which is used to evaluate the buoyancy ratio for waste group selection. However, the derivation of Eq. (3.9.3) assumes that no gas escapes from the sediment during evaporation, so all gas present initially expands as the pressure decreases. This is not a reasonable assumption for the long period over which a large volume of waste would evaporate. Therefore, the reevaluation of the tank classification and adding liquid to reverse the effects before evaporation moves the tank into Group A will also be sufficient to prevent inducing BDGReEs during evaporation.

3.9.2 Expected Gas Releases During Evaporation

Evaporation does not itself release retained gas. The only gas release hazard associated with evaporation is the potential for eventually creating a Group A tank. However, the process is very slow and can be reversed easily by adding back makeup water.

There has been no experience or analyses on the potential size of BDGReEs induced in Group B or C tanks that have not experienced them in the past. In the absence of historic BDGReEs, the waste should be more uniform, so the first induced BDGRE might be larger than indicated by the behavior of the historic BDGRE tanks.

3.9.3 Gas Monitoring Considerations During Evaporation

Evaporation is a chronic, long-term effect whose main adverse consequence is increasing the density of supernatant and a potentially changing tank classification. Because no gas release is associated with evaporation itself, no gas monitoring is required. The waste level is monitored in all tanks so the cumulative effects of evaporation can be tracked in tanks where it is of concern.

3.9.4 Potential Changes in Tank Classification After Evaporation

The amount of evaporation needed to cause a BDGRE was evaluated for four hypothetical DSTs in Waste Groups B and C. Figure 3.20 shows the buoyancy ratio calculated with Eq. (3.9.2) versus the volume of water evaporated in four tanks: Tanks 1 and 2 in Group C and Tanks 3 and 4 in Group B. The parameters used in the calculation are listed in Table 3.11. The waste level history indicates that the evaporation rate in all tanks but Tank 2 is negligible.

Because the supernatant in Tanks 1 and 4 is already somewhat concentrated (see Table 3.11), evaporation rapidly increases the supernatant density, reducing the neutral buoyancy gas fraction and increasing the buoyancy ratio. Eq. (3.9.2) predicts that Tank 4 would reach a buoyancy ratio of 1.0 after evaporation of 30,000 gallons or 10 inches of water. Tank 1 could begin exhibiting BDGReEs (though it would remain in Group C) after about 190,000 gallons or 72 inches of water have evaporated. Tanks 2 and 3 have a more dilute supernatant, so evaporation changes the buoyancy ratio more slowly. About 20,000 gallons or 7 inches of evaporation would need to evaporate to create a BDGRE in Tank 2; Tank 3 would require about 250,000 gallons, or 90 inches of evaporation.

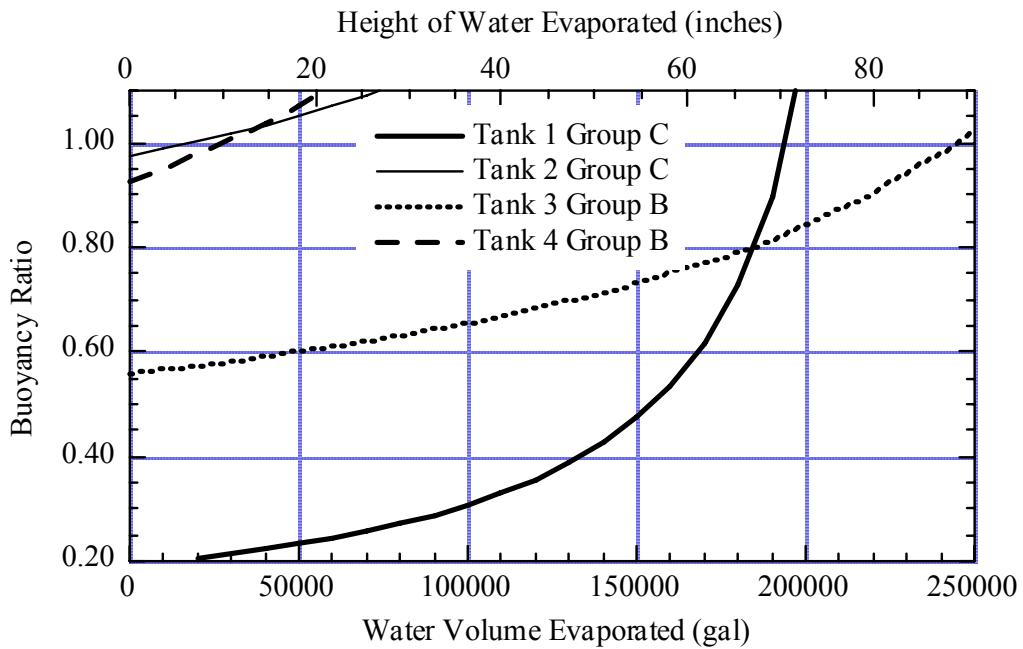


Figure 3.20. Effects of Evaporation on Buoyancy Ratio

Table 3.11. Tank Data Used for Evaporation Analysis

Property	Tank 1 Group C	Tank 2 Group C	Tank 3 Group B	Tank 4 Group B
H_{L1} (m)	9.1	4.7	7.3	7.3
H_S (m)	0.8	1.6	2.9	2.3
ρ_{L1} (kg/m^3)	1430	1150	1240	1370
ρ_S (kg/m^3)	1560	1397	1498	1560
α_{NB1}	0.08	0.18	0.17	0.12
BR_1	0.19	0.97	0.56	0.93

This analysis shows that evaporation can, in theory, change a tank's classification, though a large amount of evaporation would be required. Tank 2 is a possible exception because its initial buoyancy ratio is very close to unity. However, this is not a concern because the tank is in Waste Group C. With daily waste level monitoring, any such trends would be obvious long before serious concentration occurred, and makeup water could be added to prevent it.

The self-limiting nature of evaporation can be evaluated by correlating the water partial pressure in a concentrated salt solution with the solution density. The water partial pressures measured by Norton and Pederson (1994) in simulated DST wastes were correlated with approximate densities derived from simulant compositions. The resulting correlation is

$$P_{H_2O} = (\rho_L - 1)^{-0.00681} \exp[13.51 - 0.0038(\rho_L - 1)T - 1.669(\rho_L - 1)^2 - 5064/T] \quad (3.9.4)$$

where

- P_{H2O} = water partial pressure above the solution (atm)
- ρ_L = density of the salt solution (g/mL)
- T = temperature of the solution (K).

The water partial pressure in the waste given by Eq. (3.9.4) is plotted versus liquid density in Figure 3.21 for solution temperatures of 300, 310 and 320K (approximately 80°, 100° and 120°F). There is approximately a factor of 3 decrease in water partial pressure going from a density of 1.0 to 1.5 g/mL. The maximum evaporation rate will occur in the area where ambient air is first exposed to the waste surface. A representative partial pressure for water vapor in the Hanford atmosphere^(a) is on the order of 0.03 atm, which is shown in the figure. The maximum rate of evaporation is roughly proportional to the difference between the equilibrium partial pressure of the waste and the partial pressure of water in the incoming air. This indicates that the evaporation rate will become small for solution densities above about 1.1 g/mL at 300K and roughly 1.4 g/mL at 310K. Evaporation also decreases strongly at higher densities at 320K even though it does not go to zero.

Because evaporation effectively ceases before it can raise the liquid density to the point where BDGReS would occur, we conclude that evaporation does not pose a significant risk.

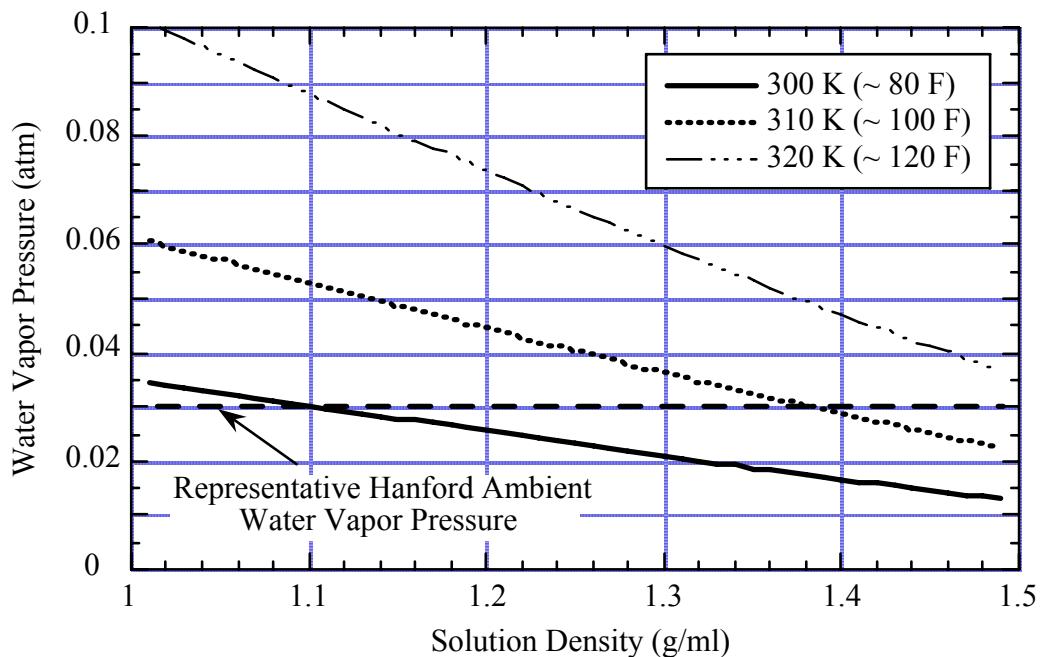


Figure 3.21. Water Partial Pressure Versus Solution Density

(a) Mahoney LA. April 2000. *Modeling Evaporation from the 241-AX-152 Catch Tank*. Letter report TWS00.44, Pacific Northwest National Laboratory, Richland, WA.

3.9.5 Potential Changes in Hydrogen Generation After Evaporation

Because evaporation removes water from the tank that would have been subject to radiolysis, the gas generation rate should decrease. Therefore, the steady-state flammability is bounded by the initial calculation and the hydrogen generation rate need not be reevaluated.

4.0 Summary and Conclusions

The analyses presented in Section 3 based on the theory summarized in Section 2 have quantified the effects of nine globally waste-disturbing activities in terms of gas release potential as well as changes in waste behavior. Analyses for modified sluicing are presented in Huckaby and Wells (2004). The practical effects of each operation differ depending on the initial waste group assignment of a tank. The three waste groups are defined in Section 2.4.1. Waste Group A and B tanks store sufficient gas to make the headspace flammable if all of it were suddenly released; Waste Group C tanks do not. Waste Group A tanks exhibit spontaneous BDGReS, while tanks in Waste Group B do not.

This section summarizes the conclusions of the analyses by waste group. Section 4.1 summarizes the gas releases that could be induced by the nine globally waste-disturbing activities, and Section 4.2 discusses potential long-term changes in waste group classification. Section 4.3 lists cases where the hydrogen generation rate and steady-state flammability need to be addressed.

4.1 Potential for Significant Gas Releases

A waste-disturbing activity can cause or influence gas releases in three ways: 1) spontaneous BDGReS in Waste Group A tanks may be amplified, 2) BDGReS can be induced by the effects of the operation, and 3) the waste disturbance itself can produce substantial gas releases that can accumulate to flammability in the headspace if the ventilation rate is low or the release rate is rapid compared with the ventilation rate. Table 4.1 lists the potential gas releases for each globally waste-disturbing activity by waste group. Note that, though gas releases occur in Waste Group C tanks, they are inconsequential by definition because these tanks have insufficient retained gas to make their headspace flammable even if all the gas is released instantaneously.

Inducing or exacerbating BDGReS is the greatest concern in any globally waste-disturbing activity in DSTs. Any addition to a tank already exhibiting spontaneous BDGReS will increase the resulting headspace hydrogen concentration by decreasing the headspace and increasing the hydrostatic pressure on the gas. However, with no dissolution, this was shown in Section 3.5.4 not to cause previously nominal BDGReS in the historic Group A tanks to exceed the LFL for addition of water or caustic to the maximum waste level. Adding volume could move a Group C tank exhibiting BDGReS directly to Group A. BDGReS can also be induced by removing low-gas overburden from the sediment by dissolution and as a result of lowering the neutral buoyancy gas fraction, α_{NB} , by suspending sediment in the supernatant (by mixer pump or airlift circulator operation), or by otherwise increasing the supernatant density (by waste or chemical additions). With dissolution, and with or without BDGReS, water addition volumes must be limited so as to not exceed the LFL, Section 3.5.2.

In SSTs there is no mechanism for large spontaneous or induced gas releases. However, these tanks are typically passively ventilated, and a relatively slow gas release during saltcake dissolution retrieval, water addition, or saltwell pumping can potentially accumulate to eventually raise the headspace hydrogen concentration to the LFL in Group A or B tanks.

Table 4.1. Potential Gas Releases During Globally Waste-Disturbing Activities

Operation	Waste Group		
	A	B	C ^(a)
Waste Removal (DST)	BDGRE (depressurization) Slow release (uncovery, depressurization)	Late BDGRE ^(b) (depressurization) Slow release (uncovery, depressurization)	BDGRE (depressurization) Slow release (uncovery, depressurization)
Waste Addition (DST)	Amplified BDGRE (higher pressure, smaller headspace, potential added sediment)	BDGRE (lower α_{NB}) ^(c)	BDGRE (lower α_{NB}) ^(c)
Waste Addition (SST)	N/A	Slow release (dissolution)	Slow release (dissolution)
Saltwell pumping (SST)	N/A	Slow release (uncovery, depressurization)	Slow release (uncovery, depressurization)
Saltcake dissolution (SST)	N/A	Slow release (dissolution)	Slow release (dissolution)
Water addition (SST)	N/A	Slow release (dissolution)	Slow release (dissolution)
Water addition (DST)	Amplified BDGRE (higher pressure, smaller headspace) ^(d) Slow release, BDGRE (dissolution)	Slow release, BDGRE (dissolution)	BDGRE (higher pressure, smaller headspace) Slow release, BDGRE (dissolution)
Mixer pump operation (DST)	BDGRE (lower α_{NB}), large disturbance	BDGRE (lower α_{NB}), large disturbance	BDGRE (lower α_{NB}), large disturbance
Airlift circulator operation (DST)	BDGRE (lower α_{NB}), large disturbance	BDGRE (lower α_{NB}), large disturbance	BDGRE (lower α_{NB}), large disturbance
Chemical addition (DST)	Amplified BDGRE (higher pressure, smaller headspace) ^(d) Slow release (dissolution)	BDGRE (lower α_{NB}) ^(c) Slow release (dissolution)	BDGRE (lower α_{NB}) ^(c) Slow release (dissolution)
Natural evaporation (DST)	None	None	None

(a) Gas releases are listed for Group C tanks but have no consequence because these tanks retain insufficient gas to make their headspace flammable.
(b) Induced BDGRES, if any, are expected late in the process where headspace is maximum and flammability potential is minimum.
(c) BDGRES can be prevented by adjusting the waste addition to keep the buoyancy ratio < 1.
(d) Water and caustic addition with no dissolution do not cause current BDGRES to exceed the LFL.

Analytical methods were developed in Section 3 to determine the conditions under which gas releases, specifically BDGRES, would occur. In most cases these analyses indicate that flammable conditions will not occur as a result. Where there is some potential for BDGRES, operations can be adjusted to avoid it in most cases.

4.2 Potential Changes in Waste Group Classification

The waste group classification depends basically on whether a tank retains enough gas to make its headspace flammable if all of it were released suddenly and whether it experiences spontaneous BDGReS. The potential changes in waste group classification for the nine activities are listed by initial waste group in Table 4.2. The retained gas condition is determined mainly by the headspace so any addition to a tank tends to move it to a higher group. A Group C tank experiencing BDGReS could move all the way into Group A by waste addition.

Table 4.2. Potential Changes in Waste Group from Globally Waste-Disturbing Activities

Operation	Waste Group		
	A	B	C ^(a)
Waste Removal (DST)	B or C	A ^(b) or C	None
Waste Addition (DST)	None ^(c)	A (lower α_{NB})	A or B (lower α_{NB} , smaller headspace)
Saltwell pumping (SST)	N/A	C	None
Saltcake dissolution (SST)	N/A	C	None
Water addition (SST)	N/A	Footnote (d)	B (smaller headspace); A not applicable ^(d)
Water addition (DST)	B or C with dissolution; See also footnote (c)	A possible with or without dissolution; C with dissolution	A or B (smaller headspace)
Mixer pump operation	None	None	None
Airlift circulator operation	None	None	None
Chemical addition	None ^(c)	A (lower α_{NB})	A or B (lower α_{NB} , smaller headspace)
Natural evaporation	None ^(e)	C ^(e)	None ^(e)

(a) Waste Group C tanks cannot move into a lower group, but removal of waste will decrease their retained gas volume and potential for releases.

(b) A Waste Group A tank can be created by waste removal only if: 1) the supernatant removal is sufficient that the resulting pressure reduction in the sediment causes the buoyancy ratio to exceed unity, 2) the supernatant removal is not sufficient to result in an energy ratio less than three, and 3) the supernatant removal is not sufficient to result in a tank classification of Waste Group C.

(c) Waste Group A tanks cannot move to a higher group, but addition of waste, water, or chemicals may amplify their spontaneous BDGReS.

(d) Water additions to SSTs are restricted to preclude BDGReS by prohibiting additions that result in an energy ratio greater than three (Kripps 2005).

(e) Significant evaporation does not occur in tanks with a high salt concentration, which historically included all Group A tanks. This also prevents initiation of BDGRE behavior by evaporation.

Though many factors influence BDGRE behavior, the dominant parameter is the neutral buoyancy gas fraction, which depends on the ratio of supernatant density to bulk sediment density. Therefore, any operation that makes the supernatant more dense can potentially decrease the neutral buoyancy gas fraction, initiate BDGRES, and move a tank from Group B to Group A. However, initiating BDGRES in a Group C tank will not move it to Group A unless the headspace is reduced sufficiently simultaneously. Waste, water, or chemical addition proves to be the only activity capable of moving a tank to a higher waste group except for waste removal in specific instances. Conversely, all of the waste removal activities can move a tank to Group C.

4.3 Potential Changes in Hydrogen Generation

The hydrogen generation rate depends on the amount of waste, specifically the liquid portion, the concentration of dissolved salts, TOC, and aluminum in the liquid. Waste addition may cause an increase in the hydrogen generation rate, and any major removal of waste, especially liquid, will decrease it.

Hu (2004b) calculated the effect of adding 10,000 gallons of water to DSTs.^(a) The calculation assumed that the added water mixes only with the supernatant liquid with no solids dissolution. The resulting reduction of the TOC concentrations lowers the unit generation rate. This effect is “balanced” in the total generation rate however by the increase in waste volume. The total generation rate was shown to typically increase due to the increase of water radiolysis.

We expect that, because natural evaporation will increase the ionic concentration while reducing the waste volume and water radiolysis, the total hydrogen generation should be reduced. Addition of caustic may or may not increase the hydrogen generation and should be evaluated.^(a)

The changes in hydrogen generation rate resulting from the nine activities are listed in Table 4.3. Because the effect of an operation on hydrogen generation rate does not depend on the waste group, only one column is shown in the table.

Energy dissipation from one or more mixer pumps being run for long periods is an issue. Mixer pump operation can elevate the waste temperature considerably. Because hydrogen generation follows Arrhenius behavior (see Section 2.1), it is sensitive to temperature. A relatively short period of continuous mixer pump operation could easily increase the hydrogen generation rate by an order of magnitude. If the mixer pump must be run such that the waste temperature remains elevated for an extended period, the steady-state flammability hazard of the tank should be evaluated at the new temperature.

(a) Personal communication with Albert Hu (CH2M HILL).

Table 4.3. Changes in Gas Generation from Globally Waste-Disturbing Activities

Operation	Effect on Hydrogen Generation
Waste removal (DST)	Decrease
Waste addition	Increase
Saltwell pumping (SST)	Decrease
Saltcake dissolution (SST)	Decrease
Water addition	Increase
Mixer pump operation	Pump energy deposition issue
Airlift circulator operation	None
Chemical addition	Potential increase
Natural evaporation	Decrease

4.4 Implications for Other Operations

This report has treated nine specific globally waste-disturbing activities that are currently performed or proposed in SSTs and DSTs. In addition to these activities, there are many local waste-disturbing activities. These activities are discussed in Sections 4.4.1 in terms of the analysis and conclusions derived for the globally waste-disturbing activities. Some globally waste-disturbing activities that have not been analyzed are noted in Section 4.4.2. A brief summary is given in Section 4.4.3.

4.4.1 Local Waste Disturbances

The local waste-disturbing activities can be grouped into four broad categories according to the general type of disturbance involved. Each category is briefly examined below. No gas release associated with these activities is expected to create headspace hydrogen concentrations exceeding the action level (see the historical summary in Section 1.2.1), nor can any of them change a tank's waste classification. However, it is important to note that spontaneous BDGREGs can occur at any time in Waste Group A tanks that have historically raised the hydrogen concentration near or above the action level (though less than 50% of the LFL).

Water lancing: operation of water lances or jets lowered vertically into the waste, usually to create a passage or cavity for installing other waste-penetrating equipment. Water lances range in design from a single vertically oriented nozzle that opens a 4-inch-diameter hole to the 42-inch cruciform lance that uses high-pressure jets to excavate for the SY-101 mixer pump in 1993. Water additions during lancing are not expected to approach the thresholds for hazardous gas release calculated in Section 3.5.

Equipment installation: Examples include installation of mixer pumps and transfer pumps, saltwell screens and liquid observation wells, and thermocouple trees and multi-function instrument trees. Diameters range from less than 4 inches to about 40 inches. Installation involves pushing equipment vertically into the waste. This causes a local waste disturbance of approximately the same volume as that of the device below the

waste surface. Water lancing is often a prerequisite for installation, or a water lance is designed into the base of the equipment to ensure that it penetrates the waste easily.

Equipment removal: Examples are removal of mixer pumps, sludge weights, air lances, liquid observation wells, thermocouple trees, and multifunction instrument trees. Removing equipment that has been installed in the waste for a long time requires rather extensive water flushes for decontamination, which adds to the direct waste disturbance of the removal itself. Flushes are not expected to approach the threshold for hazardous gas release calculated in Section 3.5.

Equipment operation: This involves operation of the ball rheometer, VFI, core sampling, densitometer readings, and liquid grab sampling. Running the validation probe inside a multifunction instrument tree or taking neutron or gamma logs in a liquid observation well are not considered because they occur inside a sealed volume and do not disturb the waste. Operation of waste-penetrating equipment is actually a combination of the installation and removal disturbances. However, because such equipment is designed specifically for operating in the waste, the disturbance attending its operation is typically less than that of devices intended for permanent installation. Measurement and sampling devices are typically of small diameter and flushes are small because they have been designed for easy decontamination.

4.4.2 Other Global Waste Disturbances

Gas releases expected during modified sluicing are described in Huckaby and Wells (2004). Vacuum retrieval is described below. This activity would require further definition and study to assess the potential hazard adequately. The main concern about gas releases during global waste disturbances is the amount of waste disturbed and its release rate. Given the retained gas volume fraction and assuming that the disturbance is sufficient to release most or all of gas, the volume of waste disturbed in a given time is a direct measure of the volume of gas released. If this is large compared with the headspace ventilation rate, high hydrogen concentrations can result. Therefore, planning an activity must be aimed at controlling the waste disturbance to limit gas release to acceptable values.

The system for the vacuum retrieval scheme consists of the Articulated Mast System (AMS) and the Mobile Retrieval System (MRS) or “crawler.” The AMS utilizes injected air and a “lubricating fluid,” either water or supernatant liquid, to vacuum waste out of the tank. High pressure water or supernatant liquid may be used to mobilize the waste via the scarifier. The MRS is a hydraulically driven device used as a bull-dozer to move waste into the range of the AMS in the larger waste tanks (i.e., all tanks except the 200 series). The MRS includes a “cannon” which can mobilize waste using water or recirculated supernatant.

In terms of the globally waste-disturbing activities discussed in this report, the vacuum retrieval system represents a combination of mixing and dissolution and has the potential for relatively large and rapid waste disturbances. If performed in a Waste Group B tank, the potential for reaching the LFL would possibly exist; planning of the operation must take this into account.

4.4.3 Summary

Considering the entire spectrum of waste-disturbing activities described above and those analyzed in Section 3 of this report, it is difficult to conceive of a larger or more rapid waste disturbance than operation of one or more large mixer pumps. Because these pumps were designed to have sufficient power to mobilize the entire waste volume and mix it with sufficient intensity to keep the sediment suspended uniformly, they can quickly release all of the gas in a tank. Waste retrieval operations will also release all of a tank's retained gas but will do so more gradually or locally, so it is easier for ventilation to keep up with the release. These combined factors indicate that mixer pump operation may be the bounding global waste disturbance. Past experience and the recent analyses could serve as a model for planning similar, but less aggressive, activities.

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Appendix

Technical Background for Dissolution and Precipitation

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A.1 Governing Principle

The governing principle of dissolution and precipitation is that one specific type of product of the thermodynamic activities of the dissolved species is constant, for a given temperature, whenever the solution is saturated with the solid produced by those dissolved species. This type of product is called the “solubility product.” As an example, for sodium phosphate dodecahydrate ($\text{Na}_3(\text{PO}_4)_2 \cdot 12\text{H}_2\text{O}$), the solubility product (also called the K_{sp}) is calculated as

$$K_{\text{sp}, \text{NaPO}_4} = a_{\text{Na}}^3 a_{\text{PO}_4}^1 a_{\text{H}_2\text{O}}^{12} \quad (\text{A.1})$$

where

- a_{Na} = the activity of dissolved sodium (a function of the concentration)
- a_{PO_4} = the activity of dissolved phosphate
- $a_{\text{H}_2\text{O}}$ = the activity of the solvent water

It can be seen that for each species in the solid the exponent in the K_{sp} expression is the same as the number of moles of the species in the solid. This is the general rule for all solubility products. Thus, when there is one sodium atom in the compound, the Na activity is raised to the first power; when there are two, the Na activity is squared, and so on.

So long as a solid’s activity product in solution is less than that solid’s K_{sp} , which is fixed for a given temperature, the solution is unsaturated in that solid and the solid does not precipitate. Once the activity product exceeds the K_{sp} , enough solid precipitates to make the product equal the K_{sp} . As long as the solid is present the activity product must equal the K_{sp} . The situation becomes more complicated when many different solids, some of which contain the same species, are present. Nevertheless, it remains the rule that the species in solution are in equilibrium with all the solids that are present.

Before the activity product can be calculated the activities of the species in solution must be determined. The thermodynamic activity of a species in solution is the product of its activity coefficient, γ_i , and its concentration. The activity coefficient is equal to unity at infinitely dilute solution for all species, meaning that the activity equals the molality. However, in solutions as concentrated as those handled at Hanford, the activity coefficient is greatly different from unity and can vary considerably depending on the species, the total ionic strength of the solution, and the other species present in the solution.

The calculation of activity coefficients is complicated and the data for it may not be available for the species of concern in Hanford waste. A simplifying assumption that uses the concentrations in solution instead of activities (as if $\gamma_i = 1$ for all species) can be made. The

assumption introduces an approximation that will be very significant, the size of which varies with the ionic strength of the solution (in effect, with the total Na concentration), but allows a concentration product to be substituted for the activity product. Any models based on this assumption are likely to be grossly inaccurate; it is used in the following development only to simplify the equations for conceptual purposes.

A.2 Conceptual Application – Salt Dissolution

The addition of caustic to the supernatant layer of waste has three chemical effects that relate to gas release by dissolution. First, caustic changes the sodium concentration of the liquid, making it more or less concentrated depending on whether the sodium concentration in the caustic is greater or less than that in the supernatant. This effect, taken alone, tends to cause precipitation of sodium salts. Second, caustic dilutes all the supernatant anions that were previously at their saturation concentrations, tending to cause dissolution of sodium salts. Only if the caustic increases the sodium concentration enough to offset the dilution of supernatant anions will dissolution be avoided (in which case precipitation of sodium salts will occur). The required amount of increase depends on which solids are available for dissolution, with 1:1 salts such as sodium nitrate requiring the most increase. Third, the caustic addition increases the alkalinity of the supernatant and therefore dissolves some portion of whatever aluminum hydroxide solids may be present.

Because of anion dilution, even caustic that is considerably denser than the supernatant can cause sodium salt waste to dissolve. Consider caustic that has a specific gravity of 1.5 and a concentration of 19 M NaOH. At this sodium concentration, sodium nitrate solid will dissolve (if it is present) unless the sodium molarity in the original supernatant is less than approximately 40% of the caustic concentration. Other solids contain more than one mole of sodium per mole of salt (such as sodium carbonate, sodium sulfate, sodium oxalate, sodium phosphate, sodium fluoride sulfate, and sodium fluoride phosphate) and are less likely to dissolve when strong caustic is added. Supernatant sodium concentrations can be in the range of 11 M and up without precipitating these salts when 19 M caustic is added. However, sodium molarity is unlikely to be this high unless solid sodium nitrate is present, in which case it would be the governing salt.

For a first example of caustic addition, consider the case of the common soluble salt sodium nitrate (NaNO_3). Because the equations in this example are developed using the certainly inaccurate assumption that the species activity equals the molality, they are strictly conceptual examples and should not be used for actual calculations. Whenever a solution is saturated with sodium nitrate, the concentration-product expression governing its dissolution equilibrium is

$$[\text{Na}^+] \llbracket \text{NO}_3^- \rrbracket = K_{\text{NaNO}_3} \quad (\text{A.2})$$

where

$[\text{Na}^+]$ = molal concentration of sodium (mol Na/kg water)

$[\text{NO}_3^-]$ = molal concentration of nitrate (mol NO_3^- /kg water)

K_{NaNO_3} = the concentration-based solubility product (molality)²

Assuming that the supernatant in a tank is known to be saturated with NaNO_3 , and that the supernatant concentrations of sodium and nitrate ions are known, the concentration-based K_{NaNO_3} can be found:

$$K_{\text{NaNO}_3} = [Na^+]_s [NO_3^-]_s \quad (\text{A.3})$$

In the above, the concentrations are those in the supernatant, denoted by s subscripts. When the added caustic, with a sodium molality $[Na^+]_c$, is mixed with the supernatant it changes the sodium and nitrate concentrations in a way that can be found by mass balance. Defining R as the ratio of the caustic and supernatant water masses, the result is

$$[Na^+] = \frac{R[Na^+]_c + [Na^+]_s}{1 + R} \quad (\text{A.4})$$

$$[NO_3^-] = \frac{[NO_3^-]_s}{1 + R} \quad (\text{A.5})$$

Only if the product of the new sodium and nitrate molalities equals the value for the original supernatant, which equaled K_{NaNO_3} , does no net dissolution or precipitation occur. If the new product is less, dissolution occurs to increase the concentration until the product equals K_{NaNO_3} . If it is greater, precipitation occurs and continues until the concentrations in the liquid have been reduced enough to allow the product to equal K_{NaNO_3} .

The condition for the breakpoint, where no dissolution or precipitation of sodium nitrate occurs, is

$$[Na^+] [NO_3^-] = [Na^+]_s [NO_3^-]_s \quad (\text{A.6})$$

Because the equation sets a product equal to a product, the approximation introduced by using molal concentrations instead of activities is decreased. That is, if the equation was expressed exactly, in terms of activities rather than molalities, the activity coefficients would be on both sides and partly cancel each other out. (However, the remaining inaccuracy can still be large.)

The condition in Equation (A.6) is met when

$$\left(\frac{R[Na^+]_c + [Na^+]_s}{1 + R} \right) \left(\frac{[NO_3^-]_s}{1 + R} \right) = [Na^+]_s [NO_3^-]_s \quad (\text{A.7})$$

Equation (A.7) can be solved for the caustic molality that is necessary to maintain the NaNO_3 solids as is:

$$[Na^+]_c = [Na^+]_s (R + 2) \quad (\text{A.8})$$

Conceptually, the molality of the caustic must be a minimum of twice the molality of sodium in the supernatant to avoid dissolution of sodium nitrate (if any is present). The greater the amount of caustic that is added, the more concentrated it needs to be to avoid dissolution. Above the breakpoint caustic concentration, sodium nitrate will be precipitated.

The same sort of approach can be applied to the other waste solids. Making the same inaccurate but simplifying assumption about activities as before, the general concentration-based solubility expression for saturated solutions of sodium salts is

$$[Na^+]^a [A^{-a}] = K_{Na_a A} \quad (A.9)$$

where the chemical equation for the dissolution of the generalized salt is



The condition for avoidance of dissolution or precipitation of the generalized salt, after caustic addition, is

$$\left(\frac{R[Na^+]_c + [Na^+]_s}{1+R} \right)^a \left(\frac{[A^{-a}]_s}{1+R} \right) = [Na^+]_s^a [A^{-a}]_s \quad (A.11)$$

which can be simplified to

$$[Na^+]_c = [Na^+]_s \frac{(1+R)^{(a+1)/a} - 1}{R} \quad (A.12)$$

For R nearly equal to zero (a relatively small amount of caustic added), the caustic concentration breakpoint is $(a+1)/a$ times the supernatant sodium concentration, with all units in molality. The breakpoint concentration increases with the amount of caustic added, though much less than proportionally. Conceptually, the more moles of sodium there are per mole of solid (the higher a is), the lower is the caustic concentration at the breakpoint between dissolution and precipitation (for equal caustic amounts, these being expressed in terms of the mass of water in the caustic).

A.3 Conceptual Application – Aluminum Dissolution

The addition of caustic can dissolve not only sodium salts but aluminum hydroxide, although the dissolution of the latter is much slower. The solubility of aluminum in solution depends on the formation of a hydroxide complex from the solid:



As a result the ratio of the aluminum and hydroxide ion activities is approximately constant, at least for fairly small changes in the hydroxide concentration. This approximation, together with

the quite inaccurate but conceptually simplifying assumption of using the molal concentrations of the species instead of the activities, can be used to explore the dissolution of Al(OH)_3 after caustic addition.

As a first step, find the post-addition hydroxide ion and dissolved aluminum concentrations before any aluminum hydroxide is dissolved. They are equal to

$$[\text{OH}^-]_i = \frac{(M_{wc} [\text{OH}^-]_c + M_{ws} [\text{OH}^-]_s)}{M_{wc} + M_{ws}} \quad (\text{A.14})$$

$$[\text{Al}]_i = [\text{Al}]_s \frac{M_{ws}}{M_{wc} + M_{ws}} \quad (\text{A.15})$$

where M_{wc} and M_{ws} are the masses of water in the caustic addition and in the supernatant, respectively.

At the point when aluminum hydroxide dissolution is complete, assuming that saturated conditions exist as they did before caustic addition, it is approximately true that

$$\frac{[\text{Al}]}{[\text{OH}^-]} = \frac{[\text{Al}]_s}{[\text{OH}^-]_s} \quad (\text{A.16})$$

A further expression is needed to relate the post-dissolution concentrations ($[\text{Al}]$ and $[\text{OH}^-]$) to those that existed after caustic addition but before dissolution ($[\text{Al}]_i$ and $[\text{OH}^-]_i$). Because hydroxide ion is consumed to form the aluminum hydroxide complex and bring aluminum into solution, a mole balance gives

$$[\text{Al}] - [\text{Al}]_i = [\text{OH}^-]_i - [\text{OH}^-] \quad (\text{A.17})$$

Combining equations yields

$$[\text{Al}] - [\text{Al}]_s \frac{M_{ws}}{M_{wc} + M_{ws}} = \frac{(M_{wc} [\text{OH}^-]_c + M_{ws} [\text{OH}^-]_s)}{M_{wc} + M_{ws}} - [\text{Al}] \frac{[\text{OH}^-]_s}{[\text{Al}]_s} \quad (\text{A.18})$$

$$[\text{Al}] = \frac{M_{wc} [\text{OH}^-]_c + M_{ws} ([\text{OH}^-]_s + [\text{Al}]_s)}{M_{wc} + M_{ws}} \left(\frac{[\text{Al}]_s}{[\text{Al}]_s + [\text{OH}^-]_s} \right) \quad (\text{A.19})$$

The change in the number of moles of aluminum in solution, ΔAl , is equal to

$$\Delta Al = (M_{wc} + M_{ws}) [\text{Al}] - M_{ws} [\text{Al}]_s \quad (\text{A.20})$$

$$\Delta Al = \frac{M_{wc} [OH^-]_c [Al]_s}{[Al]_s + [OH^-]_s} \quad (A.21)$$

The change in the moles of dissolved aluminum is always positive; note, though, that this result is driven by the limitations of the assumptions used and may not be general. This somewhat unintuitive outcome results largely from the approximation that the Al/OH molal ratio is a constant at saturation.

A.4 Conclusions

The discussion in this appendix leads to several conclusions:

- The smaller the amount of caustic that is added, the lower is the minimum caustic concentration required to avoid dissolution of salt solids. Small amounts of high-concentration caustic are less likely to cause salt dissolution than larger amounts of low-concentration caustic. However, any addition of caustic that increases the hydroxide concentration is likely to dissolve aluminum hydroxide solid, if any is present.
- The caustic amount and concentration could be such that low- α salt solids such as NaNO₃ dissolve while high- α salt solids such as Na₇F(PO₄)₂.19H₂O precipitate. So long as the main modeling concern is to determine the amount of dissolution of existing solids, which is the driver for short-term gas release, the precipitation of higher- α solids need not be calculated. Thus modeling that is aimed at estimating short-term gas releases can focus on the lower- α solids that are present.
- The net effect of caustic on waste solids cannot be calculated with any accuracy without some form of chemical modeling. The approach detailed above shows the general type of mass-balance and solubility complications that arise, but even so oversimplifies to the point of grave inaccuracy by leaving out the variations in activity coefficients. These can only be handled by a computational chemical thermodynamic model. (This is even truer for additions of diluted waste, where the composition effects are more complicated.) The identity and amount of the solids in the waste are crucial.

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